PERIODIC CLASSIFICATION OF ELEMENTS AND PERIODICITY  1. The number of shells in an element reflects its: (a) Period number (b) Group number (c) Both (d) None 2. The number of electrons in valence shell of an element reflects its: (a) Period number (b) Group number (c) Both (d) None 3. Generally metals form oxides: (a) Acidic (b) Basic (c) Amphoteric (d) None 4. Variable valency is shown by: (a) Group IA (b) Group IIA (c) Group VIIA (d) Group IIB 5. Hydrogen can be placed above the groups of the periodic table: (a) IA, IVA and VIIA elements (b) VIIA Elements (c) IIIA, IVIA and VIIA elements (d) IIA, IIIA and VIIA elements (c) Halogens 6. Members of group IA are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens 7. Members of group IB are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens 8. Keeping in view the sizes of atoms, which order is the correct one: (a) Mg>Sr (b) Ba > Mg. (c) Be>Mg (d) Ra>Ba  7. Members of group IIA are called: (a) Mg>Group IIA (b) Ray (c) Be>Mg (d) Ra>Ba  8. Keeping in view the sizes of atoms, which order is the correct one: (a) Mg>Group (b) Nav (c) SO: (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 KJ mol¹ (d) 480 KJ mol¹ (c) Govalent Compound (d) None of these  (a) +3 (b) +2 (c) +4 (d) +1  14. Hydrides can be classified into: (a) Two types (d) Five types  (b) Three types  (c) Four types (d) Five types  (d) Five types  18. Which of the following represents elements in order of increasing atomic radii: (a) I>Br> CI (b) Li>Na>K (c) He>Ne>Ar (d) None								
The number of shells in an element reflects its:   (a)   Period number (b)   Group number (c)   Both   (d)   None				CHAP	TER 1			
The number of shells in an element reflects its:   (a)   Period number (b)   Group number (c)   Both   (d)   None								
1. The number of shells in an element reflects its:  (a) Period number (b) Group number (c) Both (d) None  2. The number of electrons in valence shell of an element reflects its:  (a) Period number (b) Group number (c) Both (d) None  3. Generally metals form oxides:  (a) Acidic (b) Basic (c) Amphoteric (d) None  4. Variable valency is shown by:  (a) Group VIIA  (b) Group IIA  (c) Group VIIA  (d) Group IIB  Hydrogen can be placed above the groups of the periodic table:  (a) IA, IVA and VIIA elements (b) VIIA Elements  (c) IIIA, IVA and VA elements (d) IIA, IIIA and VIIA elements  (d) Alkaline earth metals  (c) Halogens (d) Noble gases  7. Members of group IB are called:  (a) Alkali metals (b) Alkaline earth metals  (c) Halogens (d) Noble gases  8. Keeping in view the sizes of atoms, which order is the correct one:  (a) Mg-Sr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  Ng-Sr (b) Ba > Mg (c) Be>Mg (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every:  (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in:  (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table:  (a) 13. Non-metals usually exist as:  (a) Liquids (c) 190 (d) 480 K.J mol¹  (c) 496 K.J mol¹  (d) 480 K.J mol¹  (e) 496 K.J mol¹  (f) 496 K.J mol¹  (g) 496 K.J mol¹  (h) 496 K.J mol¹  (h) 490 (d) 480 K.J mol¹  (h) 400 (e) 496 K.J mol¹  (h) 400 (f) Five types  18. Which of the following represents elements in order of increasing atomic radii:  18. Which of the following represents elements in order of increasing atomic radii:		PEKIO	DIC CLA				IS ANL	
(a) Period number (b) Group number (c) Both (d) None  The number of electrons in valence shell of an element reflects its:  (a) Period number (b) Group number (c) Both (d) None  3. Generally metals form oxides:  (a) Acidic (b) Basic (c) Amphoteric (d) None  4. Variable valency is shown by:  (a) Group 1A (b) Group IIA  (b) Group IIA  (c) Group VIIA (d) Group IB  5. Hydrogen can be placed above the groups of the periodic table:  (a) IA, IVA and VIIA elements (b) VIIA Elements  (c) IIIA, IVA and VIA elements (b) VIIA Elements  (c) IIIA, IVA and VIA elements (d) IIA, IIIA and VIIA elements  (c) Halogens (d) Noble gases  7. Members of group IB are called:  (a) Alkali metals (b) Alkaline earth metals  (c) Halogens (d) Noble gases  8. Keeping in view the sizes of atoms, which order is the correct one:  (a) Mg>Sr (b) Ba>Mg (c) SO2 (d) ZnO  9. Which one of the following oxides is amphoteric in nature:  (a) MgO (b) NazO (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every:  (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in:  (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table:  (a) 7 (b) 8 (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 K.J mol¹ (d) 480 K.J mol¹  (c) Go-ordinate Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) Two types (d) Three types  (b) Five types  18. Which of the following represents elements in order of increasing atomic radii:				PERIO	DICII	Y		
2. The number of electrons in valence shell of an element reflects its:  (a) Period number (b) Group number (c) Both (d) None  3. Generally metals form oxides:  (a) Acidic (b) Basic (c) Amphoteric (d) None  4. Variable valency is shown by:  (a) Group 1A (b) Group IIA  (c) Group VIIA  5. Hydrogen can be placed above the groups of the periodic table:  (a) IA, IVA and VIIA elements (b) VIIA Elements  (c) IIIA, IVA and VIA elements (d) IIIA, IIIA and VIIA elements  (c) IIIA, IVA and VIA elements (d) Noble gases  7. Members of group IB are called:  (a) Alkali metals (b) Alkaline earth metals  (c) Halogens (d) Noble gases  7. Members of group IB are called:  (a) Mg>Fr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  8. Keeping in view the sizes of atoms, which order is the correct one:  (a) MgO (b) Na₂O (c) Be>Mg (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of propertiest take place at every:  (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in:  (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table:  (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as:  (a) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 KJ mol¹ (d) 480 KJ mol¹  (c) 496 KJ mol² (d) 480 KJ mol²  (d) Five types  16. Members of lements recurrence in a decimal properties the correct one:  (a) 43 (d) 7 (d) 7 (d) 18  17. Hydrides can be classified into:  (a) Two types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	1.	The number	of shells in a	n element refle	ects its:			
(a)   Period number (b)   Group number (c)   Both (d)   None	2			-			: X : X	None
3. Generally metals form oxides: (a) Acidic (b) Basic (c) Amphoteric (d) None 4. Variable valency is shown by: (a) Group PIIA (b) Group IIA (c) Group VIIA (d) Group IIA (d) Group IIA (d) Group IIA (d) Group IIA (e) Group VIIA (e) Group VIIA (f) Group IIA (g) IIA, IVA and VIIA elements (g) IIIA, IVA and VIIA elements (g) Alkali metals (g	۷.							None
4. Variable valency is shown by:  (a) Group VIIA (b) Group IIA (c) Group VIIA (d) Group IIB  5. Hydrogen can be placed above the groups of the periodic table: (a) IA, IVA and VIIA elements (b) VIIA Elements (c) IIIA, IVA and VIIA and vIIIA elements (d) IIA, IIIA and VIIA elements  6. Members of group IA are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens (d) Noble gases  7. Members of group IB are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens (d) Noble gases  8. Keeping in view the sizes of atoms, which order is the correct one: (a) Mg-Sr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  9. Which one of the following oxides is Amphoteric in nature: (a) MgO (b) Nato (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids or gases (d) Waxy solids  14. The Lonization energy of Sodium is: (a) 50t K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> (e) 496 K.J mol <sup>-1</sup> (f) Ga has oxidation state: (a) 43 (b) +2 (c) +4 (d) +1  14. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	3.	Generally me	etals form ox	100 Jan 1990				
(a) Group IA (b) Group IIA (c) Group VIIA (c) Group VIIA (d) Group IIA Group IIA Group IIA (d) Group IIA Group IIA Group IIA (d) Group IIA Group IIA Group IIA Group IIA Group IIA (d) Group IIA Gro	4	\ 7	( )		(c)	Amphoteric	(d)	None
(c) Group VIIA	4.		X <del></del>	n by:	(b)	Group IIA		
(a) IA, IVA and VIIA elements (b) VIIA Elements (c) IIIA, IVA and VA elements (d) IIA, IIIA and VIIA elements  6. Members of group IA are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens (d) Noble gases  7. Members of group IB are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens (d) Noble gases  7. Members of group IB are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens (d) Coinage metals  8. Keeping in view the sizes of atoms, which order is the correct one: (a) Mg>Sr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  9. Which one of the following oxides is Amphoteric in nature: (a) MgO (b) NavO (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (b) Gases (c) Liquids or gases (d) Waxy solids  15. SnCla is a: (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:		(c) Group	VIIA		(d)	Group IB	<u>6 1 ''</u>	2
(c) IIIA, IVA and VA elements (d) IIA, IIIA and VIIA elements  Members of group IA are called: (a) Alkali metals (d) Noble gases  7. Members of group IB are called: (a) Alkali metals (b) Alkaline earth metals (c) Halogens (d) Noble gases  8. Keeping in view the sizes of atoms, which order is the correct one: (a) Mg>Sr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  9. Which one of the following oxides is Amphoteric in nature: (a) MgO (b) NavO (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) SOO K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) 17 Wydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	5.		1 <del>77</del> 8			### (Fig. 1)	pakcity.	org 🚜
6.   Members of group IA are called:   (a)							A elements	
(c) Halogens 7. Members of group IB are called: (a) Alkali metals (c) Halogens (d) Alkaline earth metals (c) Halogens (d) Alkaline earth metals (c) Halogens (d) Coinage metals  8. Keeping in view the sizes of atoms, which order is the correct one: (a) Mg>Sr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  9. Which one of the following oxides is Amphoteric in nature: (a) MgO (b) NazO (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl4 is a: (a) Co-ordinate Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	6.				(u)	IIA, IIIA dha VIIA	\ Cicinchis	
7.						1 1 V	etals	
(a) Alkali metals (b) Alkaline earth metals (c) Halogens  8. Keeping in view the sizes of atoms, which order is the correct one: (a) Mg>Sr (b) Ba > Mg (c) Be> Mg (d) Ra>Ba  9. Which one of the following oxides is Amphoteric in nature: (a) MgO (b) NazO (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a: (a) Co-ordinate Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	7			called	(d)	Noble gases		
8. Keeping in view the sizes of atoms, which order is the correct one:  (a) Mg>Sr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  9. Which one of the following oxides is Amphoteric in nature:  (a) MgO (b) Na2O (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every:  (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in:  (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table:  (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as:  (a) Liquids (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types  (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	7.		_	caneu.	(b)>>	Alkaline earth me	etals	
(a) Mg>Sr (b) Ba > Mg (c) Be>Mg (d) Ra>Ba  9. Which one of the following oxides; is Amphoteric in nature: (a) MgO (b) NazO (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a: (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:		(c) Halog	ens		( _ / \ \ )	Coinage metals		
9. Which one of the following oxides is Amphoteric in nature:  (a) MgO (b) Na2O (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every:  (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in:  (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table:  (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as:  (a) Liquids (b) Gases  (c) Liquids (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl4 is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound  (c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types  (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	8.	the teat of the te	o company		· > · ·			Dox Do
(a) MgO (b) NazO (c) SO2 (d) ZnO  10. According to Newland's arrangement of elements recurrence (periodicity) of properties take place at every: (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a: (a) Co-ordinate Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	9.	` '		T 2/ 4 \	C. A		(a)	Ka>ba
take place at every:  (a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in:  (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table:  (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as:  (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	# (#)	2 A 2 2		_ ( )	- TE	전 점	(d)	ZnO
(a) 8th element (b) 10th element (c) 18th element (d) None  11. Atomic number was discovered by Mosley in: (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a: (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	10.			rrangement of	elements	recurrence (peri	odicity) of p	properties
11. Atomic number was discovered by Mosley in:  (a) 1913 (b) 1914 (c) 1915 (d) 1916  12. Total groups in modern periodic table: (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a: (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:		ere ere III. o ara is	- ()	10 <sup>th</sup> elemen	t (c)	18 <sup>th</sup> element	(d)	None
12. Total groups in modern periodic table:  (a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as:  (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	11.		( VIV				(-)	
(a) 7 (b) 8 (c) 10 (d) 18  13. Non-metals usually exist as: (a) Liquids (b) Gases (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is: (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a: (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	40	<b>(</b> )			(c)	1915	(d)	1916
13. Non-metals usually exist as:  (a) Liquids (c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	12.		Z1 N	eriodic table:	(c)	10	(d)	18
(c) Liquids or gases (d) Waxy solids  14. The Ionization energy of Sodium is:  (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3  (b) +2  (c) +4  (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	13.		\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	as:			(4)	10
14. The Ionization energy of Sodium is:  (a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:								
(a) 500 K.J mol <sup>-1</sup> (b) 513 K.J mol <sup>-1</sup> (c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	14		0=15.	Sodium is:	(d)	Waxy solids		
(c) 496 K.J mol <sup>-1</sup> (d) 480 K.J mol <sup>-1</sup> 15. SnCl <sub>4</sub> is a:  (a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	470		,	pakci	<b>ty</b> (b)	513 K.J mol <sup>-1</sup>		
(a) Co-ordinate Covalent Compound (b) Ionic compound (c) Covalent Compound (d) None of these  16. Ga has oxidation state: (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into: (a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:		(c) 496 K	.J mol <sup>-1</sup>		*** **********************************	480 K.J mol <sup>-1</sup>		
(c) Covalent Compound (d) None of these  16. Ga has oxidation state:  (a) +3 (b) +2 (c) +4 (d) +1  17. Hydrides can be classified into:  (a) Two types (b) Three types  (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	15.		dinate Covaler	nt Compound	(b)	Ionic compound		
<ul> <li>16. Ga has oxidation state: <ul> <li>(a) +3</li> <li>(b) +2</li> <li>(c) +4</li> <li>(d) +1</li> </ul> </li> <li>17. Hydrides can be classified into: <ul> <li>(a) Two types</li> <li>(b) Three types</li> <li>(c) Four types</li> <li>(d) Five types</li> </ul> </li> <li>18. Which of the following represents elements in order of increasing atomic radii:</li> </ul>				•	. ,			
<ul> <li>17. Hydrides can be classified into:         <ul> <li>(a) Two types</li> <li>(b) Three types</li> <li>(c) Four types</li> <li>(d) Five types</li> </ul> </li> <li>18. Which of the following represents elements in order of increasing atomic radii:</li> </ul>	16.	(전) (전) 1.1.1 (	.5			The second secon		
(a) Two types (b) Three types (c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	4-7	•			(c)	+4	(d)	+1
(c) Four types (d) Five types  18. Which of the following represents elements in order of increasing atomic radii:	17.	2 <b>.</b>		u into:	(h)	Three types		
		(語) (表) (二字)	E.O. E.		프 교환	7=400 S (S)		
(a) I>Br> Cl (b) Li>Na>K (c) He>Ne>Ar (d) None	18.	Which of the	following re		v.=	der of increasing	7.15	
		(a) I>Br>	Cl (b)	Li>Na>K	(c)	He>Ne>Ar	(d)	None

19.						ns, because of t	he increase i	n numbe
	751. 550	=5	Tanel	ns and laye		is known as:		
	(a)	Resonance e Inductive eff			(d)	Shielding effect		
20.	(c) Numbe		11.12.00.00	nt in the 5 <sup>th</sup>	(d) neriod of a	None periodic table is:	•	
20.	(a)	8	(b)	10	(c)	18	<b>.</b> (d)	32
21.		element has			\ \		(4)	<b>3.</b>
	(a)	Li	(b)	Na	(c)	K	(d)	Rb
22.	12 72	of the follow	ving pair	s are chem	53A gag ( 90)			
	(a)	Na and K			(b)	Ba and Sr		
23.	(C)	Zr and Hf of the follow	vina olon	onte ie ma	(d)	Ca and Zn		
23.	(a)	Oxygen	(b)	Chlorine	(c)	Nitrogen	(d)	Fluorine
24.		of the follow	\ - J		``		(4)	Haomin
n <del></del> 0-	(a)	Mg	(b)	Ca	(c)	Al	(d)	Cs
25.	Highes	st Hydration	energy is	277 IA 19 🖦	7	3700 4 400 <del>-</del>	y: powy	20.00
	(a)	Na <sup>+</sup>	(b)	Mg <sup>+2</sup>	(c)	Al <sup>+3</sup>	(d)	Ga <sup>+3</sup>
26.		- car		<u> </u>	7. 3	ation Energy:	/ ـ / \	ΛI+3
27.	(a) The va	Na+ Jence shell <i>e</i>	(b) electronic	Al <sup>+</sup> - structure	of an elem	Al <sup>+2</sup> ent is ns²np⁵. Th	(d) he element w	Al <sup>+3</sup> vill belond
		group:		. Stracture	or an elem	A	ic cicilicit W	pelong
	(a)	IA	(b)	IIA	(c)	VA (29)	(d)	VIIA
28.	Which	of the follow	ving pair	of atomic	numbers re	presents IIA ele	ements:	
100000000000000000000000000000000000000	(a)	3, 11	(b)	3, 12	(c)	4   .	(d)	3, 20
29.		the followi		ents the hig	hest value	of electron affir		by:
20	(a)	too of attrac	(b)	Cl n alamant /	on charad r	Br <b>air of electron i</b>	(d)	1
30.	(a)	Covalent bor		i element	(b)	Ionization poter		
	(c)	Electron affi		(0)	(d)	Electronegativit		
31.	` '	metals in ea	The state of the s	have:			,	
	(a)	Smallest size		M	(b)	Lowest Ionization	on Energy	
	(c)	Highest Ioni	1 7	5-01-0 Laboratoria	(d)	Lowest atomic r	radius	
32.	191	rrect order o	of electro	n affinity a				
	(a)	F > Cl > Br Cl > F > Br	1/2		(b)	Br > Cl > F F > Br > Cl		
33.	(c) Which	10.00 PM PM 10.00 PM 10.00 PM	vina doe	s not exhib	(d) it the "peri	odicity" in prop	erties of the (	elements
<b>JJ.</b>	(a)	Ionization er	32 Table 1		(b)	n/p ratio		
	(c)	Electronegat		Y), a	(d)	Atomic radius		
34.	Polymo	eric (Interm	ediate) h	ydride sha	ll be forme	d by:	8 \$500	
<b>~</b> =	(a)	Na -baskasidik	(b)	K	(c)	Be	(d)	С
35.	2.3	ghest acidity		A PARTY	city.org	MnOs	(4)	MnO
36.	(a) The co	Mn <sub>2</sub> O <sub>7</sub> valent hydri	(b) des are u	Mn <sub>2</sub> O <sub>3</sub>	(c)	MnO <sub>2</sub>	(d)	MnO
J.	(a)	Liquids	ui e t	Jamiyi	(b)	Gases		
	(c)	Volatile Liqu	ids or gase	es	(d)	Waxy solids		
<b>37.</b>		of the follow				<b>*</b> 2		
	(a)	Na <sup>+</sup>	(b)	Mg <sup>2+</sup>	(c)	O <sup>2-</sup>	(d)	Cl-
	(Sec. 1984)	of the follow	70001.00	(1929) NATO	70G T.W	SCHOOL DES	2 K	Cl
38.	(a)	CaO	(b)	CO <sub>2</sub>	(c)	SiO <sub>2</sub>	(d)	Sb <sub>2</sub> O <sub>3</sub>
	(a)	<b>elting point</b> i Be	(b)	Mg	(c)	Ca	(d)	Sr
38. 39.	(~)	<b>-</b> 5	(-)	. יש	(~)		(4)	<del>Ų</del> .

(a) They are borderline elements that exhibit both metallic and non-metallic properties (b) They usually act as electron donors to non-metals (c) They usually act as electron acceptors from metals (d) They are good conductor of heat and electricity  41. Among the pure ionic compounds, the have the highest lattice energies: (a) Fluorides (b) Chlorides (c) Bromides (d) Iodides  42. Which electronic sub-shell in lanthanides is incompletely filled: (a) 4f (b) 5f (c) 6f (d) All  43. The number of elements in fourth period of periodic table is: (GRW 2011) (a) 32 (b) 18 (c) 10 (d) 8  44. The basis of modern periodic law is: (FSD, BWP, GRW 2012) (a) Electron affinity (b) atomic mass (c) Ionization energy (d) atomic number  45. The highest ionization energy is possessed by: (GRW 2011) (a) Nitrogen (b) Phosphorous (c) Bismuth (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of: (a) 5 & p-block (b) d-block (c) f-block (d) All  47. Mark the correct statement: (a) All lanthanides are present in the same group (b) All halogens are present in the same group (d) All the noble gases are present in the same period (c) All the alkali metals are present in the same period (d) All the noble gases are present in the same period (d) Metallic character increases down the group (b) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same down the group  19. In Potassium super oxide (KO2) oxidation state of O is: (a) -1 (b) -2 (c) -1/2 (d) -4	40	Which	of the fe	Mossin	a in n	at tua	for m	otalla:	da.					
(b) They usually act as electron donors to non-metals (c) They usually act as electron acceptors from metals (d) They are good conductor of heat and electricity  41. Among the pure ionic compounds, the have the highest lattice energies:  (a) Fluorides (b) Chlorides (c) Bromides (d) Iodides  42. Which electronic sub-shell in lanthanides is incompletely filled: (a) 4f (b) 5f (c) 6f (d) All  43. The number of elements in fourth period of periodic table is: (GRW 2011) (a) 32 (b) 18 (c) 10 (d) 8  44. The basis of modern periodic law is: (FSD, BWP, GRW 2012) (a) Electron affinity (b) atomic mass (c) Ionization energy (d) atomic mass (c) Ionization energy (d) atomic number (GRW 2011) (a) Nitrogen (b) Phosphorous (GRW 2011) (b) Rismuth (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of: (a) s & p-block (d) All  47. Mark the correct statement: (a) All lanthanides are present in the same group (d) All the noble gases are present in the same period. (a) All the noble gases are present in the same period. (b) Metallic character increases along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a peri	40.				3-3					allic an	d non-r	netallic pr	ropertie	ns.
(d) They are good conductor of heat and electricity  41. Among the pure ionic compounds, the		선생님 이를										, i e de la	. sps. s.s	
41. Among the pure ionic compounds, the have the highest lattice energies:  (a) Fluorides (b) Chlorides (c) Bromides (d) Iodides  42. Which electronic sub-shell in lanthanides is incompletely filled:  (a) 4f (b) 5f (c) 6f (d) All  43. The number of elements in fourth period of periodic table is:  (a) 32 (b) 18 (c) 10 (d) 8  44. The basis of modern periodic law is:  (a) Electron affinity (b) atomic mass (C) Ionization energy (d) atomic number  45. The highest ionization energy is possessed by:  (a) Nitrogen (c) Bismuth (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of:  (a) s & p-block (d) All  47. Mark the correct statement:  (a) All lanthanides are present in the same group  (b) All halogens are present in the same period.  48. Encircle the correct statement:  (a) Metallic character increases along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (e) Possess the property of catenation (d) Form neutral oxides   ANSWER KEY   ANSWER KEY   ANSWER KEY   ANSWER KEY   ANSWER KEY   ANSWER KEY    1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  a d c c c a b b a b c  21 22 23 24 25 26 27 28 29 30  a d d d d c d d c b d d b a  41 42 43 44 44 45 46 47 48 49 50		(c)	055											
(a) Fluorides (b) Chlorides (c) Bromides  42. Which electronic sub-shell in lanthanides is incompletely filled:  (a) 4f (b) 5f (c) 6f (d) All  43. The number of elements in fourth period of periodic table is: (GRW 2011)  (a) 32 (b) 18 (c) 10 (d) 8  44. The basis of modern periodic law is: (FSD, BWP, GRW 2012)  (a) Electron affinity (b) atomic mass (C) Ionization energy (d) atomic number  45. The highest ionization energy is possessed by: (GRW 2011)  (a) Nitrogen (b) Phosphorous (GBW 2011)  (b) All though (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of:  (a) 8 & p-block (d) All  47. Mark the correct statement:  (a) All lanthanides are present in the same group  (b) All halogens are present in the same group  (d) All the alkall metals are present in the same period.  48. Encircle the correct statement:  (a) Metallic character increases down the group  (b) Metallic character remains the same down the group  19. In Potassium super oxide (KO2) oxidation state of O is:  (a) -1 (b) -2 (c) -1/2 (d) -4  40. Hydrogen can be placed with the element of group (IV-A) because both:  (a) Act as strong oxidizing agent (b) Act as strong reducing agent  (c) Possess the property of catenation (d) Form neutral oxides  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER (EY)		A. (8)	(AT)	10-10-10-10-10-10-10-10-10-10-10-10-10-1							- · · · · · · · · · · · · · · · · · · ·	_ =22	<b>2</b> 7	
42. Which electronic sub-shell in lanthanides is incompletely filled:  (a) 4f (b) 5f (c) 66 (d) All  43. The number of elements in fourth period of periodic table is: (GRW 2011)  (a) 32 (b) 18 (c) 10 (d) 8  44. The basis of modern periodic law is: (FSD, BWP, GRW 2012)  (a) Electron affinity (b) atomic mass (c) Ionization energy (d) atomic number  (b) All milest ionization energy is possessed by: (GRW 2011)  (a) Nitrogen (b) Phosphorous  (b) Phosphorous  (c) Bismuth (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of:  (a) s & p-block (b) d-block  (b) d-block  (c) f-block (d) All  47. Mark the correct statement:  (a) All lanthanides are present in the same group  (b) All the noble gases are present in the same period.  48. Encircle the correct statement:  (a) Metallic character increases down the group  (b) Metallic character increases down the group  (b) Metallic character remains the same along a period  (c) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same along a period  (d) Metallic character remains the same down the group  10. In Potassium super oxide (KO2) oxidation state of 0 is:  (a) -1 (b) -2 (c) -1/2 (d) -4  11. 12 13 14 15 16 17 18 19 20  a b b d a a d b b d a  11. 12 13 14 15 16 17 18 19 20  a d c c c c a b b a b c  21. 22 23 24 25 26 27 28 29 30  a d d c c c c a b d b b a  31. 32 33 34 35 36 37 38 39 40  b c b c a c d d d b a  41. 42 43 44 45 46 47 48 49 50	41.	ese ser	-1		E-01 -00			E 10		1, 1952				Todidos
(a) 4f (b) 5f (c) of (d) All  43. The number of elements in fourth period of periodic table is:	42											(0	u)	iodiaes
43. The number of elements in fourth period of periodic table is: (a) 32 (b) 18 (c) 10 (d) 8 (44. The basis of modern periodic law is: (a) Electron affinity (b) atomic mass (C) Ionization energy (d) atomic number (d) Antimony  45. The highest ionization energy is possessed by: (a) Nitrogen (b) Phosphorous (GRW 2011)  46. The decrease in atomic sizes is much prominent across rows containing elements of: (a) s & p-block (b) d-block (c) f-block (d) All anthanides are present in the same group (d) All halogens are present in the same group (d) All the noble gases are present in the same group (d) All the noble gases are present in the same group (d) Metallic character increases down the group (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same down the group (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same down the group (1/2 (d) 4  49. In Potassium super oxide (KO2) oxidation state of O is: (a) -1 (b) -2 (c) -1/2 (d) -4  Hydrogen can be placed with the element of group (IV-A) because both: (a) Act as strong oxidizing agent (b) Act as strong reducing agent (c) Possess the property of catenation (d) Form neutral oxides  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER (EY)  ANSWER (E	741		-	iic Sub			namac			-	iicu:	((	d)	All
44. The basis of modern periodic law is:  (a) Electron affinity (b) atomic mass (c) Lonization energy (d) Antimony  45. The highest ionization energy is possessed by: (a) Nitrogen (b) Phosphorous (c) Bismuth (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of: (a) s & p-block (b) d-block (c) f-block (d) All lanthanides are present in the same group (b) All halogens are present in the same period (c) All the alkali metals are present in the same period (d) All the noble gases are present in the same period (d) All the noble gases are present in the same operiod.  48. Encircle the correct statement: (a) Metallic character increases down the group (b) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Possess the property of catenation (d) Form neutral oxides   ANSWER KEY	43.			f eleme			n perio		eriodic	table	is:		100	
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(c) Ionization energy (d) atomic number  The highest ionization energy is possessed by:  (a) Nitrogen (b) Phosphorous (c) Bismuth (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of: (a) s & p-block (b) d-block (c) F-block (d) All  47. Mark the correct statement: (a) All lanthanides are present in the same group (b) All halogens are present in the same period (c) All the alkali metals are present in the same period.  48. Encircle the correct statement: (a) Metallic character increases down the group (b) Metallic character remains the same along a period (d) Metallic character remains the same down the group  49. In Potassium super oxide (kO2) oxidation state of O is: (a) -1 (b) -2 (c) -1/2 (d) -4  Hydrogen can be placed with the element of group (IV-A) because both: (a) Act as strong oxidizing agent (b) Act as strong reducing agent (c) Possess the property of catenation (d) Form neutral oxides  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER (A) 4 5 6 7 8 9 10.  a b b d a a d b b d a  11 12 13 14 15 16 17 18 19 20  a d c c c c a b a b c  21 22 23 24 25 26 27 28 29 30  a d d d d c d d c b d  31 32 33 34 35 36 37 38 39 40  b c b c a c d d b a  41 42 43 44 45 46 47 48 49 50	44.					lic law	is:	4.3				(FSD, BW	VP, GRV	N 2012)
45. The highest ionization energy is possessed by:  (a) Nitrogen (b) Phosphorous (c) Bismuth (d) Antimony  46. The decrease in atomic sizes is much prominent across rows containing elements of: (a) s & p-block (b) d-block (c) f-block (d) All  47. Mark the correct statement: (a) All lanthanides are present in the same group (b) All halogens are present in the same group (d) All the noble gases are present in the same period (c) All the noble gases are present in the same period (d) Metallic character increases down the group (b) Metallic character increases along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same along a period (d) Metallic character remains the same down the group  49. In Potassium super oxide (KO2) oxidation state of O is: (a) -1 (b) -2 (c) -1/2 (d) -4  Hydrogen can be placed with the element of group (IV-A) because both: (a) Act as strong oxidizing agent (c) Possess the property of catenation (d) Form neutral oxides  ANSWER KEY		Ž., Š.						7.13						
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1.       2.       3.       4.       5.       6.       7.       8.       9.       10.         a       b       b       d       a       d       b       d       a         11       12       13       14       15       16       17       18       19       20         a       d       c       c       c       a       b       a       b       c         21       22       23       24       25       26       27       28       29       30         a       d       d       d       c       d       d       c       b       d         31       32       33       34       35       36       37       38       39       40         b       c       b       c       a       c       d       d       b       a         41       42       43       44       45       46       47       48       49       50	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the note that the control of the	charact charact charact charact charact super o	tatem ter increase ter increase ter remainder ter remainder (b) ced windizing	e preser ent: eases d eases a ains the ains the -2 th the agent	own the long a same same eleme	e group beriod along a down t on stat (c) ent of g	period.	is: /2 IV-A) et as str	ong red	se both: ducing ag		-4
a     b     b     d     a     a     d     b     d     a       11     12     13     14     15     16     17     18     19     20       a     d     c     c     c     a     b     a     b     c       21     22     23     24     25     26     27     28     29     30       a     d     d     d     c     d     d     c     b     d       31     32     33     34     35     36     37     38     39     40       b     c     b     c     a     c     d     d     b     a       41     42     43     44     45     46     47     48     49     50	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the note that the control of the	charact charact charact charact charact super o	tatem ter increase ter increase ter remainder ter remainder (b) ced windizing	e preser ent: eases d eases a ains the ains the -2 th the agent of caten	own the long a same same exidation	e group beriod along a down t (c) nt of g (b) (d)	period.	is: /2 IV-A) et as str	ong red	se both: ducing ag		-4
11     12     13     14     15     16     17     18     19     20       a     d     c     c     c     a     b     a     b     c       21     22     23     24     25     26     27     28     29     30       a     d     d     d     c     d     d     c     b     d       31     32     33     34     35     36     37     38     39     40       b     c     b     c     a     c     d     d     b     a       41     42     43     44     45     46     47     48     49     50	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the note that the control of the	charact charact charact charact charact super o	tatem ter increase ter increase ter remainder ter remainder (b) ced windizing	e preser ent: eases d eases a ains the ains the -2 th the agent of caten	own the long a same same exidation	e group beriod along a down t (c) nt of g (b) (d)	period.	is: /2 IV-A) et as str	ong red	se both: ducing ag		-4
a     d     c     c     a     b     a     b     c       21     22     23     24     25     26     27     28     29     30       a     d     d     d     c     d     d     c     b     d       31     32     33     34     35     36     37     38     39     40       b     c     b     c     a     c     d     d     b     a       41     42     43     44     45     46     47     48     49     50	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the related the control of the c	charact charact charact charact super of the proof the p	ter increase removed (b)  ced with a sidizing operty of the sidicing operation.	ent: eases deases a ains the ains the agent of caten  4.	own the long a same same same same same same same sa	e group beriod along a down t on stat (c) nt of g (b) (d)	period. period	is: /2 IV-A) et as strorm neu	ong red tral oxi	ty.org		-4
21         22         23         24         25         26         27         28         29         30           a         d         d         d         c         b         d         d         c         b         d           31         32         33         34         35         36         37         38         39         40           b         c         b         c         a         c         d         d         b         a           41         42         43         44         45         46         47         48         49         50	<b>49.</b>	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the related the control of the c	charact charact charact charact super of the proof the p	ter increase removed (b)  ced with a sidizing operty of the boundary of the control of the contr	ent: eases deases a ains the ains the agent of caten  4. d	own the long a same same same same same same same sa	e group beriod along a down t on stat (c) nt of g (b) (d)	period.  per	is: /2 IV-A) et as strorm neu  8. b	ong red tral oxi	ty.org		-4
a     d     d     c     d     d     c     b     d       31     32     33     34     35     36     37     38     39     40       b     c     b     c     a     c     d     d     b     a       41     42     43     44     45     46     47     48     49     50	<b>49.</b>	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the related the control of the c	charactecharac	ter increase removed (b)  ced with a sidizing operty of the boundary of the control of the contr	ent: eases deases a ains the ains the agent of caten  4. d	own the long a same same same same same same same sa	e group beriod along a down t on stat (c) nt of g (b) (d)	period.  per	is: /2 IV-A) et as strorm neu	pakci gpakci	ty.org		-4
31     32     33     34     35     36     37     38     39     40       b     c     b     c     a     c     d     d     b     a       41     42     43     44     45     46     47     48     49     50	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the relected the condition of the co	charactecharac	ses are statem ter increater removable (b) ced windizing operty of the statement of the sta	ent: eases deases a ains the ains the agent of caten  4. d 14. c	own the long a same same ation  NSW  5.  a 15  C	e group beriod along a down to n state (c) nt of g (b) (d)	period.  per	is: /2 IV-A) et as strorm neu	pakci pakci b	ty.org		-4
b c b c a c d d b a 41 42 43 44 45 46 47 48 49 50	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the related the content of the c	charactecharac	ses are statem ter increter removed (b) ced windizing operty of the ced windizing oper	ent: eases deases a ains the ains the agent of caten  4. d 14. c	own the long a same same ation  NSW  5.  a 15  C	e group beriod along a down t on stat (c) nt of g (b) (d)	period.  per	is: /2 IV-A) et as strorm neu	pakci pakci b	ty.org		-4
41 42 43 44 45 46 47 48 49 50	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the related the content of the c	charactecharac	ses are statem ter increater removable (b) ced windizing operty of the ced windizing o	ent: eases deases a ains the ains the agent of caten  A. d  14  C  24  d  d  14  C  24  d	own the long a same same ation  Solution  Solu	e group beriod along a down t on stat (c) nt of g (b) (d)	period.  per	is: /2 IV-A) ct as strorm neu B. b 18 a 28 c	pakci pakci b 29	ty.org  ty.org  c  a  20  c  30  d		-4
	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the related the content of the c	charactecharac	ses are statem ter increater removable (b) ced windizing operty of the ced windizing o	ent: eases deases a ains the ains the agent of caten  A. d  14  C  24  d  d  14  C  24  d	own the long a same same ation  Solution  Solution  Solution  Solution  Column 15  Column 25  Colum	e group beriod along a down t on stat (c) nt of g (b) (d)	period.  per	is: /2 IV-A) et as strorm neu b 18 a 28 c 38	pakci pakci b 29	ty.org  ty.org		-4
	<b>49.</b>	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the relected the condition of the co	charactecharac	ses are statem ter increter removide (b) ced windizing operty of the ced windizing ope	ent: eases deases a ains the ains the ains the agent of caten  A. d  14  C  24  d  14  C  24  d  14  C  24  d  14  C	own the long a same same ation  NSW  5.  a 15  C 25  C 35  a	e group beriod along a down to n state (c) nt of g (b) (d)	period. period	is: /2 IV-A) et as strorm neu 8. b 18 a 28 c 38 d	pakci	ty.org  ty.org		-4
	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the relected the condition of the co	charactecharac	ses are statem ter increter removable (b) ced windizing operty of the ced windizing op	ent: eases deases a ains the ains the ains the agent of caten  A. d  14  C  24  d  14  C  24  d  14  C  24  d  14  C	own the long a same same ation  NSW  5.  a 15  c 25  c 35  a 45	e group beriod along a down to n state (c) nt of g (b) (d) TERKE	period. period	is: /2 IV-A) et as strorm neu B. b 18 a 28 c 38 d 48	pakci	ty.org ty		-4
	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the relected the condition of the co	charactecharac	ses are statem ter increter removable (b) ced windizing operty of the ced windizing op	ent: eases deases a ains the ains the ains the agent of caten  A. d  14  C  24  d  14  C  24  d  14  C  24  d  14  C	own the long a same same ation  NSW  5.  a 15  c 25  c 35  a 45	e group beriod along a down to n state (c) nt of g (b) (d) TERKE	period. period	is: /2 IV-A) et as strorm neu B. b 18 a 28 c 38 d 48	pakci	ty.org ty		-4
	49.	(d) Encirc (a) (b) (c) (d) In Pot (a) Hydro (a)	All the relected the condition of the co	charactecharac	ses are statem ter increter removable (b) ced windizing operty of the ced windizing op	ent: eases deases a ains the ains the ains the agent of caten  A. d  14  C  24  d  14  C  24  d  14  C  24  d  14  C	own the long a same same ation  NSW  5.  a 15  c 25  c 35  a 45	e group beriod along a down to n state (c) nt of g (b) (d) TERKE	period. period	is: /2 IV-A) et as strorm neu B. b 18 a 28 c 38 d 48	pakci	ty.org ty		-4



1.	2.	3.	4.	<sub>18</sub> 5.ci	6.	7.	8.	9.	10.
a	b	b	d	а	а	d	b	d	а
11	12	13	14	15	16	17	18	19	20
а	d	С	С	С	а	b	a	b	С
21	22	23	24	25	26	27	28	29	30
а	d	d	d	С	d	d	С	b	d
31	32	33	34	35	36	37	38	39	40
b	С	b	С	a	С	d	d	b	а
41	42	43	44	45	46	47	48	49	50
а	а	b	d	а	а	С	а	С	b



	CHAP	TER 2			
	S-BLOCK	ELEMI	Pakcity.org		
				•	
1.	Which of the following sulphates is not (a) Sodium sulphate	soluble in (b)	<b>n water:</b> Potassium sulphate		
	(c) Zinc sulphate	(d)	Barium sulphate		
2.	Chile Salt peter has the chemical formul	la: Č	•		
	(a) NaNO₃	(b)	KNO <sub>3</sub>		
_	(c) Na <sub>2</sub> B <sub>4</sub> O <sub>7</sub>	(d)	Na <sub>2</sub> CO <sub>3</sub> .H <sub>2</sub> O	!4.1. 6	20 -
3.	Which one of the following compounds (a) Calcium chloride	71.3	Solved in water react	s with (	LO <sub>2</sub> :
	(a) Calcium chloride (c) Calcium Sulphate	(b) (d)	Calcium hydroxide		
4.	The main product of the reaction of pota	` '			
	(a) KOH (b) K <sub>2</sub> O <sub>2</sub>	(c)	KO <sub>2</sub>	(d)	$K_2O$
<b>5.</b>	The alkali metal whose carbonate is rel	latively le	ess stable to heat and	decon	nposes on
	heating, giving its oxides is:	2.3		7 B	
6	(a) Li (b) Ba	(c)	K Ja limlanasa	(d)	Rb
6.	Which of the following compounds has a (a) BaO <sub>2</sub> (b) CO <sub>2</sub>	a per oxio	PbO <sub>2</sub>	(d)	SiO <sub>2</sub>
7.	Which of the following does not belong	` /		(u)	3102
<b>*</b> •	(a) Be (b) Mg	(c)\s	Ra	(d)	Rb
8.	Which does not belong to alkali metals:	SHIP			
	(a) Cs (b) Fr	(Sc)	Na	(d)	Ca
9.	The oxides of beryllium are:				
	(a) Acidic	(b)	Neutral		
10.	(c) Basic Nelson's cell is used to prepare:	(d)	Amphoteric		
TO.	(a) Sodium hydroxide	(b)	Sodium carbonate		
	(c) Sodium metal	(d)	Sodium bicarbonate		
11.	What is deposited at the cathode during	the elec	trolysis of brine:		
	(a) Na (b) H <sub>2</sub>	(c)	Cl <sub>2</sub>	(d)	OH
12.	Which of the following is a rare radioact			. IN	D.I.
12	(a) Fr (b) Cs The chemical formula of Magnesite is:	(c)	Ra	(d)	Rb
13.	The chemical formula of Magnesite is: (a) MgCl <sub>2</sub>	(b)	Mg (ClO <sub>3</sub> ) <sub>2</sub>		
	(c) MgCO <sub>3</sub>	(d)	None of these		
14.	Which carbonate of alkali metals is inso	\ /			
	(a) $Na_2Co_3$ (b) $K_2CO_3$	(c)9	Li <sub>2</sub> CO <sub>3</sub>	(d)	Cs <sub>2</sub> CO <sub>3</sub>
<b>15.</b>	Ga has most common oxidation state of				
	(a) +3 (b) +2	(c)	+4	(d)	+1
16.	Alkali metals form:	<b>/</b> b)	Covalent companyed		
	<ul><li>(a) Ionic compounds</li><li>(c) Coordinate covalent compounds</li></ul>	(b) (d)	Covalent compounds None of these		
17.	Li <sub>2</sub> O is:	(u)	NOTIC OF LITESE		
-2:-	(a) Orange yellow solid	(b)	White solid		
	(c) Greenish solid	(d)	Pale yellow solid		
18.	Melting point of pure sodium chloride is	G 100 100 100 100 100 100 100 100 100 10	500		
	(a) 600 °C (b) 775 °C	(c)	750 °C	(d)	801 °C
19.	Lime (CaO) is obtained by thermal deco	E 15			# Face
	(a) $Ca(OH)_2$ (b) $CaCO_3$	(c)	CaHCO <sub>3</sub>	(d)	None

			Company of the Compan					
20.		ement neces	sary for	normal "lea				
	(a)	Phosphorus			(p)	Sulphur		
21.	(c)	Calcium s often used	261		(d)	Magnesium		
<b>ZI</b> .	(a)	Reducing Ac			(b)	Oxidizing Agent		
	(c)	Dehydrating			(d)	Catalytic Agent		
22.		m metal can	5.22	d in:	(-)	outsily are 7 i.go. i.e		
	(a)	Alcohol	(b)	Kerosene o	oil (c)	H <sub>2</sub> O	(d)	All
23.	Which		50.00 cap ( 50)	ACCIONE PROGRAM NO SELECTION	£ 62	the strongest base:	2 921	1201 - P121 W 166
5.4	(a)	LiOH	(b)	NaOH	(c)	KOH	(d)	CsOH
24.		ing soda has			Ц.Ω (a)	Na. CO. 3H.O	(4)	Na <sub>2</sub> CO <sub>2</sub>
25.	(a) Which	halide has t		Na <sub>2</sub> CO <sub>3</sub> .10		$Na_2CO_3.3H_2O$	(d)	Na <sub>2</sub> CO <sub>3</sub>
23.	(a)	NaCl	(b)	NaBr	(c)	NaF	(d)	Nal
26.					X /	y to form Nitride is:	(~)	. 101
100 mm	(a)	Li	(b)	Na	(c)	K	(d)	Rb
27.	The el	ectronic con	figuratio	n of metal (I	M) is $1s^2$ , $2$	2s <sup>2</sup> , 3s <sup>1</sup> . The formula o	f its ox	ide would
	be:							
20	(a)	MO	(b)	M <sub>2</sub> O	(c)	M <sub>2</sub> O <sub>3</sub>	(d)	MO <sub>2</sub>
28.		ueous solutio	on of sod	ium carbona	ite is alkai	ine because sodium ca	arbonat	te is a sali
	<b>of:</b> (a)	Weak acid a	nd waak l	hase	(b)	Strong acid and weak	hace	
	(a) (c)	Weak acid a			(d)	Strong acid and strong		
29.	<b>\</b>	of the follow			` '	A Congressions	buse	
	(a)	Ra	(b)	Fr	(c)	Rn	(d)	Cs
<b>30.</b>	Dolon	nite has the o	omposit	ion	1(3)			
	(a)	KCl.MgCl <sub>2</sub> .6ŀ	$H_2O$		(b)	Na <sub>3</sub> AlF <sub>6</sub>		
	(c)	CaCO <sub>3</sub> .MgC(	)3	200	(d)	CaCl <sub>2</sub> .MgCl <sub>2</sub> .6H <sub>2</sub> O		
31.		n salt is:		21212	4.3		i de la companya de	
	(a)	Magnesium	30	A Participation	(b)	Ferrous ammonium su	ilphate	
32.	(C)	Magnesium er of Paris is	11	m phosphate	(d)	Calcium sulphate		
<b>JZ.</b>	(a)	BaSO <sub>4</sub>	(b)	CaSO <sub>4</sub>	(c)	MgSO <sub>4</sub>	(d)	ZnSO <sub>4</sub>
33.			~ /			ives plaster of Paris:	(4)	211004
	(a)	Borax	(b)	Gypsum	(c)	Alum	(d)	Calome
34.	On he	ating quick l	ime with	coke in an	electric fu	rnace, we get:	Α 2	
	(a)	Ca and CO <sub>2</sub>	(b)	CaCO <sub>3</sub>	(c)	Ca+CO	(d)	CaC <sub>2</sub>
35.	2 3	ubstance not	likely to	contain Ca				
	(a)	Dolomite		SVA L	(b)	Marble		
36	(C)	Gypsum	lium ic		(d)	Sea shells		
36.	(a)	xide of Beryl Acidic	num 15:		(b)	Basic		
	(a) (c)	Acidic			(d)	None of these		
37.		out the ore	of potass	sium:	(~)	a denies si minas		
	(a)	Dolomite			(b)	Cryolite		
	(c)	Bauxite	ž.im.		(d)	Carnallite		
38.	Dolor	nite is an ore	of:		#F 8	20.2		
	(a)	Strontium			(b)	Magnesium		
20	(c)	Barium	lali mati	-12	(d)	Potassium		
39.		h is not an al			(c)	Dubidium	(d)	Padium
	(a)	Francium	(b)	Cesium	(c)	Rubidium	(d)	Radium

40. Chile salt peter has the chemical formula:  (a) NaNO3 (b) KNO3 (c) NasBr4O7 (d) NasCO3.H3O  41. The milk of magnesia is used for the treatment of:  (a) Acidity (b) Basicity (c) Rancidity (d) Jaundice  42. Cement contains gypsum:  (a) 3% (b) 2% (c) 0.2% (d) 0.3%  43. Dolomite is: (a) CaCO3 (b) MgCO3.CaCO3 (c) MgCO3 (d) Na2CO3  44. Chemical formula of Magnesite is: (a) CaMgs(SiO3)4 (b) MgCO3. (c) MgCO3 (d) MgCO3  45. Which one does not belong to alkaline earth metals?  (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15) (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell:  (BWP 14, FSD 11,13)  (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium (d) Barium sulphate (c) Potassium sulphate (c) Potassium sulphate (d) Barium sulphate (e) Potassium sulphate (f) Decrease solubility (g) Decrease solubility (h) Decrease dissociation (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY   ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  4 ANSWER KEY  ANSWER KEY  ANSWER KEY	(a) NaNO3 (b) KNO3 (c) Na2F03 H2O (d) Na2F03 (e) Rancidity (b) Basicity (c) Rancidity (d) Jaundice (d) Jaundice (a) 3% (b) 2% (c) 0.2% (d) 0.3% (d) 0.3% (d) Na2F03 (d													
C  Na <sub>2</sub> Br <sub>4</sub> O <sub>7</sub>	C  Na <sub>2</sub> Br <sub>3</sub> O <sub>7</sub>	40.	7.7		r has t	he che	emical	formu			N.O.			
41. The milk of magnesia is used for the treatment of:  (a) Acidity (b) Basicity (c) Rancidity (d) Jaundice  42. Cement contains gypsum: (a) 3% (b) 2% (c) 0.2% (d) 0.3%  43. Dolomite is: (a) CaCO3 (b) MgCO3.CaCO3 (c) MgCO3 (d) Na:CO3 (d) RGSD4 (d) MgCl2  45. Which one does not belong to alkaline earth metals? (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15) (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell: (FSD 10,14, BWP 10, LHR 14) (a) Na (b) H2 (c) Cl2 (d) O2  47. The ore CaSO4.2H2O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (SGD 2010) (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (d) Barium sulphate (c) Potassium sulphate (c) Potassium sulphate (d) Epsom salt  49. The sulphate compound insoluble in water is: (SGD 2010) (a) Berilling by (b) Decrease dissociation (c) CaClz is added to NaCl in Down's cell to: (d) Decrease conductivity   ANSWER KEY	41. The milk of magnesia is used for the treatment of:  (a) Acidity (b) Basicity (c) Rancidity (d) Jaundice  42. Cement contains gypsum: (a) 3% (b) 2% (c) 0.2% (d) 0.3%  43. Dolomite is: (a) CaCO3 (b) MgCO3.CaCO3 (c) MgCO3 (d) Na <sub>2</sub> CO3 (d) RgCO3 (d) RgCC3 (d) RgCC3 (d) RgCCC3 (d) RgCCCCCCCCCCCCCCCCCCCCCCCCCC		2 120		<u>.</u> .							٦-٨		
(a) Acidity (b) Basicity (c) Rancidity (d) Jaundice  42. Cement contains gypsum: (a) 3% (b) 2% (c) 0.2% (d) 0.3%  43. Dolomite is: (a) CaCO <sub>3</sub> (b) MgCO <sub>3</sub> .CaCO <sub>3</sub> (c) MgCO <sub>3</sub> (d) Na <sub>2</sub> CO <sub>3</sub> 44. Chemical formula of Magnesite is: (a) CaMg <sub>3</sub> (SiO <sub>3</sub> ) <sub>4</sub> (b) MgCO <sub>3</sub> (c) MgSO <sub>4</sub> (d) MgCO <sub>2</sub> 45. Which one does not belong to alkaline earth metals? (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15) (a) Be (b) Ra (c) Ba (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15) (a) Be (b) Ra (c) Ba (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15) (a) Na (b) H <sub>2</sub> (c) Cl <sub>2</sub> (d) O <sub>2</sub> 47. The substance deposited at the cathode during electrolysis of brine in diaphragm cell: (FSD 10,14, BWP 10, LHR 14) (a) Na (b) H <sub>2</sub> (c) Cl <sub>2</sub> (d) O <sub>2</sub> 48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcitum (d) Barium (sGD 2010) (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (sGD 2010) (a) Beryllium (b) Magnesium (d) Ersom salt  50. CaCl <sub>2</sub> is added to NaCl in Down's cell for (d) Decrease dissociation Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d c a a a d d d d a a 11 1 12 13 14 15 16 17 18 19 20 b c c c a a b d b c b c a 11 1 12 13 14 15 16 17 18 19 20 c b d b c a a b c b c b c a 13 1 32 33 34 35 36 37 38 39 40 a b b d c c d b b d a a 41 42 43 44 45 46 47 48 49 50	(a) Acidity (b) Basicity (c) Rancidity (c) Rancidity (d) Jaundice (Acidity (c) Rancidity (d) Jaundice (d) Jaundice (d) Jaundice (d) Jaundice (d) Jaundice (d) Jaundice (a) 3% (b) 2% (c) 0.2% (d) 0.3% (d) 0.3% (d) Na <sub>2</sub> CO <sub>3</sub> (e) MgSO <sub>4</sub> (d) MgCl <sub>2</sub> (a) CaMg <sub>3</sub> (SiO <sub>3</sub> ) <sub>4</sub> (b) MgCO <sub>3</sub> (c) MgSO <sub>4</sub> (d) MgCl <sub>2</sub> (d) Na <sub>2</sub> CO <sub>3</sub> (e) MgSO <sub>4</sub> (d) MgCl <sub>2</sub> (a) Be (b) Ra (c) Ba (d) Rn The substance deposited at the cathode during electrolysis of brine in diaphragm cell: (FSD 10,14, BWP 10, LHR 14) (a) Na (b) H <sub>2</sub> (c) Ba (d) Rn The substance deposited at the cathode during electrolysis of brine in diaphragm cell: (e) Dolomite (c) Calcite (d) Epsom salt (FSD 10,14, BWP 10, LHR 14) (a) Beryllium (b) Magnesium (c) Calcitum (d) Barium (s) Beryllium (b) Magnesium (c) Calcitum (d) Barium (s) Beryllium (b) Magnesium (c) Calcitum (d) Barium (s) Decrease solubility (c) Potassium sulphate (b) CaC <sub>2</sub> is added to NaCl in Down's cell to: (a) Decrease dissociation Decrease conductivity  ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.     d	41.				a is us	ed for	the tre			a2CO3 .1	120		
42. Cement contains gypsum:  (a) 3% (b) 2% (c) 0.2% (d) 0.3%  43. Dolomite is:  (a) CaCO <sub>3</sub> (b) MgCO <sub>3</sub> .CaCO <sub>3</sub> (c) MgCO <sub>3</sub> (d) Na <sub>2</sub> CO <sub>3</sub> 44. Chemical formula of Magnesite is:  (a) CaMg <sub>3</sub> (SiO <sub>3</sub> ) <sub>4</sub> (b) MgCO <sub>3</sub> (c) MgSO <sub>4</sub> (d) MgCl <sub>2</sub> 45. Which one does not belong to alkaline earth metals?  (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15)  (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell:  (FSD 10,14, BWP 10, LHR 14)  (a) Na (b) H <sub>2</sub> (c) Cl <sub>2</sub> (d) O <sub>2</sub> 47. The ore CaSO <sub>4-2</sub> H <sub>2</sub> O has general name:  (a) Gypsum (b) Dolomite  (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is:  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium  (b) Sodium sulphate  (c) Potassium sulphate  (c) Potassium sulphate  (d) Sodium sulphate  (e) Potassium sulphate  (fSD 2010)  (a) Beryllium (b) Decrease dissociation  (c) CaCl <sub>2</sub> is added to NaCl in Down's cell to:  (a) Decrease Melting point (d) Decrease conductivity  ANSWER KEY   ANSWER KEY   ANSWER KEY   (LHR 2012)  (d) Na <sub>2</sub> CO <sub>3</sub> (d) Na <sub>2</sub> CO <sub>3</sub> (d) MgCO <sub>3</sub> (d) MgCO <sub>3</sub> (e) MgSO <sub>4</sub> (fSD 2019, FSD 09,15, RWP 12, GRW 14,15)  (EFSD 10,14, BWP 10, LHR 14)  (b) Polomite  (FSD 10,14, BWP 10, LHR 14)  (e) Calcium (d) Barium  (fSGD 2010)  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium  (c) Calcium (d) Barium  (d) Barium  (sGD 2010)  (a) Barium sulphate  (b) Sodium sulphate  (c) Potassium sulphate  (c) Potassium sulphate  (d) Decrease dissociation  (sGD 2011)  (a) Decrease dissociation  (c) Decrease onductivity	42. Cement contains gypsum:  (a) 3% (b) 2% (c) 0.2% (d) 0.3%  43. Dolomite is:  (a) CaCO <sub>3</sub> (b) MgCO <sub>3</sub> .CaCO <sub>3</sub> (c) MgCO <sub>3</sub> (d) Na <sub>2</sub> CO <sub>3</sub> 44. Chemical formula of Magnesite is:  (AMGG) (a) CaMg <sub>3</sub> (SiO <sub>3</sub> ) <sub>4</sub> (b) MgCO <sub>3</sub> (c) MgSO <sub>4</sub> (d) MgCl <sub>2</sub> 45. Which one does not belong to alkaline earth metals?  (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15)  (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell:  (FSD 10,14, BWP 10, LHR 14)  (a) Na (b) H <sub>2</sub> (c) Cl <sub>2</sub> (d) O <sub>2</sub> 47. The ore CaSO <sub>4-2</sub> H <sub>2</sub> O has general name:  (a) Gypsum (b) Dolomite  (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is:  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (d) Barium (e) Contains uplphate  (c) Potassium sulphate  (d) Sodium sulphate  50. CaCl <sub>2</sub> is added to NaCl in Down's cell to:  (a) Beryllium (b) Decrease dissociation (d) Decrease onductivity  ANSWER KEY   ANSWER KEY   ANSWER KEY  (LHR 2012)  (d) Na <sub>2</sub> CO <sub>3</sub> (d) MgCO <sub>3</sub> (d) MgCO <sub>3</sub> (d) MgCO <sub>3</sub> (e) MgCO <sub>3</sub> (c) MgCO <sub>3</sub> (				-g						asicity			
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44. Chemical formula of Magnesite is: (a) CaMgs(SiO3)4 (b) MgCO3 (c) MgSO4 (d) MgCl2  45. Which one does not belong to alkaline earth metals? (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15) (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell: (FSD 10,14, BWP 10, LHR 14)  (a) Na (b) H2 (c) Cl2 (d) O2  47. The ore CaSO4.2H2O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (SGD 2010) (a) Beryllium (b) Magnesium (c) Calcium (d) Barium  49. The sulphate compound insoluble in water is: (SGD 2010) (a) Barium sulphate (c) Potassium sulphate (c) Potassium sulphate (c) Decrease Melting point (d) Decrease dissociation (a) Decrease Melting point (d) Decrease conductivity  ANSWER KEY   ANSWER KEY	44. Chemical formula of Magnesite is: (a) CaMg3(SiO3)4 (b) MgCO3 (c) MgSO4 (d) MgCl2  45. Which one does not belong to alkaline earth metals? (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15) (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell: (FSD 10,14, BWP 10, LHR 14) (a) Na (b) H2 (c) Cl2 (d) O2  47. The ore CaSO4.2H2O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (c) Potassium sulphate (c) Potassium sulphate (c) Decrease Melting point (d) Decrease dissociation (a) Decrease solubility (b) Decrease dissociation (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  d a d c a a a d d d a a 11 12 13 14 15 16 17 18 19 20  b c c c a a a b d b c 21 22 23 24 25 26 27 28 29 30  c b d b c a b c b c 31 32 33 34 35 36 37 38 39 40  a b b d c c d d b d a 41 42 43 44 45 46 47 48 49 50	43.				(5)	270		(0)	O.	12.70			
(a) CaMg3(SiO3)4 (b) MgCO3 (c) MgSO4 (d) MgCl2  Which one does not belong to alkaline earth metals?  (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell:  (FSD 10,14, BWP 10, LHR 14)  (a) Na (b) H2 (c) Cl2 (d) O2  47. The ore CaSO4.2H2O has general name:  (a) Gypsum (b) Dolomite  (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (SGD 2010)  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (d) Barium sulphate (c) Potassium sulphate  (b) Sodium sulphate  (c) Potassium sulphate (d) Epsom salt  49. The sulphate compound insoluble in water is: (SGD 2010)  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (SGD 2010)  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (SGD 2010)  (b) Decrease dissociation (SGD 2011)  (a) Decrease solubility (b) Decrease dissociation (SGD 2011)  (b) Decrease dissociation (SGD 2011)  (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  d a d c a a d d d d a  11 12 13 14 15 16 17 18 19 20  b c c c a a b d b c  21 22 23 24 25 26 27 28 29 30  c b d b c a b c b c  31 32 33 34 35 36 37 38 39 40  a b b d c c d b d a  41 42 43 44 45 46 47 48 49 50	(a) CaMg3(SiO3)4 (b) MgCO3 (c) MgSO4 (d) MgCl2  Which one does not belong to alkaline earth metals?  (a) Be (b) Ra (c) Ba (d) Rn  46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell: (FSD 10,14, BWP 10, LHR 14)  (a) Na (b) H2 (c) Cl2 (d) O2  47. The ore CaSO4.2H2O has general name: (BWP 14, FSD 11,13)  (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (SGD 2010)  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (SGD 2010)  (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (b) Sodium sulphate  (c) Potassium sulphate (d) Decrease dissociation (SGD 2011)  (a) Decrease Solubility (b) Decrease conductivity  ANSWER KEY		X 2						) <sub>3</sub> (c)	M	gCO₃			X 92
## Which one does not belong to alkaline earth metals?  (MTN 08,15, FSD 09,15, RWP 12, GRW 14,15)  (a) Be (b) Ra (c) Ba (d) Rn  The substance deposited at the cathode during electrolysis of brine in diaphragm cell:  (FSD 10,14, BWP 10, LHR 14)  (a) Na (b) H2 (c) Cl2 (d) O2  ## Calcite (d) Epsom salt  ## (SGD 2010)  (a) Beryllium (b) Magnesium (c) Calcium (d) Barium  ## (SGD 2010)  (a) Barium sulphate (b) CaCl2 is added to NaCl in Down's cell to:  (a) Decrease Melting point (d) Decrease conductivity  ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   1.   1.   1.   1.   1.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY    1.   2.   3.   4.   5.   6.   7.   8.   9.   10.    ## ANSWER KEY	Which one does not belong to alkaline earth metals?	44.				111			(c)	М	αSΩ₄			1,500
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46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell:	46. The substance deposited at the cathode during electrolysis of brine in diaphragm cell:  (a) Na (b) H2 (c) Cl2 (d) O2  47. The ore CaSO4.2H2O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium (d) Barium  49. The sulphate compound insoluble in water is: (a) Barium sulphate (c) Potassium sulphate (c) Potassium sulphate (d) Decrease dissociation (e) Decrease Melting point (d) Decrease conductivity  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  d a d c a a d d d d a  11 12 13 14 15 16 17 18 19 20  b c c c a a b d b c  21 22 23 24 25 26 27 28 29 30  c b d b c a b c b c  31 32 33 34 35 36 37 38 39 40  a b b d c c d b d a  41 42 43 44 45 46 47 48 49 50		(-N	D		71-3	<b>D</b> =		7 00		17.0	09,15,	RWP 1	
(a) Na (b) H <sub>2</sub> (c) Cl <sub>2</sub> (d) O <sub>2</sub> 47. The ore CaSO <sub>4-2</sub> H <sub>2</sub> O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (SGD 2010) (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (a) Barium sulphate compound insoluble in water is: (SGD 2010) (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (b) CaCl <sub>2</sub> is added to NaCl in Down's cell to: (a) Decrease solubility (b) Decrease dissociation (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d a d d d a a d d d d a a d c a a a d d d d	(a) Na (b) H <sub>2</sub> (c) Cl <sub>2</sub> (d) O <sub>2</sub> 47. The ore CaSO <sub>4</sub> .2H <sub>2</sub> O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcitum (d) Barium (d) Barium (e) Potassium sulphate (formation of the sulphate compound insoluble in water is: (a) Barium sulphate (c) Potassium sulphate (d) Barium (formation of the sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (d) Barium (formation of the sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (d) Barium (sGD 2010) (sGD 2010) (a) Decrease dissociation (b) Decrease dissociation (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  d a d c a a d d d d a a d d d d a a d d d d	46			deno			athode				s of bri	ine in d	<b>\</b> /
47. The ore CaSO <sub>4</sub> .2H <sub>2</sub> O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium (d) Epsom salt  49. The sulphate compound insoluble in water is: (a) Barium sulphate (c) Potassium sulphate (c) Potassium sulphate (d) Zinc sulphate (e) Potassium sulphate (f) Decrease dissociation (graph and the sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (d) Zinc sulphate (sGD 2010)  (a) Decrease solubility (b) Decrease dissociation (c) Decrease Melting point  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  d a d c a a d d d d a  11 12 13 14 15 16 17 18 19 20  b c c c a a b d b c  21 22 23 24 25 26 27 28 29 30  c b d b c a b c b c  31 32 33 34 35 36 37 38 39 40  a b b d c c d b d a  41 42 43 44 45 46 47 48 49 50	47. The ore CaSO4.2H <sub>2</sub> O has general name: (a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (sGD 2010) (d) Barium (sGD 2010) (d) Barium (sGD 2010) (e) Potassium sulphate (c) Potassium sulphate (d) Sodium sulphate (e) Potassium sulphate (f) Decrease dissociation (ga) Decrease solubility (ga) Decrease Melting point (ga) Decrease Melting point (ga) Decrease dissociation (ga) Decrease Melting point (ga) Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. (ga) Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. (ga) Decrease dissociation (ga) Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. (ga) Decrease dissociation		341	Julioc	. acpos	u			. 441111	<sub>3</sub> 3.66				
(a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (SGD 2010) (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (SGD 2010)  (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (c) Potassium sulphate (d) Zinc sulphate  50. CaCl <sub>2</sub> is added to NaCl in Down's cell to: (SGD 2011)  (a) Decrease solubility (b) Decrease dissociation (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d d d a 11 12 13 14 15 16 17 18 19 20 b c c a a d b d b c 21 22 23 24 25 26 27 28 29 30 c b d b c a b c b c 31 32 33 34 35 36 37 38 39 40 a b b d c c d d b d a 41 42 43 44 45 46 47 48 49 50	(a) Gypsum (b) Dolomite (c) Calcite (d) Epsom salt  48. The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium (d) Barium (SGD 2010)  (a) Barium sulphate compound insoluble in water is: (a) Barium sulphate (c) Potassium sulphate (c) Potassium sulphate (d) Zinc sulphate  50. CaCl <sub>2</sub> is added to NaCl in Down's cell to: (a) Decrease solubility (b) Decrease dissociation (d) Decrease conductivity  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d d a a d d d d a a d d d d a a d d d d a a d d d d d a a d d d d d a a d	47			211.0					C	2		/DW/D	2000
(c) Calcite (d) Epsom salt  The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium  The sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (c) Potassium sulphate (d) Decrease dissociation (e) Decrease Melting point  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d d a a 11 12 13 14 15 16 17 18 19 20 b c c c a a a b d b c c 21 22 23 24 25 26 27 28 29 30 c b d b c a b c b c 31 32 33 34 35 36 37 38 39 40 a b b d c c d b d a 41 42 43 44 45 46 47 48 49 50	(c) Calcite (d) Epsom salt  The only alkaline earth metal which forms peroxide is: (a) Beryllium (b) Magnesium (c) Calcium  The sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (d) Zinc sulphate  Tanch Service (d) Sodium sulphate  Tanch Service (d) Sodium sulphate  Tanch Sulphate (c) Potassium sulphate  Tanch Sulphate (d) Decrease dissociation (sGD 2010)  (a) Decrease solubility (b) Decrease dissociation (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10. 4. 4. 4. 5. 6. 5. 6. 7. 8. 9. 10. 5. 6. 6. 7. 8. 9. 10. 6. 7. 8. 9. 9. 10. 6. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7. 9. 7.	4/.				nas ge	eneral	name:		D	olomite		(RMb	14, FSD 11,13)
(a) Beryllium (b) Magnesium (c) Calcium (dd) Barium (SGD 2010)  The sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (dd) Zinc sulphate  50. CaCl <sub>2</sub> is added to NaCl in Down's cell to: (a) Decrease solubility (b) Decrease dissociation (c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. da a d c a a a d d d a a 11 12 13 14 15 16 17 18 19 20 b c c c a a a b d b c a b c b c a 11 12 22 23 24 25 26 27 28 29 30 c b d b c a b c b c a 11 32 33 34 35 36 37 38 39 40 a b b d c c d b d a 41 42 43 44 45 46 47 48 49 50	(a) Beryllium (b) Magnesium (c) Calcium (GGD 2010)  The sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate (c) Potassium sulphate (d) Zinc sulphate  Tinc sulphate Zinc sulphate  Tinc sulphate  Tinc sulphate  Tinc sulphate  Sodium sulphate  Zinc sulphate  (SGD 2011)  (SGD 2011)  (SGD 2011)  (SGD 2011)  ANSWER KEY   ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10. da a d d d a a a d d d d a a a d d d d			#./: #.:										
49. The sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate Zinc sulphate  SGD 2011)  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d c a a d d d d a 11 12 13 14 15 16 17 18 19 20 b c c c a a b d b c 21 22 23 24 25 26 27 28 29 30 c b d b c a b c b c 31 32 33 34 35 36 37 38 39 40 a b b d c c d b d a 41 42 43 44 45 46 47 48 49 50	49. The sulphate compound insoluble in water is: (a) Barium sulphate (b) Sodium sulphate Zinc sulpha	48.				5501 -00					11 >			
(a) Barium sulphate (b) Sodium sulphate Zinc	(a) Barium sulphate (b) Sodium sulphate Zinc sulphate Zinc sulphate Zinc sulphate (c) Potassium sulphate (d) Zinc sulphate Zinc sulphate Zinc sulphate (d) Zinc sulphate Zinc sulphate (d) Zinc sulphate (d) Zinc sulphate (d) Zinc sulphate Zinc sulphate (d) Decrease dissociation Decrease conductivity (e) Decrease conductivity (f) Decrease conductivity (h) Decrease conductivity (f) D	49.	. ,				_			7	aicium			X /
50. CaCl <sub>2</sub> is added to NaCl in Down's cell to:  (a) Decrease solubility (b) Decrease dissociation Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  d a d c a a d d d a a 11 12 13 14 15 16 17 18 19 20  b c c c a a b d b c a b c a a b c a b c a b c a a a b c a a b c a a b c a a b c a a b c a a b c a a b c a a b c a a a b c a a a b c a a a b c a a a b c a a a b c a a a b c a a a b a a a a	50. CaCl₂ is added to NaCl in Down's cell to:  (a) Decrease solubility (b) Decrease dissociation Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10.  d a d c a a a d d d a a 11 12 13 14 15 16 17 18 19 20  b c c c a a b d b c 21 22 23 24 25 26 27 28 29 30  c b d b c a b c b c 31 32 33 34 35 36 37 38 39 40  a b b d c c d b d a 41 42 43 44 45 46 47 48 49 50		월 월.	<del></del>	2000 BBD					is)Š	odium s	ulphate	)	(0.02, 2020)
(a) Decrease solubility Decrease dissociation Decrease Melting point (b) Decrease dissociation Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d d d a a d d d d a a d d d d a a d d d d d a a d d d d d a a d	(a) Decrease solubility Decrease dissociation Decrease Melting point (b) Decrease dissociation Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d d d a a d d d d a a d d d d a a d d d d d a a d d d d d a a d d d d d a a d	<b>F</b> 0	1.X 1/2/2				/-	II &	7 (gg)	Z	inc sulp	hate		(CCD 2011)
(c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d c a a d d d a 11 12 13 14 15 16 17 18 19 20 b c c c a a a b d b c 21 22 23 24 25 26 27 28 29 30 c b d b c a b c b c 31 32 33 34 35 36 37 38 39 40 a b b d c c d b d a 41 42 43 44 45 46 47 48 49 50	(c) Decrease Melting point (d) Decrease conductivity  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10. d a d c a a d d d a a 11 12 13 14 15 16 17 18 19 20 b c c c a a b d b c a b c a 21 22 23 24 25 26 27 28 29 30 c b d b c a b c b c a b c a b c a b c a b c a b c b c	50.					own s	cell to	$\sim$	D	ecrease	dissoci	iation	(SGD 2011)
1.       2.       3.       4.       5.       6.       7.       8.       9.       10.         d       a       d       c       a       a       d       d       d       a         11       12       13       14       15       16       17       18       19       20         b       c       c       c       a       a       b       d       b       c         21       22       23       24       25       26       27       28       29       30         c       b       d       b       c       a       b       c       b       c         31       32       33       34       35       36       37       38       39       40         a       b       b       d       c       c       d       b       d       a         41       42       43       44       45       46       47       48       49       50	1.       2.       3.       4.       5.       6.       7.       8.       9.       10.         d       a       d       c       a       a       d       d       d       a         11       12       13       14       15       16       17       18       19       20         b       c       c       c       a       a       b       d       b       c         21       22       23       24       25       26       27       28       29       30         c       b       d       b       c       a       b       c       b       c         31       32       33       34       35       36       37       38       39       40         a       b       b       d       c       c       d       b       d       a         41       42       43       44       45       46       47       48       49       50		<u> </u>				t	(D)						
1.       2.       3.       4.       5.       6.       7.       8.       9.       10.         d       a       d       c       a       a       d       d       d       a         11       12       13       14       15       16       17       18       19       20         b       c       c       c       a       a       b       d       b       c         21       22       23       24       25       26       27       28       29       30         c       b       d       b       c       a       b       c       b       c         31       32       33       34       35       36       37       38       39       40         a       b       b       d       c       c       d       b       d       a         41       42       43       44       45       46       47       48       49       50	1.       2.       3.       4.       5.       6.       7.       8.       9.       10.         d       a       d       c       a       a       d       d       d       a         11       12       13       14       15       16       17       18       19       20         b       c       c       c       a       a       b       d       b       c         21       22       23       24       25       26       27       28       29       30         c       b       d       b       c       a       b       c       b       c         31       32       33       34       35       36       37       38       39       40         a       b       b       d       c       c       d       b       d       a         41       42       43       44       45       46       47       48       49       50						NAIS	70						
d     a     d     c     a     a     d     d     d     a       11     12     13     14     15     16     17     18     19     20       b     c     c     c     a     a     b     d     b     c       21     22     23     24     25     26     27     28     29     30       c     b     d     b     c     a     b     c     b     c       31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50	d     a     d     c     a     a     d     d     d     a       11     12     13     14     15     16     17     18     19     20       b     c     c     c     a     a     b     d     b     c       21     22     23     24     25     26     27     28     29     30       c     b     d     b     c     a     b     c     b     c       31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50								ED I/E	1				
11       12       13       14       15       16       17       18       19       20         b       c       c       c       a       a       b       d       b       c         21       22       23       24       25       26       27       28       29       30         c       b       d       b       c       a       b       c       b       c         31       32       33       34       35       36       37       38       39       40         a       b       b       d       c       c       d       b       d       a         41       42       43       44       45       46       47       48       49       50	11       12       13       14       15       16       17       18       19       20         b       c       c       c       a       a       b       d       b       c         21       22       23       24       25       26       27       28       29       30         c       b       d       b       c       a       b       c       b       c         31       32       33       34       35       36       37       38       39       40         a       b       b       d       c       c       d       b       d       a         41       42       43       44       45       46       47       48       49       50					N	A	NSW	ER KE	<u>Y</u>				
b     c     c     c     a     a     b     d     b     c       21     22     23     24     25     26     27     28     29     30       c     b     d     b     c     a     b     c     b     c       31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d       41     42     43     44     45     46     47     48     49     50	b     c     c     c     a     b     d     b     c       21     22     23     24     25     26     27     28     29     30       c     b     d     b     c     a     b     c     b     c       31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50			1.	2.	3.	/4/3			3/1/	8.	9.	10.	
21       22       23       24       25       26       27       28       29       30         c       b       d       b       c       a       b       c       b       c         31       32       33       34       35       36       37       38       39       40         a       b       b       d       c       c       d       b       d       a         41       42       43       44       45       46       47       48       49       50	21     22     23     24     25     26     27     28     29     30       c     b     d     b     c     a     b     c     b     c       31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50			<b>1.</b> d		. / 20	/4/3	5.	6.	7.		î		
c     b     d     b     c     a     b     c     b     c       31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50	c     b     d     b     c     a     b     c     b     c       31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50			751 194	a	d/l	<b>4.</b> C	<b>5.</b>	<b>6.</b> a	<b>7.</b> d	d	d	а	
31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50	31     32     33     34     35     36     37     38     39     40       a     b     b     d     c     c     d     b     d     a       41     42     43     44     45     46     47     48     49     50			<b>11</b>	а <b>12</b> с	d 13 c	4. C 14	<b>5.</b> a <b>15</b> a	<b>6.</b> a <b>16</b> a	7. d 17 b	18 d	d <b>19</b> b	а <b>20</b> с	
a b b d c c d b d a 41 42 43 44 45 46 47 48 49 50	a b b d c c d b d a 41 42 43 44 45 46 47 48 49 50			11 b 21	a 12 c 22	d 13 c 23	4. C 14 C 24	5. a 15 a 25	6. a 16 a 26	7. d 17 b	d 18 d 28	d 19 b 29	a 20 c 30	
41 42 43 44 45 46 47 48 49 50	41 42 43 44 45 46 47 48 49 50			11 b 21 c	a 12 c 22 b	d 13 c 23 d	4. C 14 C 24 b	5. a 15 a 25 c	6. a 16 a 26 a	7. d 17 b 27 b	d 18 d 28 c	d 19 b 29	a 20 c 30 c	
a b b d b a d a c	a b b d b a d a c			11 b 21 c 31	a 12 c 22 b	d 13 c 23 d	4. C 14 C 24 b	5. a 15 a 25 c	6. a 16 a 26 a	7. d 17 b 27 b 37	d 18 d 28 c 38	d 19 b 29 b 39	a 20 c 30 c 40	
				11 b 21 c 31	a 12 c 22 b 32 b	d 13 c 23 d 33 b	4. C 14 C 24 b 34 d	5. a 15 a 25 c 35	6. a 16 a 26 a 36	7. d 17 b 27 b 37 d	d 18 d 28 c 38 b	d 19 b 29 b 39 d	a 20 c 30 c 40 a	
				11 b 21 c 31 a 41	a 12 c 22 b 32 b 42	d 13 c 23 d 33 b	4. C 14 C 24 b 34 d 44	5. a 15 a 25 c 35 c 45	6. a 16 a 26 a 36 c 46	7. d 17 b 27 b 37 d 47	d 18 d 28 c 38 b	d 19 b 29 b 39 d 49	a 20 c 30 c 40 a 50	
				11 b 21 c 31 a 41	a 12 c 22 b 32 b 42	d 13 c 23 d 33 b	4. C 14 C 24 b 34 d 44	5. a 15 a 25 c 35 c 45	6. a 16 a 26 a 36 c 46	7. d 17 b 27 b 37 d 47	d 18 d 28 c 38 b	d 19 b 29 b 39 d 49	a 20 c 30 c 40 a 50	

1.	2.	<b>∀3.</b> /	4.	5.	6.	7.	8.	9.	10.
d	а	d/	С	a	a	d	d	d	a
11	12	13	14	15	16	17	18	19	20
b	С	C	C	а	a	b	d	b	С
21	22	23	24	25	26	27	28	29	30
С	b	d	b	С	a	b	C	b	С
31	32	33	34	35	36	37	38	39	40
a	b	b	d	C	c	d	b	d	a
41	42	43	44	45	46	47	48	49	50
a	b	b	b	d	b	a	d	a	С

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		A figure and small man	•		
	CHA	APTER 3			
	CDOUD TIT A AND A	CBALIB	TV/ A ELEMER		nakaitu ara 8
	GROUP III A AND	JKUUP	IV A ELEMEN		pakcity.org &
1.	Gibbsite contain water molecul (a) 6 (b) 5	<b>es:</b> (c)	4	(d)	3
2.	Sindur used by Indian women is cher	` '		(4)	J
9 <b>-</b>	(a) PbO (b) PbO <sub>2</sub>	(c)	Pb <sub>3</sub> O <sub>4</sub>	(d)	PbCO <sub>3</sub>
3.	Corundum is: (a) Al <sub>2</sub> O <sub>3</sub>	(h)	Na <sub>3</sub> AIF <sub>6</sub>		
	(a) $Al_2O_3$ (c) $Al_2O_3.2H_2O$	(b) (d)	Al <sub>2</sub> O <sub>3</sub> .3H <sub>2</sub> O		
4.	Which one of the following is not an	X 2	3.		
argani (	(a) Corundum (b) Bauxite	(c)	Colemanite	(d)	Kaolin
5.	Which one of the following is not a us	71.3	Visit a		
	<ul><li>(a) Antiseptic</li><li>(c) Stiffening agent for candle wick</li></ul>	(b) (d)	Glaze Lubricant		
6.	(c) Stiffening agent for candle wick Which of the following reactions of A				
0.	(a) $2 \text{ Al} + 3 \text{ H}_2 \rightarrow 2 \text{AlH}_3$	(b)	4 Al + $3^{\circ}Q_2 \rightarrow 2Al_2Q_3$	3	
	(c) $2 AI + N_2 \rightarrow 2AIN$	(d)	2 Al 4 3Cl <sub>2</sub> →2AICI <sub>3</sub>	ĺ	
<b>7.</b>	Two elements frequently used for ma	aking transis	stors are:		
	(a) C and Si	(b)	Ga and In		
0	(c) P and As Which one of the following is not a w	(d) co of rad loa	Si and Ge		
8.	Which one of the following is not a use (a) Red Pigment	(b)	Flint Glass		
		90) (d)	Semiconductor		
9.	In the dried up lakes of Tibet and Cal		is found:		
WA WAY	(a) Boric acid (b) Coleman		Borax	(d)	All
10.	What is the formula of Kaolin (Clay):		41.0		
	(a) Al <sub>2</sub> O <sub>3</sub> .SiO <sub>4</sub> (c) Al <sub>2</sub> O <sub>3</sub> .2H <sub>2</sub> O	(b)	Al <sub>2</sub> O <sub>3</sub> Al <sub>2</sub> O <sub>3</sub> , 2SiO <sub>2</sub> .2H <sub>2</sub> O		
11.	Boric acid is formed when borax reac	(d)	A12O3, 23IO2 .2I12O		
	(a) NaCl (b) NaOH	(c)	HCI	(d)	H <sub>2</sub> CO <sub>3</sub>
12.	Ethyl borate is formed when boric ac	id is reacted	l with:		
	(a) Elhyl Choride	(b)	Ethyl Bormide		
12	(c) Ethyl Alcohol	(d)	Ethyl Acetate		
13.	Which of the following is a weak acid (a) Na <sub>2</sub> SO <sub>4</sub>	(b)	HCI		
	(c) Boric Acid	ceity (d)	None of these		
14.	Talc is used in:				
	(a) Talcum powders	(b)	Face powders		
2 <b>-</b> -	(c) Making of house hold articles	(d)	All		
15.	The Chemical formula of Lead sub ox		Dh₂O₂	(4)	Dh <sub>2</sub> O
16.	(a) PbO (b) Pb <sub>2</sub> O <b>The Chemical formula for white lead</b>	(c) <b>is:</b>	Pb <sub>2</sub> O <sub>3</sub>	(d)	Pb₃O <sub>4</sub>
= <b>X</b> .	(a) PbCO <sub>3</sub> (b) Pb <sub>3</sub> O <sub>4</sub>	(c)	2PbCO <sub>3</sub> .Pb(OH) <sub>2</sub>	(d)	Pb <sub>2</sub> O <sub>3</sub>
<b>17.</b>	Which metal is protected by a layer of	× 120	⇒ (M 27)		- · · <u>· ·</u>
1 <u>2</u> 8228	(a) Al (b) Zn	(c)	Sn	(d)	Pb
18.	Inert pair effect plays an important r	09 19		7.45	DL
	(a) F (b) Al	(c)	Si	(d)	Pb

19.	Alum	is not used:				
	(a)	To jam Radar	(b)	To insulate buildings	;	
	(c)	Construction of ships	(d)	Making milk storage	tanks	
20.	Al is l	padly corroded by:				
	(a)	Pure water	(b)	Salt solutions		
	(c)	Dil.H <sub>2</sub> SO <sub>4</sub>	(d)	Dil.HNO₃		
21.	300	is an example of:	21.5			
	(a)	Ionic hydride	(b)	Molecular addition co	ompound	
22	(c)	Good oxidizing agent	(d)	None		
22.	75.1 (SE)	inium reacts with caustic soda		Aluminium ovido		
	(a)	Aluminium hydroxide Sodium aluminium hydroxide	(b) (d)	Aluminium oxide		
23.	(c)	ounds of Boron behave as Lev		None use of:		
23.	(a)	Electron donation	(b)	Electron deficiency of	of Boron	
	(c)	Non-metallic nature of Boron	(d)	Small size of Boron	DOIOII	
24.		cts as acid according to the co		S.Han Size of Doron		
	(a)	Lewis	(b)	Bronsted-Lowry		
	(c)	Arrhenius	(d)	None		
25.		on-polar oxide is:	(-)	e, encommentations		
	(a)	$H_2O$ (b) $CO_2$	(c)	CO	(d) All	
26.	The s	emiconductor material among	following is:		X = 20	
	(a)	Si (b) Ge	(c)	PbS(25)	(d) All	
27.	In Bo	orax bead test, when borax is	heated with co	balt oxide it forms	bead of:	
	(a)	Black colour	(b)	⟨Blue colour		
	(c)	Red colour	(a) (5)	Green colour		
28.	Chen	nical composition of Colemani			RW 06, BWP 12)	
	(a)	$Ca_2B_6O_{11}.5H_2O$	(b)	CaB <sub>4</sub> H <sub>7</sub> .4H <sub>2</sub> O		
	(c)	CaNaB <sub>5</sub> O <sub>9</sub> .8H <sub>2</sub> O	(d)	$Na_2B_4O_7.4H_2O$	pakcity.org	b_
29.		lead chromate is formed whe			parcity.org	80
	(a)	Dilute alkali hydroxide	(b)	Dilute acid		
20	(C)	Strong alkali hydroxide	(d)	Strong acid	(LUD 2011)	
30.	, ,	th of the following is used in m Water glass		A X	(LHR 2011)	
	(a) (c)	Kaolin	(b)	Borax glass Asbestos		
31.		oboric acid when heated to re		Asbestos	(GRW 2011)	
<b>J</b> 1.	(a)	Boric anhydride	(b)	Pyroboric acid	(GKW 2011)	
	(c)	Metaboric acid	(d)	Tetraboric acid		
32.		, Te and Po are called:			(FSD 2009)	
	(a)	Coinage metals	(b)	Alkali metals	,	
	(c)	Chalcogens	(d)	Halogens		
33.	Alum	inum oxide is:		, RWP, MTN 15, LHR 1	14,15, FSD 09,13	3)
	(a)	Acidic oxide	(b)	Basic oxide		
	(c)	Amphoteric oxide	(d)	None of these		
34.	Borio	acid reacts with caustic soda			(FSD 2010)	E18-0
	(a)	NaBO <sub>2</sub> (b) NaH <sub>2</sub> Bo	O <sub>3</sub> (c)	$Na_2B_4O_7$	(d) Na₃	
35.		nary glass is:	\$1. So	- T-1	(FSD 2011)	
	(a)	Potassium silicate	(b)	Calcium silicate		
26	(c)	Sodium silicate	(d)	Calcium and sodium		
36.		compound which forms the be	2/21   3/		IR 08, SGD 09)	
	(a)	Metal oxide	(q)	Metal boride		
	(c)	Metal borate	(d)	Metal metaborate		

37.	The me	tal whi	ch doe	es not e	aive bo	orax be	ead tes	t is:				(SGD 2	010)
<b>J</b> 2.		Cu	cii doc	(b)	Cr	JIUX D	(c)	Ni Ni				(d)	Al
38.	The aqu	ueous s	olutio	**************************************	rax is:						(SWL	. 15, SGI	11)
	(a)	Acidic					(b)	Ne	eutral				
1.2	` '	Basic		2 2 2			(d)	Co	orrosive				
39.	Which		nore s		LIBO		(-)	1 1	п О			(RWP 20	and the second
40.	(a) <b>Nitric a</b>	H <sub>3</sub> BO <sub>3</sub>	ho tra	(b)	HBO:		(c)		2B2O2			(d) <b>(RWP</b> -	H <sub>6</sub> B <sub>3</sub> O <sub>9</sub>
40.		Al	DE II a	iiispoi	teu III (	a COIIC	(b)	raue u Zr	-0			(KAA b	-09)
		Cu					(d)		ne of t	hese			
41.	Valence		electro	nic co	nfigura	ation o					IIA is:	(RWP	2011)
	(a)	ns¹,np²		(b)	ns²,n		(c)		s <sup>0</sup> ,np <sup>3</sup>			(d)	ns <sup>2</sup> ,np <sup>1</sup>
42.	The chi		of alun	ninum	is:		ar s	201		2			
		NaAlF <sub>3</sub>					(q)		2O3.2H2				
43.	(c) <b>The onl</b>	Al <sub>2</sub> O <sub>3</sub>	alin ar	OUD TT	ΓΛ ic:		(d)	Al	<sub>2</sub> O <sub>3</sub> .H <sub>2</sub> C	,			
7J.		ı <b>y</b> nonmet Ar	amı yı	(b)	Ga		(c)	В				(d)	In
44.	C+SnO		Sn +			reaction			ts as a			(~)	
		Reducin					(b)		xidizing				
	(c)	Dehydra	ating ag	jent			(d)		one of t	hese			
45.	The che			a of cla	ay is:		<i>a</i> . s	<b>*</b>		\			
	5 (	Al <sub>2</sub> O <sub>3</sub> SiF	4				(b)		a <sub>3</sub> AlF <sub>6</sub>	) 2. 21. 6	_		
46.	(c) <b>Which</b> i	Al <sub>2</sub> O <sub>3</sub>	ly occi	ırrina	cubeta	nco ic	(d)	All	203.2Si	J2.2H2C	J		
70.		Haemat		(b)	Lime		(c)	(8)C	yolite			(d)	Quartz
47.	The hig			3.5			~ 1 1		, 0.1.00			(4)	Q
	in-	Silicone					(b)	222	uartz				
		Water g				20)	(d)		treous	silica			
48.	Which (		the giv		\ \ <	111	E 15		_			<i>i</i> 15	
40	<b>.</b>	MgO		(b)	Na <sub>2</sub> O		(c)		) <sub>2</sub>			(d)	ZnO
49.	When H (a)	13 <b>003</b> F Na <sub>3</sub> BO <sub>3</sub>	eacts v	(b)	Na <sub>2</sub> B	4000	(c)		aH <sub>2</sub> BO <sub>3</sub>			( <b>DGK 2</b> (d)	NaBO <sub>2</sub>
50.	Which		t form	1 2 4 7					U1 12DO3			(4)	Nubuz
		Carbon		· //			(b)	200	licone				
	(c)	Aluminu	ım			nur	(d)	Be	eryllium				
					le le			A P					
						NOW	CD VC	v 🧖					
					A	MQM	ER KE						
		1.	2.	3.	4.	5.	6.	7./	8.	9.	10		
		d	С	a	C	d	С	d	d	С	d	]	
		11	12	13	14	15	16	17	18	19	20	]	
		С	С	С	d	b	С	a	d	С	b	]	
		21	22	23	24	25	26	27	28	29	30	]	
		b	С	b	а	b	d	b	a	С	d	]	
		31	32	33	34	35	36	37	38	39	40	_	
		а	С	С	С	d	d	d	С	a	a	]	
		41	42	43	44	45	46	47	48	49	50	]	
		d	b	С	a	d	d	d	d	b	С		

1.	2.	3.	4.	5.	6.	7.	8.	9.	10
d	С	а	C	d	C	d	d	С	d
11	12	13	14	15	16	17	18	19	20
С	С	С	d	b	С	а	d	С	b
21	22	23	24	25	26	27	28	29	30
b	С	b	a	b	d	b	a	С	d
31	32	33	34	35	36	37	38	39	40
а	С	С	С	d	d	d	С	a	а
41	42	43	44	45	46	47	48	49	50
d	b	С	а	d	d	d	d	b	С



			CHAPT	ΓER 4				
			(		is a second seco	<i>*</i>		
	GR	OUP '	V A AND	IVA	ELEMENT	S	pakcity.o	rg 👺
1.	The % by volume o	f N <sub>2</sub> in a	ir is:					
_	(a) 98	(b)	88	(c)	78		(d)	68
2.	Which one of the fo	(b)	SO <sub>3</sub>	eric oxia (c)	e: CO <sub>2</sub>		(d)	$Al_2O_3$
3.	In Pyrite burner, th	0.5	oduced is:	(0)	COZ		(4)	711203
_	(a) SO <sub>3</sub>	(b)	SO <sub>2</sub>	(c)	CO <sub>2</sub>		(d)	NO
4.	In which substance		n is not prese Protein		Salt notor		(d)	Galena
5.	(a) Urea Oxidation of NO in a	(b) air prod		(c)	Salt peter		(d)	Jaicha
	(a) $N_2O$	(b)	$N_2O_3$	(c)	NO <sub>2</sub>	(d)	NO	
6.	Which of the follow					fuon in		2008)
	<ul><li>(a) It means light</li><li>(c) Bone ash is it</li></ul>			(b) (d)	Does not exist has no allotrop			
7.	Molecular formula							2009)
0	(a) P <sub>4</sub>	(b)	P	(c)	P <sub>3</sub>		(d)	P <sub>2</sub>
8.	<b>At 18 °C the specific</b> (a) 1.891	gravity (b)	2.101	(c)	1.834		(LHR (d)	<b>2010)</b> 1.740
9.	Which of the follow			\ \ \			(u)	1.7 10
	(a) Bi	(b)	Sb	(c)	As		(d)	Р
10.	The anhydride of ni	0.00	$\sim$	5 (0)	N O		(4)	NO
11.	(a) N <sub>2</sub> O <sub>4</sub> Nitrous acid is a:	(b)	N <sub>2</sub> O <sub>3</sub>	) (c)	$N_2O_5$		(d)	NO
	(a) Reducing age	nt (b)	Oxidizing age	ent (c)	Both a and b		(d)	None of
	these		VV.					
12.	Which of the follow (a) P <sub>2</sub> O <sub>3</sub>	ing is a	P <sub>2</sub> O <sub>5</sub>	copic pov	vder: P <sub>2</sub> O <sub>2</sub>		(d)	P <sub>2</sub> O <sub>4</sub>
13.	Aqua Regia is:	(100)	1205	(c)	1 202		(u)	1 204
	(a) 3 volumes of		olume of HNO <sub>2</sub>	(b)	3 Volumes of			
1.4	/Or 200		Volume of HCl		3 Volumes of	HCl+ 1V	olume o	of H <sub>2</sub> SO <sub>4</sub>
14.	Which one of the fo (a) N	(b)	P	ur free ii (c)	As		(d)	Sb
<b>15.</b>	Red phosphorous ca						1-7	
			catalyst in vacu	um at 25	0 °C			
	<ul><li>(b) Distilling it in</li><li>(c) Dissolving it i</li></ul>			y.org				
			the liquid into v	water				
16.	Phosphorus pentox		sed as:	<i>m</i> >		¥		
	(a) A cleansing a			(b)	A reducing ag			
17.	(c) A bleaching a The structure of wh	22°	sphorus is:	(d)	A dehydrating	agent		
	(a) Square plana	-	P.1.0. 4.0	(b)	Pyramidal			
	(c) Tetrahedral	_		(d)	Trigonal plane	er		
18.	Which of the follow	3.33 53	sphorus is mo	24D F 62U		oruc		
	<ul><li>(a) Red phosphor</li><li>(c) Scarlet phosp</li></ul>			(b) (d)	White phosphore Violet phosphore			
	(-)			(~)	c priosprik			

19.	Ortho	phosphoric a	cid is:						
	(a)	Monobasic			(b)	Dibasic			
	(c)	Tribasic			(d)	Tetrabasic			
20.	VPC - 0%	acts as an/a:			71-3	0 . 1			
	(a)	Acid	ont		(b)	Oxidizing ag	5.5		
21.	(c)	Reducing ago is heated with		n det	(d)	All the three	₫		
21.	(a)	Hypophosph			(b)	Phosphorou	ıs acid		
	(c)	Hypophosph			(d)	Orthophosp			
22.	(A) 17/4	e is not:	0110 0010		(4)	от иноргиоср	none dela		
	(a)	An allotrope			(b)	A powerful	oxidizing age	ent	
	(c)	Paramagneti	c species		(d)	A bent mole			
23.	Oleur	n is:							
	(a)	$H_2SO_3$	(b)	H <sub>2</sub> SO <sub>4</sub>	(c)	$H_2S_2O_7$		(d)	None
24.	SO₃ is	not directly							
	(a)	The reaction	Visit in the same and		100 m	The reaction			
25	(c)	The reaction	<b>.</b> ,		(d)		uble in water		
25.		oup VA the mo		onegative				(4)	Ch
26.	(a) Cinna	N bar is:	(b)	r	(c)	As		(d)	Sb
20.	(a)	HgS	(b)	ZnS	(c)	PbS		(d)	FeS <sub>2</sub>
27.		element whos			10.00		nt in earth		1 CO2
and s		(LHR 2011)	CO. gu.			0		0.000	
	(a)	Li	(b)	N	(c)	Na		(d)	0
28.	Gold	dissolves in a	qua regia	a to form:	463	5)	(BWP (	09, FSD	10)
	(a)	AuCl₃			(4p)	AuI <sub>3</sub>			
	(c)	AuI <sub>2</sub>			(d)	Au <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>			
29.	Arsei	nic oxides are	removed	l during m	anufacture	of H <sub>2</sub> SO <sub>4</sub> by	passing th	rough:	
	(2)	(SGD 2010) Ferric hydrox	vido.	2/0/2	(b)	Sodium hyd	Irovida		
	(a) (c)	Calcium hydi	9	My	(d)	Potassium h			
30.		compound N <sub>2</sub> (	1 1		(4)	1 otassiaiii i	iyai oxiac	(RWP 2	2011)
	(a)	Cancer	(b)		sickness (c)	Hysterical la	aughter	(d)	Tumor
31.		4 has great at	\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	A CONTRACT OF THE PARTY OF THE			<b>J</b>	(MTN 2	
	(a)	it decompose	es the acid	E PERMI	(b)	it hydrolyse	s the acid	,. <del>-</del> .	
	(c)	acid decomp	oses water		(d)	acid forms l	nydrate with	water	
32.	Atom	nic number of	Te is:					(MTN 2	2009)
	(a)	52	(b)	60	(c)	65		(d)	80
33.	22	c acid, Sulphu	ric acid a	nd Caustic	soda can b	e transporte	ed in a cont		2.13
	of:	Aluminum	/h)	Connor	(6)	Zina	(4)	(MTN 2	2009)
34.	(a) Whic	Aluminum  th of the follow	(b) wing elen	Copper	(C)	Zinc	(d) ena of allot	Teflon	
JT.		2009,10)	wing eleli	ilents dues	HOL SHOW	me hueuouu	cha or anot	. opy:	
	(a)	As	(b)	N	(c)	Sb	(d)	all of th	hese
35.	//X 20	gas emitted w	X 2		X 2			(MTN 2	
	(a)	$N_2O$	(b)	NO	(c)	$NO_2$		(d)	$N_2O_5$
36.	Whe	n sugar is trea	ated with	conc. H <sub>2</sub> S	O <sub>4</sub> the suga	r becomes b	olack due to	<b>)</b> :	
	(LHR	13, BWP 08)			F2 - 2				
	(a)	Oxidation			(b)	Reduction			
	(c)	Dehydration			(d)	Combustion			
37.	The c	catalyst used	ın manufa	acturing of	r H <sub>2</sub> SO <sub>4</sub> by (	₹0		CDW 4	21
	(a)	$V_2O_5$	(b)	Fe <sub>2</sub> O <sub>3</sub>	(c)	Ni Ni	SK, BWP 10,	(d)	2) Pt
	(a)	V 2 <b>O</b> 5	(0)	1 6203	(0)	1 11		(u)	1-0

(a) Zn (b) Cu (c) Mg (d) Sn  9. Which of the following statement is incorrect? (a) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (b) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (c) H <sub>2</sub> SO <sub>4</sub> acts as a strong reducing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (DGK 2009) (a) 80% P (b) 90% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> 1. NO <sub>2</sub> is called: (DGK 2011) (a) Nitrogen peroxide (d) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  2. Which of the following metal reacts with HNO <sub>3</sub> : (BWP 2012) (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (LHR 2012) (a) 49 °C (b) 45 °C (c) 41 °C (d) 50 °C  4. Which of the following gas is evolved when copper reacts with dil. HNO <sub>3</sub> ? (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO <sub>2</sub> (d) none of these  5. Which of the following gives brown ring with FeSO <sub>4</sub> ? (LHR 2014) (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> 6. Which metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014) (a) Pt (b) Co (c) Sn (d) Mn	(a) Zn (b) Cu (c) Mg (d) Sn (DGK 2009)  Which of the following statement is incorrect;  (a) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (c) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (e) H <sub>2</sub> SO <sub>4</sub> acts as a strong reducing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (DGK 2009)  80. Bone ash contains: (DGK 2009)  (a) 80% P (b) 90% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (1) NO <sub>2</sub> is called: (DGK 2011)  (a) Nitrogen peroxide (d) Nitric oxide (c) Nitric oxide (d) Nitric anhydride  (b) Iridium (d) Nitric anhydride  (c) Platinum (d) Magnesiu  (d) Magnesiu  (o) Platinum (d) Magnesiu  (o) Platinum (d) Magnesiu  (o) Platinum (d) Magnesiu  (o) NO <sub>2</sub> (d) 50°C  (o) NO <sub>2</sub> (d) none of these  (o) NO <sub>2</sub> (d) none of these  (o) NO <sub>2</sub> (d) none of these  (o) NO <sub>2</sub> (d) NO <sub>3</sub> (GGD 2014)  (a) NO <sub>2</sub> (b) NO (c) NO <sub>3</sub> (d) NO <sub>3</sub> (b) NO (c) NO <sub>2</sub> (d) NO <sub>3</sub> (c) NO <sub>2</sub> (d) NO <sub>3</sub> (d) NO <sub>3</sub> (GGD 2014)  (a) Yellow phosphorous (d) white phosphorous red phosphorous red phosphorous red phosphorous  (c) b b a a a d c b c d d a c a c c d d d a c b c d d a a a b c a a c a c a c a c d d d a a a b c c a a b c a c c a c a c a c a	(a) Zn (b) Cu (c) Mg (d) Sn  Which of the following statement is incorrect;  (a) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (c) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (e) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (DGK 2009)  (a) 80% P (b) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (b) NO <sub>2</sub> is called: (DGK 2011)  (a) Nitrogen peroxide (d) Nitric anhydride  (c) Nitric oxide (d) Nitric anhydride  (d) Nitric anhydride  (e) Nitric oxide (d) Nitric anhydride  (f) Nitric oxide (h) Nitric anhydride  (g) Tirdium (h) Tirdium (h) Tirdium (h) Magnesiu  (g) Tirdium (h) Tirdium (h) Magnesiu  (g) Tirdium (h) NO <sub>2</sub> (h) NO  (g) Patinum (h) 45°° (c) 41°° (d) 50°°  (g) NO <sub>2</sub> (h) NO  (g) NO <sub>2</sub> (h) NO  (g) NO <sub>2</sub> (h) NO  (h) NO  (h) NO  (c) NO <sub>2</sub> (d) none of these  (h) NO  (d) none of these  (h) NO  (d) NO <sub>3</sub> (d) NO <sub>3</sub> (e) Pt (h) Co (c) No <sub>2</sub> (d) NO <sub>3</sub> (a) Yellow phosphorous (b) NO  (b) NO  (c) NO <sub>2</sub> (d) NO <sub>3</sub> (d) NO <sub>3</sub> (e) Pt (b) Co (c) Sn (d) Mn  (LHR 2014)  (a) Yellow phosphorous (b) NO  (b) NO  (c) No <sub>3</sub> (d) NO  (d) NO  (d) NO  (d) NO  (d) NO  (e) Patinum (h) Tirdium (h) NO  (f) NO <sub>2</sub> (h) NO  (g) NO <sub>2</sub> (h) NO	38.	The giv	on olon	nent a	ives N	nac v	with di	I HNO				(RWD 2	011)	
(a) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (c) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (e) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (e) Bone ash contains: (DGK 2009)  (a) 80% P (b) 90% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> 1. NO <sub>2</sub> is called: (DGK 2011)  (b) Nitrous oxide (d) Nitric oxide (e) Nitric oxide (f) Nitric	(a) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (c) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (DoK 2009)  (a) 80% P (b) 90% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (b) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (c) Nitrogen peroxide (d) Nitrous oxide (d) Nitric oxide (e) Nitric oxide (d) Nitric anhydride (e) Nitric oxide (e) Nitric oxide (f) Nitric oxide (g) Nitric oxide (g) Nitric oxide (h) Nitric oxide (g) Nitric oxide (h) Ni	(a) H <sub>2</sub> SO <sub>4</sub> acts as a strong oxidizing agent (c) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a dehydrating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> as supported to the color as sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> as sulphonating agent (D) H <sub>2</sub> SO <sub>4</sub> (D) H <sub>2</sub> S	<b>J</b> 0.			nent g		~	with a	7.5		g		(BWP 2		Sn
(c) H <sub>2</sub> SO <sub>4</sub> acts as a strong reducing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (a) 80% P (b) 90% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (DGK 2011)  1. NO <sub>2</sub> is called: (DGK 2011)  (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  2. Which of the following metal reacts with HNO <sub>3</sub> : (BWP 2012)  (a) Titanium (b) Iridium (c) Platinum  (b) Iridium (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (LHR 2012)  (a) 49°C (b) 45°C (c) 41°C (d) 50°C  4. Which of the following gas is evolved when copper reacts with dil. HNO <sub>3</sub> ? (RWP 2009)  (a) N <sub>2</sub> O (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (b) NO (c) NO <sub>2</sub> (d) NO <sub>3</sub> 6. Which of the following sives brown ring with FeSO <sub>4</sub> ? (LHR 2014)  (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (c) Hhich metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014)  (a) Pt (b) Co (c) N <sub>3</sub> O <sub>3</sub> (d) Mn (LHR 2008)  (b) NO (c) N <sub>3</sub> O <sub>3</sub> (d) Mn (LHR 2008)  (c) Dlack phosphorous (d) white phosphorous red phosphorous red phosphorous	(c) H <sub>2</sub> SO <sub>4</sub> acts as a strong reducing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (a) 80% P (b) 90% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (DGK 2011)  11. NO <sub>2</sub> is called: (DGK 2011)  12. Which of the following metal reacts with HNO <sub>3</sub> : (BWP 2012)  13. Ortho-phosphoric acid has melting point: (d) Magnesiu  14. Which of the following gas is evolved when copper reacts with dil. HNO <sub>3</sub> ? (RWP 2009)  15. Which of the following gives brown ring with FeSO <sub>4</sub> ? (LHR 2014)  16. Which metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014)  17. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) White phosphorous (c) black phosphorous (d) Fe D C (d) MSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10  1. 1 12 13 14 15 16 17 18 19 20  1. 2 2 23 24 25 26 27 28 29 30  1. 2 41 42 43 44 45 46 47	(c) H <sub>2</sub> SO <sub>4</sub> acts as a strong reducing agent (d) H <sub>2</sub> SO <sub>4</sub> acts as a sulphonating agent (DGK 2009)  (a) 80% P (b) 90% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> 41. NO <sub>2</sub> is called: (DGK 2011)  (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  42. Which of the following metal reacts with HNO <sub>3</sub> : (BWP 2012)  (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  43. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  44. Which of the following gas is evolved when copper reacts with dil. HNO <sub>3</sub> ?  (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) N <sub>2</sub> O <sub>2</sub> (d) NO <sub>3</sub> 45. Which of the following gives brown ring with FeSO <sub>4</sub> ? (LHR 2014) (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> 46. Which metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014) (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) Mn  47. Which of the following shows phosphorescence? (d) Fellow phosphorous (e) black phosphorous (f) white phosphorous red phosphorous red phosphorous (d) Tall 1 12 13 14 15 16 17 18 19 20  (c) b b a a a d c b c d a c a c a c a c a c a c b c b c d d a c c a a c a c a c a c a c a c a	39.				122			etc Variable	720 TO VO.			-		
0. Bone ash contains: (a) 80% P (c) 80% Ca₃(PO₁)₂ (d) 90% Ca₃(PO₁)₂  1. NO₂ is called: (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  2. Which of the following metal reacts with HNO₃: (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (a) 49 ℃ (b) 45 ℃ (c) 41 ℃ (d) 50 ℃  4. Which of the following gas is evolved when copper reacts with dil. HNO₃? (RWP 2009) (a) N₂O (b) NO (c) NO₂ (d) none of these  5. Which of the following gives brown ring with FeSO₄? (a) NO₂ (b) NO (c) NO₂ (d) none of these  5. Which metal is rendered passive by HNO₃? (a) Pt (b) Co (c) Sn (d) NO₃ (e) Sn (d) Mn (LHR 2008) (h) Yellow phosphorous (c) black phosphorous (d) red phosphorous (e) black phosphorous (f) D (f) RS Sn (f) RS	(a) 80% P (b) 90% P (c) 80% Cas(PO4)2 (d) 90% Cas(PO4)2 (d) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride (c) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Platinum (d) Magnesiu (c) Platinum (d) Magnesiu (LHR 2012) (a) 49° (b) 45° (c) 41° (d) 50° (c) Platinum (d) Magnesiu (LHR 2012) (d) 49° (d) 45° (c) 41° (d) 50° (e) NO2 (d) none of these (RWP 2009) (a) N2O (b) NO (c) NO2 (d) none of these (LHR 2014) (a) NO2 (b) NO (c) NO3° (d) NO3 (d) NO3 (d) NO3 (d) NO4 (d) NO4 (d) NO5 (e) No4 (d) NO5 (e) No4 (d) NO5 (e) No5 (d)	40. Bone ash contains: (a) 80% P (b) 90% P (c) 80% Cas(PO4)2 (d) 90% Cas(PO4)2 41. NO2 is called: (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide 42. Which of the following metal reacts with HNO3: (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  43. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  44. Which of the following gas is evolved when copper reacts with dil. HNO3? (RWP 2009) (a) N2O (b) NO (c) NO2 (d) none of these  45. Which of the following gives brown ring with FeSO4? (a) NO2 (b) NO (c) NO3 (d) NO3 (d) NO3 (d) NO4 (a) Pt (b) CO (c) Sh (d) Mn (LHR 2014) (a) Pt (b) CO (c) Sh (d) Mn (LHR 2008) (b) NO (c) Sh (d) Mn (LHR 2008) (c) black phosphorous (d) red phosphorous (e) black phosphorous (f) Titalium (DGK 2010) (BWP 2011) (BWP 2012) (BWP 2012) (ILHR 2014) (d) NO3 (d) NO								-				: <b>-</b> 1		
(a) 80% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (1. NO <sub>2</sub> is called: (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  2. Which of the following metal reacts with HNO <sub>3</sub> : (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  4. Which of the following gas is evolved when copper reacts with dil. HNO <sub>3</sub> ? (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO <sub>2</sub> (d) none of these  5. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) NO <sub>2</sub> (d) NO <sub>3</sub> (d) NO <sub>3</sub> (SGD 2014) (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) Mn (LHR 2008) (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) Which of the following shows phosphorecee? (b) white phosphorous (c) black phosphorous (d) T R R R P R P R P R P R P R P R P R P R	(a) 80% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (DGK 2011)  NO <sub>2</sub> is called: (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  Noth of the following metal reacts with HNO <sub>3</sub> : (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  Which of the following gas is evolved when copper reacts with dil. HNO <sub>3</sub> ? (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (b) NO (c) NO <sub>2</sub> (d) NO <sub>3</sub> (c) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (d) NO <sub>3</sub> (a) Pt (b) Co (c) N <sub>2</sub> (d) Mn  Mich of the following shows phosphorous (e) black phosphorous (c) black phosphorous (d) Which of the following shows phosphorous red phosphorous (e) black phosphorous (f) Vellow phosphorous (g) Yellow phosphorous (h) Yello	(a) 80% P (c) 80% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (d) 90% Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> (DGK 2011)  41. NO <sub>2</sub> is called: (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  42. Which of the following metal reacts with HNO <sub>3</sub> : (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  43. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  44. Which of the following gas is evolved when copper reacts with dil. HNO <sub>3</sub> ? (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO <sub>2</sub> (d) none of these  45. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) NO <sub>2</sub> (d) none of these  46. Which metal is rendered passive by HNO <sub>3</sub> ? (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (c) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) Mn (d) Mn (HR 2008) (e) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (f) Which of the following shows phosphoreace? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) Which of the following and a d c d c d d d c d d d c d d d c d d d c d d d c d d d c d d d c d d d c d d d d c d	40	- 1.6			Strong	reduci	ng agei	11 (0	i) п	2 <b>5</b> 04 aC	is as a	11-2	,	agent
1. NO2 is called: (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride  2. Which of the following metal reacts with HNO3: (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  4. Which of the following gas is evolved when copper reacts with dil. HNO3? (RWP 2009) (a) N2O (b) NO (c) NO2 (d) none of these  5. Which of the following gives brown ring with FeSO4? (a) NO2 (b) NO (c) NO2 (d) none of these  5. Which of the following sives brown ring with FeSO4? (a) Pt (b) CO (c) N2O3 (d) NO3 (d) NO3 (d) NO3 (d) Mn (LHR 2014) (a) NO4 (a) Pt (b) CO (c) N2O3 (d) NO3 (d) NO3 (d) Mn (LHR 2016) (d) Mn (LHR 2016) (e) NO4 (f) NO5 (f) NO6 (g) NO7 (hich of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) red phosphorous (e) black phosphorous (f) NSWER KEY  ANSWER KEY  ANSWER KEY	11. NO2 is called: (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Nitric oxide (for Nitric oxide (g) Nitric oxide (h) Nitric anhydride (h) Nagesiu (LHR 2012) (d) Magnesiu (lHR 2014) (d) NO3 (p) NO2 (d) NO3 (d)	41. NO2 is called:  (a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Nitric anhydride (fill Magnesiu (fill Magn		E 81						(	b) 9	0% P			,	
(a) Nitrogen peroxide (c) Nitric oxide (d) Nitric oxide (d) Nitric oxide (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Platinum (d) Magnesiu (d) Magnesiu (e) Platinum (d) Magnesiu (d) Magnesiu (d) 49°C (b) 45°C (c) 41°C (d) 50°C (d) 70°C (d) NO2 (e) NO2 (d) NO3 (d) NO3 (e) NO2 (d) NO3 (e) NO2 (d) NO3 (e) NO3 (d) NO3	(a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (e) Nitric anhydride (d) Nitric anhydride (e) Nitric anhydride (e	(a) Nitrogen peroxide (b) Nitrous oxide (c) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (e) Nitric oxide (d) Nitric anhydride (d) Nitric anhydride (d) Nitric anhydride (e) Nitric anhydride (d) Nitric anhydride (e) Platinum (d) Magnesiu (d) Magnesiu (d) Magnesiu (d) Magnesiu (d) 49°C (b) 45°C (c) 41°C (d) 50°C (d) 49°C (b) 45°C (c) 41°C (d) 50°C (e) No (	·= :=	Se 2.5		3 <b>(PO</b> 4 <b>)</b> 2				(	d) 9	0% Ca₃	(PO <sub>4</sub> ) <sub>2</sub>		works are to the	
(c) Nitric oxide (d) Nitric anhydride  Which of the following metal reacts with HNOs: (BWP 2012)  (a) Titanium (b) Iridium (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  4. Which of the following gas is evolved when copper reacts with dil. HNO3? (RWP 2009)  (a) N2O (b) NO (c) NO2 (d) none of these  5. Which of the following gives brown ring with FeSO4? (LHR 2014)  (a) NO2 (b) NO (c) N2O3 (d) NO3  6. Which metal is rendered passive by HNO3? (SGD 2014)  (a) Pt (b) Co (c) Sn (d) Mn  7. Which of the following shows phosphorescence? (b) white phosphorous (c) black phosphorous  (c) black phosphorous (d) red phosphorous  (d) Mn  (LHR 2008)  ANSWER KEY  ANSWER KEY	(c) Nitric oxide (d) Nitric anhydride  Which of the following metal reacts with HNOs: (BWP 2012)  (a) Titanium (b) Iridium (d) Magnesiu  B3. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  Which of the following gas is evolved when copper reacts with dil. HNO3? (RWP 2009)  (a) N <sub>2</sub> O (b) NO (c) NO2 (d) none of these  Which of the following gives brown ring with FeSO4? (LHR 2014)  (a) NO2 (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO3  B6. Which metal is rendered passive by HNO3? (SGD 2014)  (a) Pt (b) Co (c) Sn (d) Mn  Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous  (c) black phosphorous (d) red phosphorous  Pakcity.org ANSWER KEY   ANSWER KEY	(c) Nitric oxide (d) Nitric anhydride  Which of the following metal reacts with HNO3: (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  43. Ortho-phosphoric acid has melting point: (LHR 2012)  (a) 49°C (b) 45°C (c) 41°C (d) 50°C  44. Which of the following gas is evolved when copper reacts with dil. HNO3? (RWP 2009)  (a) N <sub>2</sub> O (d) none of these  Which of the following gives brown ring with FeSO4? (LHR 2014)  (a) NO2 (d) none of these  Which metal is rendered passive by HNO3? (SGD 2014)  (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) NO  47. Which of the following shows phosphorescence? (a) Yellow phosphorous  (c) black phosphorous  (d) whith phosphorous  (e) black phosphorous  (f) Which of the following shows phosphorescence? (a) Yellow phosphorous  (b) Which of the following shows phosphorescence? (b) C d  47. Which of the following shows phosphorescence? (b) C d  48. Which of the following shows phosphorescence? (d) Mn  (LHR 2008)  (a) Yellow phosphorous  (b) Whith phosphorous  (c) black phosphorous  (d) Titalium  (LHR 2014)  (a) NO  (b) NO  (c) N <sub>2</sub> O <sub>3</sub> (d) NO  (d) NO  (SGD 2014)  (a) Yellow phosphorous  (b) Whith phosphorous  (c) black phosphorous  (d) Titalium  (LHR 2014)  (a) NO  (b) NO  (c) NO  (c) NO  (d) NO  (d) NO  (SGD 2014)  (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) NO  (d) Mn  (LHR 2008)  (LHR 2008)  (LHR 2014)  (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) NO  (c) NO  (d) NO  (d) NO  (SGD 2014)  (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) NO  (ED NO  (LHR 2008)  (LHR 2014)  (d) NO  (ED NO  (LHR 2008)  (D) NO	41.			norov	ido				(h) N	itrous o	vido	(DGK 2	2011)	
2. Which of the following metal reacts with HNO3:  (a) Titanium (b) Iridium (c) Platinum (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (a) 49 °C (b) 45 °C (c) 41 °C (d) 50 °C  4. Which of the following gas is evolved when copper reacts with dil. HNO3?  (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO2 (d) none of these  5. Which of the following gives brown ring with FeSO4? (a) NO2 (b) NO (c) N <sub>2</sub> O (d) NO3 (d) NO3 (d) NO3 (d) NO3 (e) NO4 (a) Pt (b) Co (c) N <sub>2</sub> O <sub>3</sub> (d) NO3 (d) NO3 (d) NO3 (sGD 2014) (a) Pt (a) Pt (b) Co (c) Sn (d) Mn (LHR 2008) (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) red phosphorous (e) Dlack phosphorous (f) Pakcity.org  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 (c) D D C (d) NO3	Which of the following metal reacts with HNOs:   (a)   Titanium   (b)   Irridium   (c)   Platinum   (d)   Magnesiu	42. Which of the following metal reacts with HNOs:  (a) Titanium (b) Irridium (c) Platinum (d) Magnesiu  43. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  44. Which of the following gas is evolved when copper reacts with dil. HNO3?  (RWP 2009) (a) N2O (b) NO (c) NO2 (d) none of these  45. Which of the following gives brown ring with FeSO4? (a) NO2 (b) NO (c) N2O3 (d) NO3  46. Which metal is rendered passive by HNO3? (a) Pt (b) Co (c) Sn (d) Mn  47. Which of the following shows phosphorescence? (c) black phosphorous (d) MSWER KEY   ANSWER KEY				-	ue				` '					
(c) Platinum (d) Magnesiu  3. Ortho-phosphoric acid has melting point: (LHR 2012) (a) 49°C (b) 45°C (c) 41°C (d) 50°C  4. Which of the following gas is evolved when copper reacts with dil. HNO3? (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO2 (d) none of these  5. Which of the following gives brown ring with FeSO4? (LHR 2014) (a) NO2 (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> 6. Which metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014) (a) Pt (b) Co (c) Sn (d) Mn  7. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (c) black phosphorous  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 C d b d c d a c a c 11 12 13 14 15 16 17 18 19 20 C b b a a a d c b c d 21 22 23 24 25 26 27 28 29 30 d d c c a a b a a c 31 32 33 34 35 36 37 38 39 40 d a a a b c a a b c c 41 42 43 44 45 46 47	(c) Platinum (d) Magnesiu (d) Soft (d) Soft (d) Soft (EHR 2012) (d) NO3 (d) Magnesiu (d) Soft (EHR 2014) (d) NO3 (SGD 2014) (d) NO3 (SGD 2014) (d) Magnesiu (d) Soft (EHR 2015) (d) NO3 (SGD 2014) (d) Magnesiu (d) Soft (EHR 2016) (EHR 2017) (G) No3 (G)	43. Ortho-phosphoric acid has melting point: (a) 49°C (b) 45°C (c) 41°C (d) 50°C  44. Which of the following gas is evolved when copper reacts with dil. HNO3? (RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO2 (d) none of these  45. Which of the following gives brown ring with FeSO4? (a) NO2 (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO3 (46. Which metal is rendered passive by HNO3? (a) Pt (b) Co (c) Sn (d) Mn  47. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) Which phosphorous (c) black phosphorous (d) Which phosphorous (e) Pakcity.org  ANSWER KEY	42.		of the f	ollowii	ng me	tal rea	cts wit		·		,	(BWP 2	2012)	
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(RWP 2009) (a) N <sub>2</sub> O (b) NO (b) NO (c) NO <sub>2</sub> (d) none of these  5. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (5GD 2014) (a) Pt (b) Co (c) Sn (d) Mn (LHR 2008)  7. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) red phosphorous (e) Pakcity.org  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 (a) C d b d C d a C a C (b) C d a C d a C a C (c) C d d d C d a C d a C d C (d) Mn (LHR 2008)  ANSWER KEY	(RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO <sub>2</sub> (d) none of these  35. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (d) NO <sub>3</sub> (d) NO <sub>3</sub> (e) NO <sub>6</sub> (e) N <sub>2</sub> O <sub>3</sub> (f) NO <sub>7</sub> (g) NO <sub>8</sub> (h) NO <sub>8</sub> (g) NO <sub>8</sub> (g) NO <sub>9</sub> (h) NO <sub>9</sub> (h) NO <sub>9</sub> (g) NO <sub>9</sub> (h) NO <sub>9</sub> (h) NO <sub>9</sub> (g) NO <sub>9</sub> (d) NO <sub>9</sub> (d) Mn (LHR 2008) (e) NO <sub>9</sub> (f) NO <sub>9</sub> (g) NO <sub>9</sub>	(RWP 2009) (a) N <sub>2</sub> O (b) NO (c) NO <sub>2</sub> (d) none of these  45. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> 46. Which metal is rendered passive by HNO <sub>3</sub> ? (a) Pt (b) Co (c) Sn (d) Mn (47. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) red phosphorous (e) Pakcity.org  ANSWER KEY   ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 (a) C (c) C (d)	10.	.g g.		oric ac	29 - 20	_		열 25	4	1℃		(-:::x <b>-</b>		50℃
(a) N <sub>2</sub> O (b) NO (d) none of these  5. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (6. Which metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014) (a) Pt (b) Co (c) Sn (d) Mn (LHR 2008)  7. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) red phosphorous (c) black phosphorous (d) Table Parallel	(a) N <sub>2</sub> O (b) NO (d) none of these  Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> Which metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014) (a) Pt (b) Co (c) Sn (d) Mn  Which of the following shows phosphorescence? (LHR 2014) (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) red phosphorous (c) black phosphorous (d) red phosphorous  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10	(a) N2O (b) NO (d) none of these  45. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> 46. Which metal is rendered passive by HNO <sub>3</sub> ? (SGD 2014) (a) Pt (b) Co (c) Sn (d) Mn  47. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous  (c) black phosphorous (d) red phosphorous  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10 (LHR 2008)  ANSWER KEY  ANSWER KEY  ANSWER KEY  2. 3. 4. 5. 6. 7. 8. 9. 10 (LHR 2008)  4. 4. 5. 6. 7. 8. 9. 10 (LHR 2008)  4. 5. 6. 7. 8. 9. 10 (LHR 2008)  4. 6. 6. 7. 8. 9. 10 (LHR 2008)  4. 6. 6. 7. 8. 9. 10 (LHR 2008)  4. 6. 7. 8. 9. 10 (LHR 2008)  4. 6. 6. 7. 8. 9. 10 (LHR 2008)  4. 6. 7. 8. 9. 10 (LHR 2008)	44.			ollowii	ng gas	is evo	lved w	hen co	pper	reacts	with d	il. HNC	) <sub>3</sub> ?	
(c) NO <sub>2</sub> (d) none of these  5. Which of the following gives brown ring with FeSO <sub>4</sub> ? (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> 6. Which metal is rendered passive by HNO <sub>3</sub> ? (a) Pt (b) Co (c) Sn (d) Mn  7. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (c) black phosphorous  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10  C d b d C d a C a C  11 12 13 14 15 16 17 18 19 20  C b b a a d c b C d  21 22 23 24 25 26 27 28 29 30  d d C C a a b C  31 32 33 34 35 36 37 38 39 40  d a a b C C a b C  41 42 43 44 45 46 47	(c) NO2	45. Which of the following gives brown ring with FeSO4? (a) NO2 (b) NO (c) N2O3 (d) NO3 (46. Which metal is rendered passive by HNO3? (a) Pt (b) Co (c) Sn (d) Mn (LHR 2014) (47. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (c) black phosphorous (d) red phosphorous  ANSWER KEY   ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 (c d b d c d a c a c 11 12 13 14 15 16 17 18 19 20 (c b b a a a d c b c d 21 22 23 24 25 26 27 28 29 30 (d d c c a a b c c a a b c c 31 32 33 34 35 36 37 38 39 40 (d a a a b c c a a b c c 41 42 43 44 45 46 47		2 .	ener see					(h)	N	0				
5. Which of the following gives brown ring with FeSO4?  (a) NO2 (b) NO (c) N2O3  (b) NO (c) N2O3  (c) N2O3  (d) NO3  (SGD 2014)  (d) Mn  (d) Mn  (LHR 2008)  7. Which of the following shows phosphorescence? (a) Yellow phosphorous (c) black phosphorous (c) black phosphorous  (d) red phosphorous (e) Pakcity.org  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10  c d b d c d a c a c  11 12 13 14 15 16 17 18 19 20  c b b a a d c b c d  21 22 23 24 25 26 27 28 29 30  d d c c a a b a a c  31 32 33 34 35 36 37 38 39 40  d a a b c c a b c  41 42 43 44 45 46 47	## Which of the following gives brown ring with FeSO <sub>4</sub> ?  (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (SGD 2014)  (a) Pt (b) Co (c) Sn (d) Mn  (b) Which of the following shows phosphorescence?  (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous  (b) Which of the following shows phosphorescence?  (a) Yellow phosphorous (d) white phosphorous red phosphorous  (c) black phosphorous (d) red phosphorous  (d) Mn  (LHR 2008)  (LHR 2014)  (d) NO <sub>3</sub> (SGD 2014)  (d) Mn  (LHR 2008)  (LHR 2008)  (a) Pt (b) Co (c) Sn (d) Mn  (LHR 2008)  (b) White phosphorous red phosphorous  (c) black phosphorous (d) red phosphorous  (d) Mn  (LHR 2008)  (a) Pt (b) Co (c) Sn (d) Mn  (LHR 2008)	45. Which of the following gives brown ring with FeSO <sub>4</sub> ?  (a) NO <sub>2</sub> (b) NO (c) N <sub>2</sub> O <sub>3</sub> (d) NO <sub>3</sub> (a) Pt (b) Co (c) Sn (d) Mn  (LHR 2014)  (a) Pt (b) Co (c) Sn (d) Mn  (LHR 2008)  (a) Yellow phosphorous (d) white phosphorous (c) black phosphorous  (b) Which of the following shows phosphorescence?  (c) Dack phosphorous (d) white phosphorous (e) pakeity.org  ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10  c d b d c d a c a c  11 12 13 14 15 16 17 18 19 20  c b b a a d c b c d  21 22 23 24 25 26 27 28 29 30  d d c c a a b a a c  31 32 33 34 35 36 37 38 39 40  d a a b c c a b c C  41 42 43 44 45 46 47								3 ×			hese			
6. Which metal is rendered passive by HNO <sub>3</sub> ? (a) Pt (b) Co (c) Sn (d) Mn (LHR 2008)  7. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) red phosphorous  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10  c d b d c d a c a c  11 12 13 14 15 16 17 18 19 20  c b b a a a d c b c d  21 22 23 24 25 26 27 28 29 30  d d c c a a b a a c  31 32 33 34 35 36 37 38 39 40  d a a b c c a b c C  41 42 43 44 45 46 47	Mhich metal is rendered passive by HNO <sub>3</sub> ?  (a) Pt (b) Co (c) Sh (d) Mn  (LHR 2008)  (b) Which of the following shows phosphorescence?  (a) Yellow phosphorous (d) white phosphorous red phosphorous  (c) black phosphorous  (d) FR 2008  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10  c d b d c d a c a c  11 12 13 14 15 16 17 18 19 20  c b b a a a d c b c d  21 22 23 24 25 26 27 28 29 30  d d c c a a b a a c  31 32 33 34 35 36 37 38 39 40  d a a b c c a b c C  41 42 43 44 45 46 47	46. Which metal is rendered passive by HNO <sub>3</sub> ? (a) Pt (b) Co (c) Sn (d) Mn (LHR 2008)  47. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous (c) black phosphorous (d) white phosphorous red phosphorous  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 c d b d c d a c a c 11 12 13 14 15 16 17 18 19 20 c b b a a d c b c d 21 22 23 24 25 26 27 28 29 30 d d c c a a b a a c 31 32 33 34 35 36 37 38 39 40 d a a b c c a b c C 41 42 43 44 45 46 47	45.	2 20		ollowii	ng give	es brov	wn ring	with		(0	\		(LHR	
(a) Pt (b) Co (c) Sin (d) Mn (LHR 2008)  7. Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous red phosphorous (c) black phosphorous  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10  c d b d c d a c a c  11 12 13 14 15 16 17 18 19 20  c b b a a a d c b C d  21 22 23 24 25 26 27 28 29 30  d d c c a a b a a c  31 32 33 34 35 36 37 38 39 40  d a a b c c a b c C  41 42 43 44 45 46 47	(a) Pt (b) Co (c) Sin (d) Mn (LHR 2008)  (a) Yellow phosphorous (c) black phosphorous (d) red phosphorous (c) Pakcity.org  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 (c) d b d c d a c a c (11 12 13 14 15 16 17 18 19 20 (c) c b b a a d c b C d (21 22 23 24 25 26 27 28 29 30 (d) d c c a a b a a c (31 32 33 34 35 36 37 38 39 40 (d) Mn (LHR 2008)	(a) Pt (b) Co (c) Sin (d) Min (LHR 2008)  (b) Which of the following shows phosphorescence? (a) Yellow phosphorous (b) white phosphorous red phosphorous  (c) black phosphorous  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10  c d b d c d a c a c  11 12 13 14 15 16 17 18 19 20  c b b a a a d c b C d  21 22 23 24 25 26 27 28 29 30  d d c c a a b a a c  31 32 33 34 35 36 37 38 39 40  d a a b c c a b c C  41 42 43 44 45 46 47	16	. /		ronde			by UN		N	2037	<u>)</u>			
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1. 2. 3. 4. 5. 6. 7. 8. 9. 10 c d b d c d a c a c 11 12 13 14 15 16 17 18 19 20 c b b a a a d c b c d 21 22 23 24 25 26 27 28 29 30 d d c c a a b a a c 31 32 33 34 35 36 37 38 39 40 d a a a b c c a b c 41 42 43 44 45 46 47	C	1.   2.   3.   4.   5.   6.   7.   8.   9.   10		(a)			\ /			\ /	MINI				( ) ( ) ( ) ( ) ( ) ( ) ( ) ( ) ( ) ( )	
1.   2.   3.   4.   5.   6.   7.   8.   9.   10	1.   2.   3.   4.   5.   6.   7.   8.   9.   10	1.   2.   3.   4.   5.   6.   7.   8.   9.   10	47.			ollowii	ng sho	ws pho	osphor	escen	(é?				(LHR	
c     d     b     d     c     d     a     c     a     c       11     12     13     14     15     16     17     18     19     20       c     b     b     a     a     d     c     b     c     d       21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     b     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     c       41     42     43     44     45     46     47	c     d     b     d     c     d     a     c     a     c       11     12     13     14     15     16     17     18     19     20       c     b     b     a     a     d     c     b     c     d       21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	c     d     b     d     c     d     a     c     a     c       11     12     13     14     15     16     17     18     19     20       c     b     b     a     a     d     c     b     c     d       21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     c       41     42     43     44     45     46     47	47.	Which (a)	<b>of the f</b> Yellow p	hospho	rous	ws pho		(db)	> w	-	1.5		(LHR	
11       12       13       14       15       16       17       18       19       20         c       b       b       a       a       d       c       b       c       d         21       22       23       24       25       26       27       28       29       30         d       d       c       c       a       a       b       a       a       c         31       32       33       34       35       36       37       38       39       40         d       a       a       b       c       c       a       b       c       c         41       42       43       44       45       46       47	11       12       13       14       15       16       17       18       19       20         c       b       b       a       a       d       c       b       c       d         21       22       23       24       25       26       27       28       29       30         d       d       c       c       a       a       a       c         31       32       33       34       35       36       37       38       39       40         d       a       a       b       c       c       a       b       c       c         41       42       43       44       45       46       47	11       12       13       14       15       16       17       18       19       20         c       b       b       a       a       d       c       b       C       d         21       22       23       24       25       26       27       28       29       30         d       d       c       c       a       a       a       c         31       32       33       34       35       36       37       38       39       40         d       a       a       b       c       c       a       b       c       C         41       42       43       44       45       46       47	47.	Which (a)	<b>of the f</b> Yellow p	hospho	rous	ws pho	19 E)	(d)	w re	-	1.5		(LHR	
c     b     b     a     a     d     c     b     c     d       21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     c       41     42     43     44     45     46     47	c     b     b     a     a     d     c     b     c     d       21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	c     b     b     a     a     d     c     b     c     d       21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     c       41     42     43     44     45     46     47	47.	Which (a)	<b>of the f</b> Yellow p	pakcit	rous	PA A	NSW	ER KE	w re	ed phos	phorous	5	(LHR	
21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     b     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	21       22       23       24       25       26       27       28       29       30         d       d       c       c       a       a       b       a       a       c         31       32       33       34       35       36       37       38       39       40         d       a       a       b       c       c       a       b       c       C         41       42       43       44       45       46       47	21     22     23     24     25     26     27     28     29     30       d     d     c     c     a     a     b     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	47.	Which (a)	<b>of the f</b> Yellow p	pakcity	orous ous ous 3.	4.	NSW	(d) ER KE 6.	ν re <b>7</b> .	ed phos	<b>9.</b>	10	(LHR	
d     d     c     c     a     a     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	d     d     c     c     a     a     b     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	d     d     c     c     a     a     b     a     c       31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	47.	Which (a)	of the formal property	pakcity 2.	orous ous ous 3.	4. d	NSW 5.	(b) (d) ER KE 6.	<b>Y 7.</b> a	8.	<b>9.</b> a	<b>10</b>	(LHR	
31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	31     32     33     34     35     36     37     38     39     40       d     a     a     b     c     c     a     b     c     C       41     42     43     44     45     46     47	47.	Which (a)	the formal property of	pakcity  2.  d 12	orous ous 3. b	4. d 14 a	5. c 15	6. d 16	7. a 17 c	8. C 18 b	9. a 19 C	<b>10</b> c <b>20</b> d	(LHR	
d a a b c c a b c C 41 42 43 44 45 46 47	d a a b c c a b c C 41 42 43 44 45 46 47	d a a b c c a b c C 41 42 43 44 45 46 47	47.	Which (a)	the formal property of	pakcity 2. d 12 b 22	orous ous 3. b	4. d 14 a 24	5. c 15 a 25	6. d 16 d 26	7. a 17 c 27	8. C 18 b	9. a 19 C 29	10 c 20 d 30	(LHR	
41 42 43 44 45 46 47	41 42 43 44 45 46 47	41 42 43 44 45 46 47	47.	Which (a)	the formal property of	pakcity 2. d 12 b 22 d	orous ous 3. b 13 b 23	4. d 14 a 24	5. c 15 a 25 a	6. d 16 d 26 a	7. a 17 c 27 b	8. c 18 b 28	9. a 19 C 29	10 c 20 d 30	(LHR	
b d c b b b	b d c b b b	b d c b b b	47.	Which (a)	the formal property of	pakcity b 2. d 12 b 22 d 32	orous ous 3. b 13 b 23 c	4. d 14 a 24 c 34	5. c 15 a 25 a 35	6. d 16 d 26 a	7. a 17 c 27 b 37	8. c 18 b 28 a 38	9. a 19 C 29	10 c 20 d 30 c 40	(LHR	
			47.	Which (a)	the formal property of	pakcity 2. d 12 b 22 d 32 a	orous ous 3. b 13 b 23 c 33	4. d 14 a 24 c 34 b	5. c 15 a 25 a 35	6. d 16 d 26 a 36 c	7. a 17 c 27 b 37	8. c 18 b 28 a 38	9. a 19 C 29	10 c 20 d 30 c 40	(LHR	
			47.	Which (a)	1. c 11 c 21 d 41	pakcity 2. d 12 b 22 d 32 d 32 42	orous ous 3. b 13 b 23 c 33	4. d 14 a 24 c 34 b 44	5. c 15 a 25 a 35 c 45	6. d 16 d 26 a 36 c 46	7. a 17 c 27 b 37 a 47	8. c 18 b 28 a 38	9. a 19 C 29	10 c 20 d 30 c 40	(LHR	
			. <b>7.</b>	Which (a)	the formal property of	pakcity 2. d 12 b 22 d 32 a	orous ous 3. b 13 b 23 c 33	4. d 14 a 24 c 34 b	5. c 15 a 25 a 35	6. d 16 d 26 a 36 c	7. a 17 c 27 b 37	8. c 18 b 28 a 38	9. a 19 C 29	10 c 20 d 30 c 40	(LHR	

1.	2.	♥3./	4.	5.	6.	7.	8.	9.	10
С	d	b	d	С	d	a	C	а	С
11	12	13	14	15	16	17	18	19	20
С	b	b	a	а	d	С	b	С	d
21	22	23	24	25	26	27	28	29	30
d	d	C	C A	а	a	b	а	а	С
31	32	33	34	35	36	37	38	39	40
d	а	a	b	dc	y <sub>C</sub>	a	b	С	Ç
41	42	43	44	45	46	47			
b	d	С	b	b	b	b	]		



	C	HAPTER	₹ 5		
	THE HALOGENS	SANDT	HE NOBLE G	AES &	pakcity.org
1.	Halogen acid in gaseous state found a	as equilibriu	m mixture of monom	ers and	hexamers
	is: (a) HF (b) HCl	(c)	HBr	(d)	HI
2.	(a) HF (b) HCl  Color of which halogen is not correct	(c) Iv related:	HDI	(d)	111
	(a) F <sub>2</sub> Colorless gas	(b)	Cl <sub>2</sub> greenish yellow g	jas	
	(c) Br <sub>2</sub> Reddish brown liquid	(d)	I2 grayish black solid		
3.	Mark the element, which can displace	e three halo	gens from their com	pounds:	_
А	(a) F <sub>2</sub> (b) Cl <sub>2</sub> The chemical formula of iodic acid is:	(c)	Br <sub>2</sub>	(d)	$I_2$
4.	The chemical formula of iodic acid is: (a) HI (b) HIO	(c)	HIO <sub>2</sub>	(d)	HIO <sub>3</sub>
5.	The most inert element in noble gas	X 2	11102	(u)	11103
in <del>Ta</del> . Th	(a) He (b) Ne	(c)	Ar	(d)	Kr
6.	In which compound, the oxidation st	ate of xeno	n is not +6:		
_	(a) XeOF <sub>2</sub> (b) XeOF <sub>4</sub>	(c)	XeO <sub>3</sub>	(d)	XeF <sub>6</sub>
7.	Radon is formed by the removal of al		~ ~ ~ )		
	(a) Radium (c) Rhodium	(b) (d)	Rhenium Rutherfordium		
8.	Which of the following gases is us	(-)		er treatr	ment and
	earthquake predictions?	SHE!			
	(a) Ar (b) Ne	(C)	Xe	(d)	Rn
9.	Chlorine heptaoxide (Cl <sub>2</sub> O <sub>7</sub> ) reacts w	MA 5"	2797	RW 08, BV	VP 09)
	<ul><li>(a) hypochlorous acid</li><li>(c) chloric acid</li></ul>	(b) (d)	perchloric acid chlorine and oxygen		
10.	The anhydride of HClO <sub>4</sub> is:	(u)	chiornie and oxygen		
_0.	(a) $Cl_2O_7$ (b) $Cl_2O_5$	(c)	ClO₃	(d)	ClO <sub>2</sub>
11.	Silver bromide is used in:			(LHR	2010)
	(a) Paints (b) Photogra	phy (c)	Ceramics	(d)	
10	Gasoline Which of the following acid is used for	r stabing o	f alocs:		
12.	Which of the following acid is used for (a) HF (b) HCI	c) (c)	HBr	(d)	HI
13.	Oxidation state of chlorine in HClO <sub>4</sub> is			(u)	1 11
	(a) -7 (b) +7	(c)	-1	(d)	+1
14.	Which is used for making unshrinkab			as NACO	g specialists
	(a) HBr (b) I <sub>2</sub>	(c)	Bleaching powder	(d)	HCI
15.	(a) Rn (b) Kr	rediction (c)	Xe	(d)	Ar
16.	Which of the following compound is	\ /	ΛC	(u)	Λl
	(a) KCl. Mg (OH) <sub>2</sub> . 6H <sub>2</sub> O	(b)	KOH. MgCl <sub>2</sub> 6H <sub>2</sub> O		
	(c) KOH. Mg(OH) <sub>2</sub> . 6H <sub>2</sub> O	(d)	KCl. MgCl <sub>2</sub> .6H <sub>2</sub> O		
17.	Chlorine dioxide is a:	20.5			
	(a) Red gas	(d)	Pale yellow gas		
18.	(c) Orange gas  Iodine pentoxide acts as a:	(d)	Green gas		
<b>-9</b> :	(a) Reducing Agent	(b)	Oxidizing Agent		
	(c) Dehydrating Agent	(d)	None of these		
19.	The Chemical formula of Perchloric a			gy provi	g sign time. And
	(a) HClO (b) HClO <sub>3</sub>	(c)	HCIO <sub>4</sub>	(d)	HClO <sub>2</sub>

20.	The o	xidation state	es of Xe	in its compou	unds rang	e from:			
	(a)	+2 to +8			(b)	+3 to +			
	(c)	+11 to +8			(d)	+3 to +	<b>⊦7</b>		
21.	000 000	ompounds of	Xe are:		71.5				
	(a)	Saturated			(b)	Unsatu			
22	(C)	Stable	ulua la u	and to fill floor	(d)	Unstab	ie		
22.		of the follow		G-1	0-1-	2.20		(4)	Noon
23.	(a) Which	Krypton  is the strong	(b)	Argon •	(c)	Xenon		(d)	Neon
25.	(a)	HI	(b)	HCl	(c)	HBr		(d)	HF
24.		of the follow	X 2					(4)	
	(a)	Fluorine	(b)	Chlorine	(c)	Bromin	e	(d)	Iodine
25.	Which	amongst the	e followi	ng is the sma	allest átoi	m?			
	(a)	F	(b)	Cl	(c)	Br		(d)	I
26.	Fluori		nave pos		on states	12	e absence of	p (840)	
	(a)	d-orbital	(b)	s-orbital	(c)	p-orbita	al	(d)	None
27.		of the follow	71.		ucing pov			(A. 15)	1.14
20	(a)	HI f.thfalles	(b)	HBr	(c)	HCI		(d)	HI
28.	75. A.	of the follow	71.5	ments show (		_	state in its co	mpou	nas:
20	(a)	r . balagan !a	(b)	Ul Stronocities	(c)	Br	^	(a)	1
29.	7 3	n halogen is n	71.5	CI opositive:	(c)	Br a	0	(d)	т
30.	(a) Fluori	ne is a stro	(b)	idizina agen	V		ري in aqueous so	` ,	ı n This i
<b>30.</b>	Tolk the Section	uted to many	2000		it tilali t		iii aqueous si	Jiucioi	1. 11115 13
	(a)	Heat of disso		Схеореі	(b)	Flectro	n affinity		
	(c)	Ionization po	_		(d)	~	hydration		
31.	` '	hing powder		ith a few dro			2 Table 1 Tabl		
	(a)	Chlorine	(b)	Hypochloro	/ <b>-</b> /	(c)	A - A - A - A - A - A - A - A - A - A -	(d)	Oxygen
32.		leaching acti				. ,		. ,	,,,
	(a)	Reduction		MIDD	(b)	Hydrog	enation		
	(c)	Chlorination	2	10	(d)	Oxidati	on		
33.	Eleme		1141	ollowing grou	ips will fo		ns most readily	y:	
	(a)	Oxygen fam	ly		(b)		n family		
	(c)	Halogens	_//		(d)	Alkali m	netals		
34.	Z	alogen that is						(-I)	T.,
<b>2</b> E	(a)	t <sub>2</sub>	(b)	Cl <sub>2</sub>	(c)	Br <sub>2</sub>		(d)	$I_2$
35.		n of the follow	ving is m (b)	HBr	(6)	HCI		(d)	HF
36.	(a) Sodiu				(C) C. <b>H</b> 2 <b>SO</b> 4 <b>a</b>		potassium dic		
	(a)	Chromic chlo		Jon Willi Colle	(b)		/l chloride	91116	31463
	(c)	Chromous cl		nake	(d)	None o			
<b>37.</b>	. ,	of the follow		nonoatomic q	The state of the s				
	(a)	Oxygen	===		(b)	Neon			
	(c)	Fluorine			(d)	Nitroge	n		
38.	Which	n of the follow	ving fluo	orides of xend	on is impo	ssible?			
<u>-</u>	(a)	XeF <sub>2</sub>	(b)	XeF₃	(c)	XeF <sub>4</sub>		(d)	XeF <sub>6</sub>
39.		ollowing show	2301 -20	-	2 00	<u> 2004</u>		مين پر	E/S
	(a)	Kr	(b)	Be	(c)	Al	•	(d)	Na
40.	V	of the follow		**************************************	is <del>es</del>	- X-	here:	7.15	
A 4	(a)	He	(b)	Ne	(c)	Ar Cun and	than an tha as	(d)	Rn
41.	* *	: <b>-</b> :	OWN I WY	2.127	02 03	-	then on the ea	754 6447	Holium
	(a)	Argon	(b)	Xenon	(c)	Neon		(d)	Helium

42.	The last	: memb	er of t	he fan	nilv of i	inert a	ases is	•					
	(a)	Argon		(b)	Rado	n -	(c)		enon			(d)	Neon
43.	XeF <sub>6</sub> on (a)	partial XeF2	hydro	olysis p (b)	roduce XeOF		(c)	Y	eOF <sub>4</sub>			(d)	XeO₃
44.			ollowin				. ,			of elec	trons	/A 2	outermost
	shell:	N		<i>(</i> 1.)	- -		(-)	. 200				<i>(</i> 1)	1.T P
45.	(a) <b>The val</b> ı	Neon ue of io	nizatio	(b) on pote	Rado ential f		(c) rt gase		rgon			(d)	Helium
	(a)	Zero		(b)	Low		(c)		igh			(d)	
46.	Negative The low		lina na	nint of	halium	ı ic du	a to:						
-101		Inertnes		Jille OI	liciidii	i is du	c to:						
	<b>Y</b> - 4	Gaseous											
		High po Weak V			forces b	oetweei	n atoms	e }					
47.	Which (	of the f											
	(LHR 20 (a)	1 <b>4)</b> Bond er	ergy o	f F <sub>2</sub> is le	ess thar	ı Cl <sub>2</sub>	(b)	В	ond ene	ergy of	F <sub>2</sub> is le	ss than	$I_2$
	(c)	Bond er	ergy o	f Cl <sub>2</sub> is I	ess tha	n F <sub>2</sub>	(d)			ergy of			
48.	Goiter i	i <b>s caus</b> e Flourine		to the	deficion Brom	3.5	<b>f:</b> (c)	C	hlorine			(d)	Iodine
49.	Which							soluti	on?	0)		TB. T. 250	
	(a)	HF		(b)	HBr		(c)	() ()	3WP 14	, FSD, I	.HR 13	<b>, GRW</b> : (d)	<b>13,14)</b> HCl
50.	The cor	npound	l whic	h caus	es buri	n to sk	• •	heels	slowly	y:		<b>11</b> (I	LHR 2011)
	(a)	F <sub>2</sub>		(b)	Cl <sub>2</sub>	NOW	(c)	B B	r <sub>2</sub>			(d)	Acid
					<u>A</u>	NZW	EH KE	<u>Y</u> -&	pakcity	org 🛞			
		1.	2.	3.	4,\<	5.	6.	7.	8.	9.	10.	]	
		а	a	a	d	a	a	a	d	b	a	]	
		11	12	13	14	15	16	17	18	19	20	-	
		21	22	23	24	25	<b>26</b>	<b>27</b>	28	2 <b>9</b>	<b>30</b>	-	
			b	a	а	a	a	d	a	d	С		
		31	32	33	34	35	36	37	38	39	40		
		4 <b>1</b>	d <b>42</b>	<b>43</b>	<b>44</b>	<b>45</b>	4 <b>6</b>	<b>47</b>	48	<b>49</b>	<b>50</b>	-	
		d	<b>b</b>	C	d	C	d	a	d	a	C	1	
						oakci	ty.or	9 7		<u> </u>		_1	

1.	2.	3.	4.	\\ <b>5</b> .	6.	7.	8.	9.	10.
a	a	a	d	a	a	a	d	b	a
11	12	13/	14	15	16	17	18	19	20
·þ	a <	DP	C	a	d	b		С	a
21	22	23	24	25	26	27	28	29	30
С	b	a	a	a	a	d	а	d	С
31	32	33	34	35	36	37	38	39	40
а	d	С	a	d	ь	b	b	a	а
41	42	43	44	45	46	47	48	49	50
d	b	С	d	С	d	a	d	а	С

					ADTE	2.6		
				CH	APTER			
			Ī	RANSIT:	ION E	LEMENTS		
1.		dination number	er of Fe	e in [Fe(CN) <sub>6</sub> ]	<sup>-4</sup> ion is:			
_	(a)	<b>2010)</b> 4	(b)	2	(c)	6	(d)	-4
2.	In [Ca	o(NH <sub>3</sub> ) <sub>6</sub> ]+3 the Zero	coordi (b)	nation numb	er of coba (c)	It is: Four	(d)	Six
3.	100 000		0.5		25-7-26	structure of [Cu		SIX
.25	(a)	Square plana			` '	Octahedral	(d)	Linear
4.	(a)	n one of the fo K4CrO4	llowing (b)	K <sub>2</sub> Cr <sub>2</sub> O <sub>4</sub>	has oxida (c)	tion state of chro Cr <sub>2</sub> O <sub>6</sub>	mium other (d)	r <b>than +6:</b> CrCl <sub>3</sub>
5.	X 2	idic medium, p	V 2		. ,		(u)	CICIS
	(a)	Oxidizing age	nt (b)	Reducing a		An Acid	(d)	A base
6.	Interaction (a)	stitial compou Fe	n <b>ds are</b> (b)	tormed by:	(c)	Со	(d)	All
7.	<b>`</b>	ге h element is al	(-)	resent with i	\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \		(u)	ΔII
	(a)	Aluminium	(b)	Copper	(c)	Carbon	(d)	Nickle
8.	951. AND	20.5	100 Eq. (20)		E 15	highest oxidation	원 한번()	7n
9.	(a) <b>Follo</b> v	Mn wina property	(b) <b>of tran</b>	Cr <b>sition eleme</b> r	(c) n <b>ts does n</b>	ot vary with a reg	(d) Jular patter	Zn 'n:
		2011)			332	3)	, and a particular	
	(a)	Binding energ	У		(b)	Covalent radius		
10.	(c) The c	Melting point oordination nu	ımber d	of transition	element ir	Cationic radius  [Co(NO2)3(NH3	3)31 is:	
10.	(a)	3	(b)	4	(c)	6	(d)	0
11.	75 AND 11	IB of transiti	on elen	nents contair	24 2			
	(a) (c)	Zn, Cd, Hg Fe, Ru, Os	N	Dir	(b) (d)	Cu, Au, Ag Cr, Mo, w		
12.		e	ding e	nergy of tra		ements depends	upon: (LHI	R 13, GRW
		, SGD 11)	octron n		(b)	Number of uppai	rad alactrons	
	(a) (c)	Number of ele Number of Ne		ENIM	(b) (d)	Number of unpai Number of protoi		
13.	First	transition serie		s with:		The second secon		
4.4	(a)	Y	(b)	Sc	(c)	Zn	(d)	Cd
14.	Group (a)	IIB of transit Zn, Cd, Hg	ion ele	ments conta	ins (b)	Cu,Au, Ag		
	(c)	Fe, Ru, Os			(d)	Cr, Mo, W		
<b>15.</b>	12 TE	hape of ions co	ontaini	ng dsp³ hybri				
	(a) (c)	Tetrahedral Octahedral			(b) (d)	Trigonal bipyram Square planar	ıdal	
16.	. ,	dination number	er of ire	on in K₃ [Fe((		oquare planai		
	(a)	6	(b)	4	(c)	1	(d)	2
17.		aramagnetic b	ehavio	r is the stron	_	Eo3+ and M=2+		
	(a) (c)	Fe and Mn Fe <sup>2+</sup> and Mn <sup>3-</sup>	٠		(b) (d)	Fe <sup>3+</sup> and Mn <sup>2+</sup> Fe <sup>2+</sup> and Mn <sup>2+</sup>		
18.	The o	re of iron:			(~)			
4 ~	(a)	Fe <sub>3</sub> O <sub>4</sub>	(b)	Fe <sub>2</sub> O <sub>2</sub>	(c)	FeO	(d)	Fe(OH)
19.	The C	Chemical formu MnSiO	l <b>la of S</b> (b)	lag is: MnSiO <sub>2</sub>	(c)	MnSiO <sub>3</sub>	(d)	Mn <sub>2</sub> SiO <sub>2</sub>
	(a)	HIISIU	(0)	11113102	(0)	11113103	(u)	1111/2310/2

20.	Chromates are salts of:	7. IV	
21.	(a) HCrO <sub>3</sub> (b) H <sub>2</sub> CrO <sub>4</sub> (c) HCr <sub>2</sub> O <sub>7</sub> The color of all the chromates is:	(d)	H <sub>2</sub> CrO <sub>6</sub>
22.	(a) White (b) Red (c) Blue  Which of the following can also be prepared by Stadeler's process	(d) <b>s:</b>	Yellow
23.	(a) H <sub>2</sub> SO <sub>4</sub> (b) K <sub>2</sub> SO <sub>4</sub> (c) KMnO <sub>4</sub> <b>Medium carbon steel is used in making:</b>	(d)	H <sub>2</sub> S
	(a) Castings (b) Hammer		
24.	(c) Tubes (d) All of above Which of following is a very powerful oxidant:		
	(a) Sulphates (b) Dichromates (c) Nitrates (d) Chromates		
25.		te value of m	agnetic
26	(a) $Sc^{3+}$ (b) $Ti^{3+}$ (c) $Cu^{+}$	(d)	Zn <sup>2+</sup>
26.	Which of the following metal exhibits more than one oxidation states (a) Na (b) Mg (c) Fe	ates: (d)	Al
27.	The equilibrium $Cr_2O_7^{2-} \leftrightarrow 2CrO_4^{2-}$ is shifted to right in: (a) An acidic medium (b) A basic medium		
20	(c) A neutral medium (d) It does not exist		
28.	Bessemer converter is used in the manufacture of: (a) Pig iron (b) Steel		
29.	(c) Wrought iron (d) Cast iron  The number of unpaired electrons in Ferrous ion (Z = 26) is		
	(a) 1 (b) 2 (c) 4	(d)	5
30.	(a) Zn (b) Sn		
31.	(c) Ni  Choose the correct answer about transition elements:	e	
	<ul> <li>(a) Transition elements have low melting points</li> <li>(b) Transition elements do not have catalytic activity</li> </ul>		
	(c) Transition elements exhibit variable oxidation states		
32.	(d) Transition elements exhibit inert pair effect  The total number of inner transition elements in the periodic table	le is:	
33.	(a) 10 (b) 14 (c) 28  The number of unpaired electrons in Mn <sup>2+</sup> (Z=25) is:	(d)	30
7020 72	(a) 5 (b) 4 (c) 3	(d)	2
34.	(a) 5 (b) 6 (c) 3	(d)	4
35.	In the manufacture of steel by open hearth process, the slag obta (a) CaSiO <sub>3</sub> (b) FeSiO <sub>3</sub> (c) MnSiO <sub>3</sub>	ained is: (d)	All
36.	Which of the following is not an element:  (a) Graphite (b) Diamond		
	(c) 22-Carat gold (d) Rhombic sulphu		
37.	How many moles of acidified FeSO <sub>4</sub> solution can be completely ox KMnO <sub>4</sub> :	kidized by one	mole of
38.	(a) 10 (b) 5 (c) 6  An element in +3 oxidation state has the electronic configuration	(d) n <b>(Ar) 3d³. Its</b>	2 atomic
-	number is:		
39.		75 557	21
	(a) Zn (b) Fe <sup>2+</sup> (c) Ni <sup>3+</sup>	(d)	Cu <sup>+</sup>

40.	Group VI B of transition elements contains:	
	(a) Zn,Cd,Hg (b) Fe, Fu, Os	
um ma	(c) Cr, Mo, W (d) Mn, Te, Re	ia venotati – pet Bratilli
41.		1.00
42.	(a) Cr <sub>2</sub> OCl <sub>2</sub> (b) CrO <sub>2</sub> Cl <sub>2</sub> (c) Cr <sub>2</sub> OCl <sub>3</sub> (d)  The geometrical shape of PCl <sub>5</sub> is: (RW	CrOCl <sub>2</sub> <b>P 2009)</b>
741	(a) Octahedral (b) Square planar	P 2009)
	(c) Tetrahedral (d) Trigonal bipyramidal	
43.	Coinage metals are present in the periodic table in group: (MT	N 2008)
	(a) I-A (b) I-B (c) II-A (d)	II-B
44.		N 2009)
45.	(a) sp (b) dsp <sup>2</sup> (c) spd <sup>2</sup> (d)  The chemical formula of hematite is: (MTN 09, BV	dsp <sup>3</sup>
77.	(a) Fe <sub>2</sub> O <sub>3</sub> (b) Fe <sub>3</sub> O <sub>4</sub>	WP 12)
	(c) FeO (d) Fe <sub>2</sub> O <sub>3</sub> .3H <sub>2</sub> O	
46.		N 2010)
	(a) 2 (b) 3 (c) 6 (d)	7
47.		2010)
40	(a) Zinc (b) Iron (c) Scandium (d)	Copper
48.	and the second control of the contro	P 2011)
49.	(a) Wrought iron (b) Pig iron (c) Cast iron (d)  The central metal atom along with ligand is called: (DG)	Steel <b>K 2008)</b>
73.	(a) Coordination number (b) Coordination sphere	K 2000)
	(c) Chelates (d) none of these	
<b>50.</b>		R 2012)
	(a) 0.1 to 0.2% 0.2 to 0.7%	
	(c) 0.2 to 0.6% (d) 0.1 to 0.6%	
51.		P 2012)
	(a) 0.25 to 2.5% (b) 0.12 to 0.20% (c) 3.0 to 4.5% (d) 2.0 to 4.5%	
<b>52.</b>		electrons?
	(SGD,GRW 2015)	2
<b>F</b> 2	(a) $O_2$ (b) $O_2^+$ (c) $O_2^-$ (d)	O <sub>2</sub> -2
53.	Which of the following species has the maximum number of unpaired (DGK,BWP 2014)	electrons?
	(a) Fe (b) $Fe^{+2}$ (c) $Mn^{+2}$ (d)	Cr <sup>+3</sup>
54.		R 2014)
	(a) Zero (b) Two (c) Four (d)	Six
	pakcity.org	
	ANSWER KEY	
	1. 2. 3. 4. 5. 6. 7. 8. 9. 10.	
	c d a abd a d c a d c	
	11         12         13         14         15         16         17         18         19         20           b         b         b         a         b         a         c         b	
	21 22 23 24 25 26 27 28 29 30	
	d c a b b c b c d	
	31 32 33 34 35 36 37 38 39 40	
	C C a a d d b a b C 41 42 43 44 45 46 47 48 49 50	
	b d b d a b a b a	
	51 52 53 54	
	a a c c	

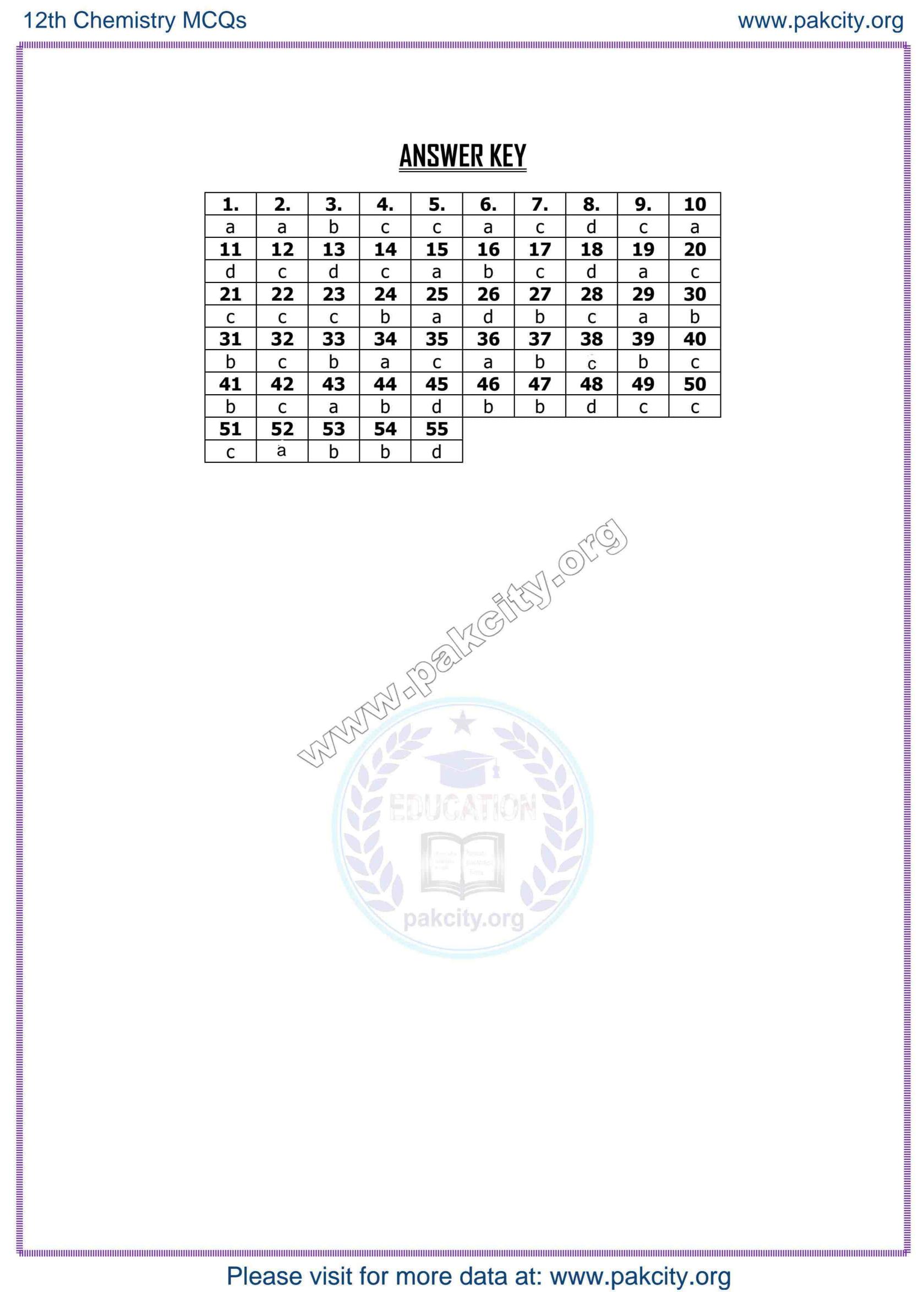
1.	2.	3.	4.	5.	6.	7.	8.	9.	10.
С	d	a	a b d	a	d	С	a	d	С
11	12	13	14	15	16	17	18	19	20
b	b	b	a	b	a	b	a	С	b
21	22	23	24	25	26	27	28	29	30
d	С	a	b	b	С	b	b	С	d
31	32	33	34	35	36	37	38	39	40
С	С	a	a	d	d	b	a	b	С
41	42	43	44	45	46	47	48	4.9	50
b	d	b	d	a	а	b	а	b	a
51	52	53	54						
a	a	С	С						

The total coal resources of Pakistan estimated by geological survey are:							
1. The total coal resources of Pakistan estimated by geological survey are:  (a) 184 billion tonnes (b) 184 million tonnes (c) 841 billion tonnes (d) 184 tonnes  2. How many isomers are possible for C <sub>6</sub> H1 <sub>24</sub> :  (a) 4 (b) 5 (c) 6 (d) 7  3. Which one of the following is an isomer of Dimetly ether:  (a) (CH <sub>3</sub> ) <sub>2</sub> CHOH (b) CH <sub>3</sub> CH <sub>2</sub> OH (c) CHO-CHO (d) None  4. The hardest form of coal is:  (a) Bituminous (b) Sub-bituminous coal (c) Anthracite (d) Lignite  5. Which one of the following compounds is heterocyclic?  (a) Anthracene (b) Phenol (c) Pyridine (d) Aniline  6. Which molecule has a tetrahedral shape?  (a) CH <sub>3</sub> -CH <sub>3</sub> (b) CH <sub>2</sub> =CH <sub>2</sub> (c) CH=CH (d) None  7. The formula of aniline is:  (a) C <sub>6</sub> H <sub>5</sub> -NH <sub>2</sub> (b) C <sub>4</sub> H <sub>2</sub> -NH <sub>2</sub> -C <sub>6</sub> H <sub>5</sub> (c) CH <sub>5</sub> -NH <sub>2</sub> -C <sub>6</sub> H <sub>5</sub> (d) C <sub>6</sub> H <sub>5</sub> -ND <sub>2</sub> 8. The formula of cyclopropane is:  (a) C <sub>6</sub> H <sub>8</sub> (b) C <sub>3</sub> H <sub>4</sub> (c) C <sub>3</sub> H <sub>5</sub> (d) C <sub>3</sub> H <sub>5</sub> (d) C <sub>3</sub> H <sub>5</sub> 9. The Octane (d) n-Heptane  (c) Iso-Octane (d) n-Heptane  (c) Iso-Octane (d) n-Hexane  10. Homologues of alkanes differ from each other by:  (a) CH <sub>2</sub> group (b) CH <sub>3</sub> group  (c) CH <sub>2</sub> group (b) CH <sub>3</sub> group  (c) CH <sub>2</sub> group (d) CH <sub>3</sub> group  (d) CH <sub>2</sub> group  (e) CH <sub>2</sub> group (d) CH <sub>3</sub> group  (f) CH <sub>2</sub> group (h) Alcohol  11. The major components of coal gas are:  (a) H <sub>2</sub> and CH <sub>4</sub> (c) H <sub>2</sub> and CO  (b) Alcohol  12. The compound in which two alkyl groups are attached to carbonyl group is called:  (a) Ether (b) Alcohol  (b) Ketone (d) NH <sub>4</sub> ) <sub>2</sub> CO  (c) (NH <sub>4</sub> ) <sub>2</sub> CO  (d) (NH <sub>4</sub> ) <sub>2</sub> CO  (e) NH <sub>4</sub> ) <sub>2</sub> CO  (f) NH <sub>4</sub> ) <sub>2</sub> CO  (g) NH <sub>4</sub> ) <sub>2</sub> CO  (h) Gasoline  (h) Gasoline  (c) (C <sub>2</sub> H <sub>3</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>3</sub> ) <sub>4</sub> Pb  (c) (C <sub>2</sub> H <sub>3</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>3</sub> ) <sub>4</sub> Pb  (c) (C <sub>2</sub> H <sub>3</sub> ) <sub>2</sub> Pb (b) (Hoffman			CHA	PTE	R 7		
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(a) 4 (b) 5 (c) 6 (d) 7  3. Which one of the following is an isomer of Dimetty ether: (a) (CH3)2 CHOH (b) CH3CH2OH (c) CHO-CHO (d) None  4. The hardest form of coal is: (a) Bituminous (b) Sub-bituminous coal (c) Anthracite (d) Lignite  5. Which one of the following compounds is heterocyclic? (a) Anthracene (b) Phenol (c) Pyridine (d) Aniline  6. Which molecule has a tetrahedral shape? (a) CH3-CH3 (b) CH2-CH2 (c) CHECH (d) None  7. The formula of aniline is: (a) CaH5 (b) CJ4-CH3 (d) CGH5-NO2  8. The formula of cyclopropane is: (a) C3H8 (b) C3H4 (c) C3H5 (d) C3H6  8. The formula of cyclopropane is: (a) C3H8 (b) C3H4 (c) C3H5 (d) C3H6  10. Homologues of alkanes differ from each other by: (a) CH2 group (b) CH3 group  (c) CH4 group (d) CH group  11. The major components of coal gas are: (a) H2 and CH4 (c) H2 and CO (d) H2 and CO (e) (H2)2CO (d) (MH2)2CO (f) (H2)2CO (g) (MH2)2CO (d) (MH2)2CO (g) (MH2)2CO (g) (MH2)2CO (d) (MH2)2CO (g) (MH2)2CO (g) (H2)3Pb (d) (C2H5)4Pb (g) (C2H5)4Pb (g) (C3H5)2Pb (d) (C2H5)4Pb (g) (C2H5)4Pb (g) (C3H5)4Pb (g) (G2H5)4Pd (h) Hoffman		2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2		5 - 10 - 1 - 1 - 1 - 1 - 1 - 1 - 1 - 1 -		3	
3.   Which one of the following is an isomer of Dimetly ether:   (a)	2.	real of the second of the seco	_	E 15	6	(4)	7
(a) (CHs): CHOH (b) CH3CH2OH (c) CHO-CHO (d) None  1 The hardest form of coal is:  (a) Bituminous (b) Sub-bituminous coal (c) Anthracite (d) Lignite  5. Which one of the following compounds is heterocyclic?  (a) Anthracene (b) Phenol (c) Pyridine (d) Aniline  6. Which molecule has a tetrahedral shape?  (a) CH3-CH3 (b) CH2=CH2 (c) CHECH (d) None  7. The formula of aniline is:  (a) C6H5-NH2 (b) CH2=CH2 (d) CH3CH4 (d) None  8. The formula of cyclopropane is:  (a) C3H6 (b) C3H4 (c) C3H5 (d) C3H6  9. The Octane number is 100 for:  (a) n-Octane (d) n-Hexane  10. Homologues of alkanes differ from each other by:  (a) CH2: GY group (d) CH3 group  (c) CH4: group (d) CH3 group  11. The major components of coal gas are:  (a) H2: and CO (d) H2: and CO  12. The compound in which two alkyl groups are attached to carbonyl group is called:  (a) Ether (c) Ketone (d) Aldehyde  (c) Aldehyde (c) Aldehyde  (c) Alcohols  14. The Chemical formula of urea is:  (a) (NH4): CO (d) (MH2): CO  (b) (NH2): CO  (c) (CH3:) 3Pb (b) (C2H5:) 4Pb  (c) (C15:) Gasoline (d) (C2H5:) 4Pb  (c) (C15:) Gasoline (d) (C2H5:) 4Pb  (c) (C15:) Gasoline (d) (C2H5:) 4Pb  (c) (C15:) Gasolice (d) (C2H5:) 4Pb  (c) (C2H5:) 4Pc  (d) (C3H5:) 4Pc  (d) (C3	3.				Ħ .	(u)	/
(a)   Bituminous   (b)   Sub-bituminous coal   (c)   Anthracite   (d)   Lignite		(a) (CH <sub>3</sub> ) <sub>2</sub> CHOH (b)	) CH₃CH₂OH		•	(d)	None
Section   Color   Section   Secti	4.	- 10 : 40: 10: 10: 10: 10: 10: 10: 10: 10: 10: 1	의 교육(대통령)의 20	ااد دراعا	(c) Anthracito	(d)	Lianite
6.	5.				X 2	(u)	Ligitic
(a) CH₃-CH₃ (b) CH₂=CH₂ (c) CH∃CH (d) None	•		And the second s		Pyridine	(d)	Aniline
7. The formula of aniline is: (a) C6H5 (b) C6H5-NH2 (d) C6H5-ND2  8. The formula of cyclopropane is: (a) C3H8 (b) C3H4 (c) C3H5 (d) C3H6  9. The Octane number is 100 for: (a) n-Octane (d) n-Hexane  10. Homologues of alkanes differ from each other by: (a) CH2 group (d) CH3 group (c) CH4 group (d) CH group  11. The major components of coal gas are: (a) H2 and CO (d) H2, CH4 and CO  12. The compound in which two alkyl groups are attached to carbonyl group is called: (a) Ether (b) Alcohol (c) Ketone (d) Aldehyde  13. Carbonyl group is present in: (a) Aldehyde (b) Ketone (c) Alcohols  14. The Chemical formula of urea is: (a) (NH4)2CO (b) (NH2)4C (c) (NH2)2CO (c) (NH2)2CO (d) (NH2)4C (c) 15% Gasoline (d) 10% Gasoline (c) 15% Gasoline (d) 10% Gasoline (c) (C2H5)3Pb (b) (C2H5)4Pb (c) (C2H5)3Pb (d) (C2H5)4Pd  17. Father of Organic Chemistry is: (a) Faraday (b) Hoffman	о.		CO. Check Co. P. H. Co. Co. Co.		CHECH	(d)	None
8. The formula of cyclopropane is: (a) C <sub>3</sub> H <sub>8</sub> (b) C <sub>3</sub> H <sub>4</sub> (c) C <sub>3</sub> H <sub>5</sub> (d) C <sub>3</sub> H <sub>6</sub> 9. The Octane number is 100 for: (a) n-Octane (b) n-Heptane (c) Iso-Octane (d) n-Hexane  10. Homologues of alkanes differ from each other by: (a) CH <sub>2</sub> group (b) CH <sub>3</sub> group (c) CH <sub>4</sub> group (d) CH group  11. The major components of coal gas are: (a) H <sub>2</sub> and CH <sub>4</sub> (b) Ethane and CO (c) H <sub>2</sub> and CO (d) H <sub>2</sub> , CH <sub>4</sub> and CO  12. The compound in which two alkyl groups are attached to carbonyl group is called: (a) Ether (b) Alcohol (c) Ketone (d) Aldehyde (e) Alcohols (d) both a and b  14. The Chemical formula of urea is: (a) (NH <sub>4</sub> ) <sub>2</sub> CO (b) (NH <sub>2</sub> ) <sub>4</sub> CO (c) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>2</sub> ) <sub>2</sub> CO (e) (15% Gasoline (d) 10% Gasoline (c) I5% Gasoline (d) C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (e) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd  17. Father of Organic Chemistry is: (a) Faraday (b) Hoffman	7.			(-)	200	(-)	
8. The formula of cyclopropane is:  (a) C₃H₃ (b) C₃H₄ (c) C₃H₅ (d) C₃H₆  9. The Octane number is 100 for:  (a) n-Octane (b) n-Heptane (c) Iso-Octane (d) n-Hexane  10. Homologues of alkanes differ from each other by:  (a) CH₂ group (b) CH₃ group (c) CH₄ group (d) CH group  11. The major components of coal gas are:  (a) H₂ and CH₂ (b) Ethane and CO (c) H₂ and CO (d) H₂₂ CH₄ and CO  12. The compound in which two alkyl groups are attached to carbonyl group is called:  (a) Ether (b) Alcohol (c) Ketone (d) Aldehyde  13. Carbonyl group is present in: (a) Aldehyde (b) Ketone (c) Alcohols (d) both a and b  14. The Chemical formula of urea is: (a) (NH₂)₂CO (b) (NH₂)₂CO (c) (NH₂)₂CO (d) (OH₂)₃Pb (d) (C₂H₅)₃Pb (d) (C₂H₅)₃Pb (d) (C₂H₅)₃Pb (d) (C₂H₅)₃Pd  17. Father of Organic Chemistry is: (a) Faraday (b) Hoffman							
(a) C <sub>3</sub> H <sub>8</sub> (b) C <sub>3</sub> H <sub>4</sub> (c) C <sub>3</sub> H <sub>5</sub> (d) C <sub>3</sub> H <sub>6</sub>     The Octane number is 100 for:   (a) n-Octane (b) n-Heptane (c) Iso-Octane (d) n-Hexane     (b) n-Hexane (d) n-Hexane (d) n-Hexane (d) n-Hexane (e) (e) CH <sub>3</sub> group (f) CH <sub>3</sub> group	8.		pane is:	(u)	C6D5-INO2		
(a) n-Octane (b) n-Heptane (c) Iso-Octane (d) n-Hexane  10. Homologues of alkanes differ from each other by: (a) CH <sub>2</sub> group (b) CH <sub>3</sub> group (c) CH <sub>4</sub> group (d) CH group  11. The major components of coal gas are: (a) H <sub>2</sub> and CH <sub>4</sub> (b) Ethane and CO (c) H <sub>2</sub> and CO (d) H <sub>2</sub> , CH <sub>4</sub> and CO  12. The compound in which two alkyl groups are attached to carbonyl group is called: (a) Ether (b) Alcohol (c) Ketone (d) Aldehyde  (b) Ketone  13. Carbonyl group is present in: (a) Aldehyde (b) Ketone (c) Alcohols (d) both a and b  14. The Chemical formula of urea is: (a) (NH <sub>4</sub> ) <sub>2</sub> CO (b) (NH <sub>2</sub> ) <sub>4</sub> C (c) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> 15. The fractional distillation of petroleum yields only: (a) 20% Gasoline (b) 30% Gasoline (c) 15% Gasoline (d) 10% Gasoline  16. Antiknocking agent is: (a) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd  17. Father of Organic Chemistry is: (a) Faraday (b) Hoffman	<u> </u>	(a) C <sub>3</sub> H <sub>8</sub> (b)	) C <sub>3</sub> H <sub>4</sub>	(c)	C₃H₅	(d)	C₃H <sub>6</sub>
(c) Iso-Octane (d) n-Hexane  10. Homologues of alkanes differ from each other by:	9.	Service Carlos	00 for:	\$ (b)	n Hontono		
10. Homologues of alkanes differ from each other by:			200	5.5			
(c) CH <sub>4</sub> group  The major components of coal gas are:  (a) H <sub>2</sub> and CH <sub>4</sub> (b) Ethane and CO (c) H <sub>2</sub> and CO (d) H <sub>2</sub> , CH <sub>4</sub> and CO  12. The compound in which two alkyl groups are attached to carbonyl group is called: (a) Ether (b) Alcohol (c) Ketone (d) Aldehyde (e) Ketone (d) Aldehyde (e) Alcohols (f) Ketone (h) Ket	10.	Homologues of alkanes	differ from each o				
11. The major components of coal gas are:  (a)			MAN				
(a) H <sub>2</sub> and CH <sub>4</sub> (b) Ethane and CO (c) H <sub>2</sub> and CO (d) H <sub>2</sub> , CH <sub>4</sub> and CO  12. The compound in which two alkyl groups are attached to carbonyl group is called: (a) Ether (b) Alcohol (c) Ketone (d) Aldehyde  13. Carbonyl group is present in: (a) Aldehyde (b) Ketone (c) Alcohols (d) both a and b  14. The Chemical formula of urea is: (a) (NH <sub>4</sub> ) <sub>2</sub> CO (b) (NH <sub>2</sub> ) <sub>4</sub> C (c) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> 15. The fractional distillation of petroleum yields only: (a) 20% Gasoline (b) 30% Gasoline (c) 15% Gasoline (d) 10% Gasoline  16. Antiknocking agent is: (a) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd  17. Father of Organic Chemistry is: (a) Faraday (b) Hoffman	11.		of coal gas are:	(u)	Ci i gioup		
12. The compound in which two alkyl groups are attached to carbonyl group is called:		(a) H <sub>2</sub> and CH <sub>4</sub>					
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13. Carbonyl group is present in:  (a) Aldehyde (b) Ketone (c) Alcohols (d) both a and b  14. The Chemical formula of urea is:  (a) (NH <sub>4</sub> ) <sub>2</sub> CO (b) (NH <sub>2</sub> ) <sub>4</sub> C (c) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> 15. The fractional distillation of petroleum yields only:  (a) 20% Gasoline (b) 30% Gasoline (c) 15% Gasoline (d) 10% Gasoline  16. Antiknocking agent is:  (a) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd  17. Father of Organic Chemistry is: (a) Faraday (b) Hoffman			ciro anty, groups	20.02		oup is co	aneur
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(c) Alcohols (d) both a and b  14. The Chemical formula of urea is:  (a) (NH <sub>4</sub> ) <sub>2</sub> CO (b) (NH <sub>2</sub> ) <sub>4</sub> C (c) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> 15. The fractional distillation of petroleum yields only:  (a) 20% Gasoline (b) 30% Gasoline (c) 15% Gasoline (d) 10% Gasoline  16. Antiknocking agent is:  (a) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd  17. Father of Organic Chemistry is:  (a) Faraday (b) Hoffman	13.		ic in:	(b)	Ketone		
(a) (NH <sub>4</sub> ) <sub>2</sub> CO (b) (NH <sub>2</sub> ) <sub>4</sub> C (c) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> <b>15.</b> The fractional distillation of petroleum yields only:     (a) 20% Gasoline     (b) 30% Gasoline     (c) 15% Gasoline     (d) 10% Gasoline <b>16.</b> Antiknocking agent is:     (a) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb     (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb     (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb     (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd <b>17.</b> Father of Organic Chemistry is:     (a) Faraday     (b) Hoffman		(c) Alcohols	The state of the s				
(c) (NH <sub>2</sub> ) <sub>2</sub> CO (d) (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> <b>15. The fractional distillation of petroleum yields only:</b> (a) 20% Gasoline (b) 30% Gasoline (c) 15% Gasoline (d) 10% Gasoline <b>16. Antiknocking agent is:</b> (a) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd <b>17. Father of Organic Chemistry is:</b> (a) Faraday (b) Hoffman	14.	선생님 이 전 전 이 이	urea is: Pakcit	y.org	(NH-)-C		
<ul> <li>15. The fractional distillation of petroleum yields only:         <ul> <li>(a) 20% Gasoline</li> <li>(b) 30% Gasoline</li> </ul> </li> <li>(c) 15% Gasoline</li> <li>(d) 10% Gasoline</li> <li>16. Antiknocking agent is:         <ul> <li>(a) (C<sub>2</sub>H<sub>5</sub>)<sub>2</sub>Pb</li> <li>(b) (C<sub>2</sub>H<sub>5</sub>)<sub>4</sub>Pb</li> <li>(c) (C<sub>2</sub>H<sub>5</sub>)<sub>3</sub>Pb</li> <li>(d) (C<sub>2</sub>H<sub>5</sub>)<sub>4</sub>Pd</li> </ul> </li> <li>17. Father of Organic Chemistry is:         <ul> <li>(a) Faraday</li> <li>(b) Hoffman</li> </ul> </li> </ul>		127 27 Statement of Statement o			Paragon and 18 meson ser		
(c) 15% Gasoline (d) 10% Gasoline  16. Antiknocking agent is:  (a) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd  17. Father of Organic Chemistry is:  (a) Faraday (b) Hoffman	<b>15</b> .	The fractional distillation	n of petroleum yi	elds on	y:		
<ul> <li>16. Antiknocking agent is:         <ul> <li>(a) (C<sub>2</sub>H<sub>5</sub>)<sub>2</sub>Pb</li> <li>(b) (C<sub>2</sub>H<sub>5</sub>)<sub>4</sub>Pb</li> <li>(c) (C<sub>2</sub>H<sub>5</sub>)<sub>3</sub>Pb</li> <li>(d) (C<sub>2</sub>H<sub>5</sub>)<sub>4</sub>Pd</li> </ul> </li> <li>17. Father of Organic Chemistry is:         <ul> <li>(a) Faraday</li> <li>(b) Hoffman</li> </ul> </li> </ul>							
(a) (C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> Pb (b) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pb (c) (C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub> Pb (d) (C <sub>2</sub> H <sub>5</sub> ) <sub>4</sub> Pd <b>17. Father of Organic Chemistry is:</b> (a) Faraday (b) Hoffman	16.			(u)	10 /0 00301116		
17. Father of Organic Chemistry is: (a) Faraday (b) Hoffman		(a) $(C_2H_5)_2Pb$					
(a) Faraday (b) Hoffman	17		strv is:	(d)	$(C_2H_5)_4Pd$		
	1/.		30 y 13.	(b)	Hoffman		
					Democritus		

18.	~ .	xample of sp l					~ .s
10	(a)	Methane number of chair	(b)	Benzene	(c)	Ethene	(d) Ethyne
19.	(a)	Three	n isome (b)	Two	e are: (c)	One	(d) Zero
20.		merism is only		11 1101 1000	(0)	Offic	(u) 2010
	(a)	Ethers		<u> </u>	(b)	Ketones	
	(c)	Both a and b			(d)	every organic f	amily
21.		ne number 100	) is give	en to:	<b>(</b> L)	n Hantana	
	(a) (c)	n – octane 2 2 4 – Trime	thylnent	ane (Iso-octan	(b) e) (d)	n – Heptane 2,2,4 – Trimeth	nyl octane
22.		type of hybridi					Tyr Octanic
	(a)	sp			(b)	sp <sup>2</sup>	
	(c)	sp <sup>3</sup>			(d)	None of these	
23.		poiling point ra	inge of	petroleum et	(22)	10 2000	(GRW 2007)
	(a)	5 − 20°C 20 − 60°C			(b) (d)	10 − 30°C 30 − 90°C	
24.	Whic	ch one of the fo	llowing	ı is not a hete	\ <i>'</i>		(LHR 2008)
	(a)	Thiophene	(b)	Anthracene		Furan	(d) Pyrrol
25.	The s	state of hybrid	ization	of carbon ato	m in eth	ane is:	(GRW 2008,12)
~ =	(a)	sp <sup>3</sup>	(b)	sp <sup>2</sup>	(c)	sp	(d) dsp <sup>2</sup>
26.	7 V	netric isomeris	m is pr	esent in:	(b)	Ethano	(GRW 2010)
	(a) (c)	Methane Propane			(b) (d)	Ethane 2-Butene	
27.	` '	ole bond consis	sts of:		(u)	1 Duteric	
	(a)	Two sigma bo			(b)	One sigma one	pi bond
	(c)	One sigma ar			((q))	Two pi bonds	
28.		ch set of hybric		- (0)		~ -	(d) day 2
29.	(a)	sp³ <b>ype of hybridi</b> :	(b)	sp (o)	· (c) m in meth	sp²	(d) dsp <sup>2</sup> (LHR 2007,09)
<b>LJ</b> 1	(a)	sp <sup>3</sup>	(b)	Sp	(c)	sp <sup>2</sup>	(d) dsp <sup>2</sup>
30.	8 8		\	AVIV		ium cyanate wa	
	(a)	G.N. Lewis	My	7/02	(b)	Fredrick Wohle	r
24	(c)	Kolbe	an		(d)	Berzelius	// UD 204 4V
31.	09 NV	process used to Thermal crack	- 17	ve the quality	y or gaso (b)	Reforming	(LHR 2014)
	(a) (c)	Steam crackir		EDUC	(d)	Combustion	
32.	St. 242	omer of C <sub>2</sub> H <sub>5</sub> O					(SGD, GRW 2015)
	(a)	CH₃OH			(b)	$(C_2H_5)_2O$	
22	(c)	CH₃OCH₃			(d)	CH <sub>3</sub> COCH <sub>3</sub>	
33.	Linea	ar shape is ass	ociated	with which s			SD 09, GRW, BWP 09,15
	(a)	sp <sup>3</sup>	(b)	sp pakel	(c)	sp <sup>2</sup>	(d) dsp <sup>2</sup>
34.		rs show the ph				•	14, LHR 14,15, FSD 11)
	(a)	Metamerism			(b)	Functional grou	± 1
25	(c)	Position isom		a aw? bb!	(d)	cis-trans isome	
35.	7 1	on atom in foll CH₃CN	owing i	s sp² hybridiz CHECH	zed (c)	НСООН	(LHR 2011) (d) CH <sub>2</sub> Cl <sub>2</sub>
36.	(a) <b>The</b> i	somerism sho	<b>*</b>		(C)	TICOOTT	(u) CH2CH2 (LHR 2011)
- A	(a)	Skeletal			(b)	Position	(
	(c)	Geometric			(d)	Metamerism	
37.		idization of ca	71.3				D, SGD 15, RWP 10)
	(a)	sp <sup>3</sup>	(b)	sp	(c)	sp <sup>2</sup>	(d) dsp <sup>2</sup>

38.	How many icomore are there in next	tano?	(ECD 00 1 HD 13 CDW 11)
30.	How many isomers are there in pent (a) 6 (b) 5	(c)	( <b>FSD 09, LHR 12, GRW 11)</b> 3 (d) 2
39.	Vital force theory was rejected by:	(0)	(LHR 2012,14)
	(a) G.N. Lewis	(b)	F.Wholer
	(c) Greek Philosophers		cientists of 20th century
40.	Select one which shows cis-trans is	omerism.	<del>-</del> 2
	(FSD 2010)		
	(a) $Cl_2C=CCl_2$	(b)	$CH_2=CH_2$
<b></b>	(c) CICH=CHCI	(d)	Br <sub>2</sub> C=CBr <sub>2</sub>
41.	Which of the following compound m	973 설계를	대한 전 등 함께 등 등 기계 등
40	(a) 1-butene (b) 2-butene		Cyclopropane (d) Aceton
42.	Urea belongs to which class of comp		(RWP 2009)
	(a) Imides	(b)	Amines
43.	(c) Amides  In ethane, each carbon atom is:	(d)	carboxylic acid
73.	2	(b)	(GRW, FSD 12, LHR 10, DGK 09) sp <sup>2</sup> hybridized
	(a) sp <sup>3</sup> hybridized (c) sp hybridized	(d)	unhybridized
44.	The hybridization of carbon atom in	` '	(MTN, RWP 2011)
.a 5.5	(a) sp (b) $sp^2$	(c)	$sp^3$ (d) $dsp$
45.	Geometric isomerism in alkenes is d		(MTN 2008)
(MATASA)	(a) Oscillation of H atoms between to		
	<ul><li>(b) Optical rotation due to multiple b</li></ul>	onds	$\mathcal{A}(\mathcal{S})$
	(c) Free rotation about C=C bond	omo o T	pakcity.org
46	(d) Restricted rotation about C=C bo	ond A	110
46.	Metamerism is shown by:	E 6 2 3 3 3 3 3 3 3 3 3 3 3 3 3 3 3 3 3 3	Ethoro
	(a) Amines	(b)	Ethers  Poth a and b
47.	(c) neither a nor b  Dimethyl ether and Ethyl alcohol are	(d)	Both a and b
47.		(b)	(MTN 2009) Functional group isomers
	(a) Metamers (c) Position isomers	(d)	cis-trans isomers
48.	The self-linking property of element		(MTN 2009)
FG			erization (c) Association
	(d) Catenation		
49.	The carbon atoms in propene are:	4.	(MTN 2010)
	(a) sp $\operatorname{sp}^3$ and $\operatorname{sp}^2$	(b)	sp <sup>2</sup> and sp
50.	(c) sp <sup>3</sup> and sp <sup>2</sup> Peat before conversion into coal is c	(d)	sp <sup>2</sup> and sp (BWP 2008)
<b>JU.</b>	(a) Anthracite	(b)	Asphalt
	(c) Lignite	(d)	all of these
51.	Tetra ethyl lead is added to petrol	to:	(BWP 2010)
	(a) Prevent its freezing point	(b)	Increase its boiling point
	(c) Prevent knocking	ccity (d)	Increase its viscosity
52.	The crude petroleum is separated in	fractions by	
	(a) Filtration	(b)	Fractional distillation
	(c) Steam distillation	(d)	Fractions sublimation
53.	Carbon atom in Dimethyl ether is:	21.3	(DGK 2009)
	(a) sp <sup>3</sup> hybridized	(p)	sp <sup>2</sup> hybridized
E4	(c) sp hybridized	(d)	unhybridized (DCK 2010)
54.	Number of isomers of C <sub>4</sub> H <sub>10</sub> is:	(0)	(DGK 2010)
55	(a) 1 (b) 2  In t-butyl alcohol, the tertiary carbo	(C)	3 (d) 4 to: (MTN SCD 2015)
55.	In t-butyl alcohol, the tertiary carbo (a) Three hydrogen atoms	2/21   3/	to: (MTN, SGD 2015) Two hydrogen atoms
	<ul><li>(a) Three hydrogen atoms</li><li>(c) One hydrogen atom</li></ul>	(b) (d)	No hydrogen atom
	(c) One hydrogen atom	(u)	No mydrogen atom

1.	2.	3.	4.	5.	6.	7.	8.	9.	10
		Francisco Sparition of the Company					4		
a	a	b	С	С	a	С	d	С	a
11	12	13	14	15	16	17	18	19	20
d	С	d	С	a	b	С	d	a	С
21	22	23	24	25	26	27	28	29	30
С	С	С	b	a	d	b	С	a	b
31	32	33	34	35	36	37	38	39	40
b	С	b	a	С	а	b	С	b	С
41	42	43	44	45	46	47	48	49	50
b	С	a	b	d	b	b	d	С	С
51	52	53	54	55					
C	a	b	b	d					





	CH	APTER 8
	ALIPHATIC	HYDROCARBONS * pakcity.org
1.	An Aldehyde is reduced to Alkane w (a) KOH+N <sub>2</sub> H <sub>4</sub>	ith: (b) NaOH
	(c) CaO and NaOH	(d) Co(OH) <sub>2</sub>
2.	Which of the following compound will	not form Metal alkynide:
4.	(a) Ethyne	(b) Propyne
	(c) 1- Butyne	(d) 2-Butyne
3.	(学) (会) (1) (1) (2) (3) (3) (4) (4) (4) (4) (4) (4) (4) (4) (4) (4	passed over heated silver under pressure we get:
	(a) Superoxide	(b) Epoxide
4.	(c) Ozonide  Ethyl chloride when boiled with alco	(d) Benzene holic KOH gives:
# #	(a) Acetylene	(b) Ethylene
	(c) Ether	(d) Ethyl alcohol
<b>5.</b>	What type of reaction occurs between	
	(a) Addition (c) Oxidation	(b) Substitution (d) Dehydration
6.	Which ion is most stable:	(u) Denyaration
0.	(a) CH <sub>3</sub> +	(b) CH <sub>3</sub> -CH <sub>2</sub> +
	(c) (CH <sub>3</sub> ) <sub>2</sub> CH <sup>+</sup>	(d) (CH <sub>3</sub> ) <sub>3</sub> C <sup>+</sup>
7.	Mustard gas is:	
	(a) Highly viscous liquid	(b) Low boiling liquid (d) Colourloss and
8.	(c) High boiling liquid  Select the compound which has acid	(d) Colourless gas
<b>9.</b>	(a) Methane	(b) Ethene
	(c) Butadiene	(d) Acetylene
9.	The characteristic reactions of alkar	
	(a) Polymerization (c) Elimination	(b) Addition (d) Substitution
10.		f acetylene while passing through Cu tube at 300°C
	gives:	
	(a) Benzene	(b) n-Hexane
4 4	(c) Naphthalene	(d) Cyclohexane
11.	The characteristic reactions of alker (a) Polymerization	(b) Addition
	(c) Elimination	kcity.org(d) Addition
12.	The presence of pi bond in a molecu	
	(a) Unsaturation	(b) Inertness
12	(c) Stabilty	(d) Saturation
13.	Vinyl acetylene combines with HCl t (a) Neoprene rubber	o yield: (b) Chloroprene
	(c) Poly vinyl acetylene	(d) White ppt.
14.		
	(a) Alkenyl group	(b) Alkane series
	(c) Alkyl group	(d) None of these
15.	(a) Na <sub>4</sub> C	heated with Sodalime, we get alkane along with: (b) H <sub>2</sub> O
	(a) $Na4C$ (c) $CO_2 + N_2$	(d) Na <sub>2</sub> CO <sub>3</sub>
	(5) 552. 112	(4) 1142003

16.	For each double bond, the heat of hydronic	rogenati	on of Alkene is:	
	(a) 110 K.J mol <sup>-1</sup>		(b) 130 K.J mol <sup>-1</sup>	
	(c) 115 K.J mol <sup>-1</sup>		(d) 120 K.J mol <sup>-1</sup>	
17.	The alkenes react with aqueous solution	on of ha	logen acid to form:	•
	(a) Alcohols		Aldehyde	
92 CZ	(c) Alkanes	(d)	Alky halides	
18.	Which of following is used as a genera	71.3		
	(a) Ethane	(d)	Propane Ethono	
19.	(c) Ethenol  Alkyl halides on treatment with active	. ,	Ethene like Zn vield:	
IJ.	(a) Alkene		Alkyne	
	(c) Alkane		Alcohol	
20.				
	(a) Acetone		Ethyl alcohol	
	(c) Formic acid	(d)	None	
21.	Acetylene gives:			
	(a) White ppt. with ammonical AgNO <sub>3</sub> and I			year the second of the second
	(b) With ppt. with ammonical AgNO <sub>3</sub> and re	ed ppt. w	th ammonical Cu <sub>2</sub> Cl <sub>2</sub>	
	(c) White. ppt. with both			
22	(d) Red ppt. with both	11141		•
22.	,			ions:
	(a) $Br_2 > Cl_2 > F_2$ (c) $F_2 > Cl_2 > Br_2$	` '	$Cl_2 > Br_2 > F_2$ $F_2 > Br_2 > CI_2$	
23.	A CONTROL OF THE CONT	` ,	1	CH = CH2 is:
25.	(a) 1, 1 – Dimethyl-3-butene		1, 1, 1 – Trimethyl-3	
	(c) 3, 3 – Dimethyl-1-butene	~ / /	3, 3,3 – Trimethyl-1	- <del> </del>
24.	For preparing a symmetrical alkane,			
	potassium salt of saturated carboxylic	117		
	(a) Hydrolysis		Oxidation	
	(c) Hydrogenation	(d)	Electrolysis	- → pakcity.org
25.				
	(a) Catalytic hydrogenation of Alkene		Wurtz reaction	
26	(c) Hydrolysis of alkyl magnesium bromide			- 1777
26.	A fuel has the same knocking proper Trimethylpentane) and 30% n-Heptan	_		
	(a) 100	(b)	70 The Octane in	uniber of that fuel is.
	(c) 50	(d)	40	
27.	Hydrocarbon which is liquid at room to			
	(a) Hexane	(b)	Butane	
	(c) Ethane	(d)	Propane	
28.	Marsh gas was the name given to:	city.or		
	(a) Methane	(b)	Ethane	
Samuel	(c) Propane	* 11.5%	Butane	
29.	Each different compound should have	e a dif	ferent name" was	published by IUPAC
	system of nomenclature in:	(2001)	4000	
	(a) 1892	(b)	1830	
20	(c) 1947	(d)	1979	
30.	Write the name of following Alkene Ch			
	(a) 1, 3-Butadiene	38 90	Buta -1, 3 diene	
21	(c) Both a & b  CnH <sub>2n</sub> is the general formula of:	(d)	None	
31.	(a) Alkanes	(b)	Alkenes	
	(c) Alkynes	(d)	None of above	
		(4)	THE OF GROVE	

32.	An Alkane is produced when an Alkyl catalyst; the reaction is called:	halide reacts with Zinc in the presence of a
	(a) Sabatier-Sendern's reaction	(b) Wurtz Synthesis
	(c) Frankland's Reaction	(d) Clemmenson's reduction
33.		d by Alkyl halide in the presence of Palladium
	charcoal, is: (a) Hydrolysis	(b) Electrolysis
	(c) Hydrogenation	(d) Hydrogenolysis
34.	Kolbe's method is not useful for the pro	
	(a) Methane	(b) Ethane
	(c) Butane	(d) Hexane
35.	-	
	<ul><li>(a) Use of electrical energy</li><li>(c) Number of side products</li></ul>	<ul><li>(b) Slow reaction</li><li>(d) Salts used are very expensive</li></ul>
36.	The reaction in which a Ketone is reduce	The state of the s
<b>J J J</b>	(a) Kolbe's electrolysis	(b) Clemmensen's reduction
	(c) Cannizzaro	(d) None
37.	Alkanes containing carbons C-18 onwar	Salt as it is is in the
	(a) Gases	(b) Liquids
20	(c) Waxy solids	(d) Solids
38.	Alkanes are soluble in all except:  (a) Benzene	(b) Ether
	(c) Water	(d) Carbon tetra chloride
39.		ot increase with increase in molar mass:
	(a) Boiling point	(b) Melting point
	(c) Density	(d) Solubility
40.	The low reactivity of alkanes is based up	원보 보다
	(a) Inertness of sigma-bond (c) Both a and b	(b) Non-polarity of the bonds (d) None of above
41.	Complete combustion of alkane yields:	(d) None of above
<del>- 11</del>	(a) CO <sub>2</sub> +H <sub>2</sub> O	(b) CO <sub>2</sub> + Heat
	(c) $CO_2 + H_2O + CO$	(d) CO <sub>2</sub> +H <sub>2</sub> O + Heat
42.	The major reaction occurring in the eng	ACCURATION OF THE PROPERTY OF
	(a) Oxidation	(b) Reduction
12	(c) Combustion  Incomplete oxidation of alkanes yields:	(d) Decomposition
43.	(a) CO <sub>2</sub> & carbon black	(b) CO <sub>2</sub> +CO
	(c) CO+H <sub>2</sub> O+carbon black	(d) CO <sub>2</sub> +heat
44.	The order of reactivity of halogen acids	
	(a) HCl > HBr >HI	(b) HBr >HCl>HI
gge-v	(c) HCl > HBr>HF	(d) HI > HBr >HCl
45.	Raney – Nickel is the alloy of Ni with:	ZI-X AT
	(a) Pt	(b) Al
46.	(c) Cu  Polymerization of Ethene to Polyethyle	(d) Pd ene take place at pressure of 100 atm and a
70.	temperature of:	she take place at pressure of 100 atili and a
	(a) 200°C	(b) 400°C
	(c) 600°C	(b) 400°C (d) 800°C

47. During the preparation of alkynes the active metals that reacts with Tetrahalo-alka is:  (a) Zn (b) Mg (c) Both a and b (d) None  48. Alkynes are colorless & odorless except:  (a) Acetylene (b) Propyne (c) Butyne  49. An Alkyne having Carbon atoms =15 is most probably a:  (a) Gas (c) Solid (d) Waxy solid  50. The reaction of Acetylene with water in H2504 and H9504 yields:  (a) Vinyl alcohol (b) Acetaldehyde (c) Mixture of both a and b (d) None  51. The gas used for illumination:  (a) Methane (b) Ethene (d) None  52. The Alkynides are used for the	17	Duning H		nava!	on of -	llama	. <b>.</b>	ativa -	notal-	+b=+	opets -	with T	otrobolo ollese
(c) Both a and b  48. Alkynes are colorless & odorless except:  (a) Acetylene (b) Propyne (c) Butyne  49. An Alkyne having Carbon atoms = 15 is most probably a: (a) Gas (c) Solid (d) Waxy solid  50. The reaction of Acetylene with water in H2SO4 and HgSO4 yields: (a) Vinyl alcohol (b) Acetaldehyde (c) Mixture of both a and b (d) None  51. The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  52. The Alkynides are used for the	4/.		ie pre	parati	on of a	пкупе	o uie a	cuve r	netals	uidt f	cacts \	WILN I	cu anaio-aikan
48. Alkynes are colorless & odorless except: (a) Acetylene (b) Propyne (c) Butyne (d) Pentyne  49. An Alkyne having Carbon atoms =15 is most probably a: (a) Gas (b) Liquid (c) Solid (d) Waxy solid  50. The reaction of Acetylene with water in H <sub>2</sub> SO <sub>4</sub> and H <sub>2</sub> SO <sub>4</sub> yields: (a) Vinyl alcohol (b) Acetaldehyde (c) Mixture of both a and b (d) None  51. The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  52. The Alkynides are used for the										559			
(a) Acetylene (b) Propyne (c) Butyne (d) Pentyne  49. An Alkyne having Carbon atoms =15 is most probably a: (a) Gas (b) Liquid (c) Solid (d) Waxy solid  50. The reaction of Acetylene with water in H2SO4 and HgSO4 yields: (a) Vinyl alcohol (b) Acetaldehyde (c) Mixture of both a and b (d) None  51. The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  52. The Alkynides are used for the	48	Sec. 1990			& odo	rless e	ycent	•	(a) No	one			
49. An Alkyne having Carbon atoms =15 is most probably a: (a) Gas (b) Liquid (d) Waxy solid  50. The reaction of Acetylene with water in H <sub>2</sub> SO <sub>4</sub> and HgSO <sub>4</sub> yields: (a) Vinyl alcohol (b) Acetaldehyde (c) Mixture of both a and b (d) None  51. The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  52. The Alkynides are used for the	10.			.011033	u ouo	11035	жеерс	41.5	Propyr	ne			
(a) Gas (c) Solid (d) Waxy solid  50. The reaction of Acetylene with water in H2SO4 and HgSO4 yields: (a) Vinyl alcohol (b) Acetaldehyde (c) Mixture of both a and b (d) None  51. The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  52. The Alkynides are used for theof alkynes (a) Preparation (c) Separation (d) All of above  53. Formula of chloroform is: (a) CH <sub>2</sub> Cl <sub>2</sub> (b) CH <sub>2</sub> Cl <sub>2</sub> (c) CHCl <sub>3</sub> ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 a d b b d d c d d a a 11 12 13 14 15 16 17 18 19 20 b a b d d d d c b 21 22 23 24 25 26 27 28 29 30 b c d d b a a d d c 31 32 33 34 35 36 37 38 39 40 b c d a c b c c d c 41 42 43 44 45 46 47 48 49 50 d c c d b b c a c c 51 52 53	40					i			( <b>*</b> )/				
(c) Solid (d) Waxy solid  The reaction of Acetylene with water in H <sub>2</sub> SO <sub>4</sub> and HgSO <sub>4</sub> yields: (a) Vinyl alcohol (b) Acetaldehyde (c) Mixture of both a and b (d) None  51. The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  52. The Alkynides are used for the	49.	S.=11	e navi	ing Cai	rbon at	toms =	:15 IS I	71.5		-			
(a) Vinyl alcohol (c) Mixture of both a and b (d) None  51. The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  52. The Alkynides are used for the								24	5 6 8 <sup>5</sup>	N.528 95			
(c) Mixture of both a and b  The gas used for illumination: (a) Methane (b) Ethene (c) Ethyne (d) None  The Alkynides are used for the	50.				lene w	ith wa	ater in						
51. The gas used for illumination: (a) Methane (c) Ethyne (d) None  52. The Alkynides are used for the					nd b			2 .JZ(		aenyae			
(c) Ethyne  (d) None  The Alkynides are used for the	51.					n:		(4)					
52. The Alkynides are used for the								- 1 - 1 - 1 - 1 - 1 - 1 - 1 - 1 - 1 - 1		е			
(a) Preparation (b) Purification (c) Separation (d) All of above  53. Formula of chloroform is:  (a) CH <sub>3</sub> Cl (c) CHCl <sub>3</sub> (b) CH <sub>2</sub> Cl <sub>2</sub> (d) CCl <sub>4</sub> ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10  a d b b a d c d d a  11 12 13 14 15 16 17 18 19 20  b a b c d d d d c b  21 22 23 24 25 26 27 28 29 30  b c c d d b a a a d c  31 32 33 34 35 36 37 38 39 40  b c d a c b c d d  41 42 43 44 45 46 47 48 49 50  d c c d b b c a c c  51 52 53	<b>52</b> .	(74) (187) (196)		are us	ed for	the							
53. Formula of chloroform is:  (a) CH <sub>3</sub> Cl (c) CHCl <sub>3</sub> (b) CH <sub>2</sub> Cl <sub>2</sub> (d) CCl <sub>4</sub> ANSWER KEY   1. 2. 3. 4. 5. 6. 7. 8. 9. 10 a d b b d d c d d a 11 12 13 14 15 16 17 18 19 20 b a b c d d d d c b 21 22 23 24 25 26 27 28 29 30 b c c d d b a a d c 31 32 33 34 35 36 37 38 39 40 b c d a c b c c d c 41 42 43 44 45 46 47 48 49 50 d c c d b b c a c c 51 52 53	2	(a) Prepar	ration				•		Purific				
(a) CH <sub>3</sub> Cl (c) CHCl <sub>3</sub> (b) CH <sub>2</sub> Cl <sub>2</sub> (d) CCl <sub>4</sub> ANSWER KEY  ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10  a d b b a d c d d a  11 12 13 14 15 16 17 18 19 20  b a b c d d d d c b  21 22 23 24 25 26 27 28 29 30  b c d d b a a d c  31 32 33 34 35 36 37 38 39 40  b c d a c b c c d c  41 42 43 44 45 46 47 48 49 50  d c c d b b c a c c  51 52 53	EO	(M) (M) (M)		orof	m le-			(d)	All of a	above			
ANSWER KEY  1. 2. 3. 4. 5. 6. 7. 8. 9. 10 a d b b d d c d d a 11 12 13 14 15 16 17 18 19 20 b a b c d d d d c b 21 22 23 24 25 26 27 28 29 30 b c c d d b a a d c 31 32 33 34 35 36 37 38 39 40 b c d a c b c c d c 41 42 43 44 45 46 47 48 49 50 d c c d b b c a c c 51 52 53	53.		от спіс	orotori	m is:			(b)	CH <sub>2</sub> Cl <sub>2</sub>				
1.       2.       3.       4.       5.       6.       7.       8.       9.       10         a       d       b       d       c       d       d       a         11       12       13       14       15       16       17       18       19       20         b       a       b       c       d       d       d       d       c       b         21       22       23       24       25       26       27       28       29       30         b       c       c       d       d       a       a       d       c         31       32       33       34       35       36       37       38       39       40         b       c       d       a       c       b       c       c       d       c         41       42       43       44       45       46       47       48       49       50         d       c       c       d       b       c       a       c       c         51       52       53										MO			
1.       2.       3.       4.       5.       6.       7.       8.       9.       10         a       d       b       d       c       d       d       a         11       12       13       14       15       16       17       18       19       20         b       a       b       c       d       d       d       d       c       b         21       22       23       24       25       26       27       28       29       30         b       c       c       d       d       a       a       d       c         31       32       33       34       35       36       37       38       39       40         b       c       d       a       c       c       c       d       c         41       42       43       44       45       46       47       48       49       50         d       c       c       d       b       c       a       c       c         51       52       53						7	11011	1/-	W 1	0)			
a       d       b       d       c       d       d       a         11       12       13       14       15       16       17       18       19       20         b       a       b       c       d       d       d       d       c       b         21       22       23       24       25       26       27       28       29       30         b       c       c       d       d       b       a       a       d       c         31       32       33       34       35       36       37       38       39       40         b       c       d       a       c       c       c       d       c         41       42       43       44       45       46       47       48       49       50         d       c       c       d       b       c       a       c       c         51       52       53						<u>A</u>	<u> </u>	<u>FK KE</u>					
a       d       b       d       c       d       d       a         11       12       13       14       15       16       17       18       19       20         b       a       b       c       d       d       d       d       c       b         21       22       23       24       25       26       27       28       29       30         b       c       c       d       d       b       a       a       d       c         31       32       33       34       35       36       37       38       39       40         b       c       d       a       c       c       c       d       c         41       42       43       44       45       46       47       48       49       50         d       c       c       d       b       c       a       c       c         51       52       53			4	2	2	1		r (2)	7	0	0	10	Ī
11       12       13       14       15       16       17       18       19       20         b       a       b       c       d       d       d       c       b         21       22       23       24       25       26       27       28       29       30         b       c       c       d       d       b       a       a       d       c         31       32       33       34       35       36       37       38       39       40         b       c       d       a       c       b       c       c       d       c         41       42       43       44       45       46       47       48       49       50         d       c       c       d       b       b       c       a       c       c         51       52       53		,	<b>т.</b>			(0)	(0)	V .	<b>7.</b>				-
21       22       23       24       25       26       27       28       29       30         b       c       c       d       d       b       a       a       d       c         31       32       33       34       35       36       37       38       39       40         b       c       d       a       c       b       c       c       d       c         41       42       43       44       45       46       47       48       49       50         d       c       c       d       b       b       c       a       c       c         51       52       53		,		- 33	120	<	10	122	17			r	;
b       c       c       d       d       b       a       a       d       c         31       32       33       34       35       36       37       38       39       40         b       c       d       a       c       c       c       d       c         41       42       43       44       45       46       47       48       49       50         d       c       c       d       c       a       c       c         51       52       53			b	a	b	(E)	d	d	d	d	С	b	
31     32     33     34     35     36     37     38     39     40       b     c     d     a     c     b     c     c     d     c       41     42     43     44     45     46     47     48     49     50       d     c     c     d     b     b     c     a     c     c       51     52     53		,	21	22	23	24	25		27	28	29	30	<u> </u>
b         c         d         a         c         b         c         c         d         c           41         42         43         44         45         46         47         48         49         50           d         c         c         a         c         c           51         52         53		5		C <	15C								i i
41       42       43       44       45       46       47       48       49       50         d       c       c       d       b       c       a       c       c         51       52       53					11/25	V.				4 \ 1			1
d c c d b b c a c c  51 52 53							<del></del>			A Table	===	+	-}
			d	С		d	b		С				
a d c pakcity.org			51	52	53				100				
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							oakci	ty.or	9				



		IIABTEB-6	•		
		HAPTER 9			
	AROMATI	C HYDRO	CARONS * pa	kcity.org	<b>*</b>
			<b>3</b>		
1.	Most common reactions of benze	ne and its deriv	atives are:		
	(a) Electrophilic addition reaction	• •	Electrophilic substitu		
	(c) Nucleophilic addition reaction		Nucleophilic substitu	ution react	ions
2.	Benzene + Ozone →Y, in this seq		Danzona diazanida		
	<ul><li>(a) Benzene monozonide</li><li>(c) Benzene triozonide</li></ul>	(b) (d)	Benzene diozonide Succinic acid		
3.	Which species represents the ele	X 2			
4 <del>-1</del> 42	(a) NO <sub>2</sub> (b) NO <sub>2</sub>	ସ 📅 .୫.୯୫.	NO <sup>+</sup>	(d)	NO <sub>3</sub>
4.	Heating a mixture of Sodium ben	zoate and soda	_	2 2	
	(a) Benzene	(b)	Methane		
	(c) Sodium benzoate	(d)	Calcium benzoate		
5.	Which of the following species part (a) $SO_3^{-2}$ (b) $S_2O_3$	12	pnonation of benzer SO3	i <b>e ring:</b> (d)	SO <sub>2</sub>
6.	The compound prepared by elect	X 2	Cont. Mar C		302
	(a) Acetophenone	(b)	Glyoxal		
	(c) Cyclohexane	(d)	Hexabromo cyclohe	xane	
7.	The term 'Aromatic' was derived	from:	110		
	(a) Greek word (b) Latin	Y Y \	Russian	(d)	English
8.	Which compound was recognized	~ /\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \		7.15	the state of the s
9.	(a) Aniline (b) Phen Which one of the following is not	$\sim$	Benzene matic hydrocarbon:	(d)	Toluene
<b>J.</b>	(a) Benzaldehyde	(b)	Benzoic acid		
	(c) Benzene sulphonic acid	(d)	Anthracene		
10.	In which one of the following con	pound rings are	Service of the servic	at ortho	positions
	(a) Phenanthrene	(b)	Naphthalene		
4 4	(c) Dipheny methane Toluene is called:	(d)	Anthracene		
11.	(a) Hydroxyl benzene (b) Metl	nyl henzene (c)	Ethyl benzene	(d)	None
12.			Edity Denzene	(4)	Hone
	(a) Aryl groups	(b)	Phenyl groups		
	(c) Acyl groups	(d)	Alkyl groups		
13.		را القوصوران ويدعونان ال		.2. W	
4.4	(a) 1824 (b) 1825		1826	(d)	1827
14.	The empirical formula of Benzeno	U CIN GILY A GILG	138		
	<ul><li>(a) IR spectra</li><li>(c) Elemental analysis</li></ul>	(b) (d)	U.V NMR spectra		
15.	How many molecules of chlorine		•	sunlight:	
ar escendist	(a) One (b) Two		Three	(d)	Four
16.	The bond angles in benzene ring	are:		27 0290.	
	(a) $90^{\circ}$ (b) $120^{\circ}$	X - 7	145 <sup>0</sup>	(d)	None
17.	All C-H bond lengths of benzene	5 W 5	1 0040	( <del></del> )	None
18.	(a) 1.07A° (b) 1.09  A six membered ring containing (	<b>\</b>	1.08A <sup>o</sup> l is called:	(d)	None
10.	(a) Cyclohexene (b) Cycl		Benzene	(d)	None
19.	Hybridization of each carbon ato	Loren Loren Some Unit		(4)	INCLIC
	(a) sp (b) $sp^2$	(c)	sp <sup>3</sup>	(d)	dsp <sup>2</sup>

20.	The stability	, of aromatic	compounds	wit	th the increase	in the numl	ner of its
20.	resonance st		compounds _		in the increase	in the name	or or its
	(a) Decre	eases		(b)	Increases		
	· /	ains constant		(d)	partially decreas		
21.		22 전(	2	2 45.	by treating it w	2 550	
22	(a) HCl	(b)	NaOH	(c)	H <sub>2</sub> O	(d)	NaHCO:
22.	20 Ta	ropnilic substit	ution reaction	n is too vi (b)	gorous to contro Bromination	OI:	
	(a) Chlor (c) Iodina			(d)	Fluorination		
23.	A service serv	id generates n	itronium ion	` '			
	(a) Nitric			(b)	Nitrogen gas		
		us acid		(d)	Potassium nitrat	:e	
24.	100 mm m	cts with ozone	and gives:	AV 30			
	(a) Glyce			(b)	Glyoxal		
25		c anhydride	vidized by se	(d)	Benzoic acid		
25.	(a) KMnC	ne are readily on the control of the	K <sub>2</sub> CO <sub>3</sub>	(c)	MnO <sub>2</sub>	(d)	Mn
26.	\ /				gnificant and it r	X 2	
		rophilic		(b)	Nucleophilic		<i>g g</i> · · ·
	(c) Nucle	eophobic		(d)	Hydrophobic		
27.	Meta direction	ng groups dec	rease the	of ber	zene ring:		
	AND REAL RAY (MADE) PAGE	ng point (b)			(c) Density	(d)	None
28.	56 SV 2006	of compound i	s most reacti	20.5			
	(a) Alkan			(b)	Alkene		
29.	(c) Alkyn		d can be used	as a cata	) Benzene <b>lyst in Friedel C</b> i	raft's reactio	nc
23.	(a) AlCl <sub>3</sub>	(b)	HNO <sub>3</sub>	(c)	BeCl <sub>2</sub>	(d)	NaCl
30.	N Z	drocarbons are				(-)	
		al series of para	NAIO		Benzene	(d)	None
31.		f the following	VIIV	e):	a anno constant anno L. L.		
		imethyl benzene		(b)	1,3-Dimethyl be		
22	<b>`</b>	imethyl benzend		(d)	1,4-Dimethyl be		na UNIÒ
32.		SO <sub>4</sub> at 50°C in		men it is	heated with a n	nixture or co	nc. HNO
	(a) 1:2	** ** ** ** ** ** ** ** ** ** ** ** **	1:1	(c)	1:3	(d)	2:1
33.	<b>\</b>	( )			e benzene ring	X /	
	(a) Conju	ugate bonds		(b)	Resonating bond	ds	
(I	( - )	a and b		(d)	None of above		
34.			amount of he	eat actua	lly released an	d the exper	imentally
		eat is called: energy		(b)	Reconance oner	αv	
	Man the control of th	ng energy		(b)	Resonance ener None	9)	
35.	S 2	& Para directi	na aroups ex		140110		
2000 Editor	(a) -X	(b)	-OH	(c)	-NR₃	(d)	-NH <sub>2</sub>
36.	Substitution	of halogens in	the benzene	ring requ	ires which cata		
	(a) NaCl			(b)	FeCl <sub>3</sub>		
<b>~</b> =	(c) $SiO_2$	£41.		(d)	Organo - nickel		
37.	20.00	f the following	Aller -		52-45	/4/	ΛII
38.	(a) Alken		Alkyne n-Hexane in t	(C) he preser	Benzene nce of which cat	(d) alvet?	All
JO.	(a) $Cr_2O_3$	# <del>#</del>	ii iiexalle III (	(b)	Al <sub>2</sub> O <sub>3</sub>	aryst:	
	(c) $SiO_2$			(d)	Fused mixture o	f a, b and c	
				77	3 - 10 - 10 - 10 - 10 - 10 - 10 - 10 - 1		

39.		ene cann					/L)	Λ.	م در د ۱:۱:۱ اد				
	(a) (c)	Substitu Oxidation					(b) (d)		dditiona Iiminatio				
40.		, para de			obtai	ned by				JII I Eac	LIOIT	(LHR	2011)
	(a)	Nitrobe					(b)		oluene			ζ	,
	(c)	Benzald	ehyde				(d)	В	enzene				
41.	0= =1	ene is pro		from	Cycloh	exane						(LHF	R 2008)
	(a)	Hydroge Dehydro		nn -			(b) (d)		ehydrat one of t				
42.	(c) <b>The c</b>	onversio	4400		e into	benze	` '				ence o	f Pt is	called:
	(a)	Isomeri				201120	(b)		romatiz	Section 1			
	(c)	Dealkyla					(d)		earrang	ement			
43.	2 20	n compoi		221 40				=	·			/ D	H11
<b>11</b>	(a)	Benzene is the m		(b)	Ethei		(c)	Et	thane			(d)	Ethyne
44.	(a)	C <sub>6</sub> H <sub>2</sub> (NC			iuia oi	INI	(b)	C	5H₂(NO2	)-CH2			
	(c)	C <sub>6</sub> H <sub>2</sub> (NC	)53				(d)		5H2(NO2	5			
45.		cule of be	E-200 (10)		in		(-)			.,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,	(G	RW 20:	11,14)
	(a)	Three d					(b)		wo doul			Nett.	
	(c)	One do			st and <b>1</b> % output topics of the	na 🎒 an	(d)	D	elocaliz	ed л-ele	ectron	charge	
46.	2 3	tion of To		takes	place a	at:	(h)	n	nocitio				
	(a) (c)	m-position of the m-position o					(b) (d)		-position oth 'o' a	/	osition	ıs	
47.		nance en		f benz	ene is:		(4)			a p p	703161011		2012)
	(a)	150.5 K					(b)	3314	40.5 KJ	/mol			•
2.22	(c)	155 KJ/		9		$\wedge$	1 (C)		45 KJ/m	ol			
48.		ich com	•		ne ring	js are	1 )		anhthal	000			
	(a) (c)	Dipheny Anthrac		ane		(6)	(b) (d)		aphthal henanth				
49.	Prese	nce of do		or triple	e bond	is a si				irene		(SGD	2011)
	(a)	Un-satu		~~	M		(b)		aturatio	n		<u> </u>	· •
	(c)	Addition		Was	7/12		(d)		ubstitut			_# 74	_ 2
50.		ene react 2008)	ts with	анку	and a	cyi na	lides ii	i the p	oresen	ce of A	MCI3.	inis re	action is:
	(a)	Freidel	Crafts re	eaction			(b)	Al	dol con	densati	on		
	(c)	Haloger					(d)		itration				
					A.A	MCM	CD VC	V		a nak	city.org	- Dec	
					A	MOM	ER KE			- Man	city.org	- Base	
		1.	2.	3.	4.	pa <sub>5</sub> .ci	6.	7.	8.	9.	10.		
		b	C	b	a	С	a	a	С	d	С	_	
		11	12	13	14	15	16	17	18	19	20	-	
		b	a	b	C	C	b	b	a	b	b		
		21	22	23	24	25	26	27	28	29	30	<u> </u>	
		C	a	a	D 24	a	D 26	b	D 20	a 20	C 40	-	
		31	32	33	34	35	36	37	38	39	40		
		D	D 42	(C)	D 44	C 4F	D 46	C 47	40	40	D <b>FO</b>	_	
		41	42	43	44	45	46	47	48	49	50	-	
			D	ען	a	a	<sub>l</sub> u	a	l a	a	a	_	



				الضمالضا					
1.	2.	3.	4.	5.CI	6.	7.	8.	9.	10.
b	С	b	a	С	a	а	С	d	С
11	12	13	14	15	16	17	18	19	20
b	а	b	С	С	b	b	a	b	b
21	22	23	24	25	26	27	28	29	30
С	d	a	b	a	b	b	b	а	С
31	32	33	34	35	36	37	38	39	40
b	b	С	b	С	b	С	d	d	b
41	42	43	44	45	46	47	48	49	50
С	b	b	а	d	d	а	a	a	а

		CHAPTE	R 1			
		ALKYL HA	AI TD	=S  pakcity.org		
		<u> </u>				
1.	Electrophile amongst the fol	lowing is:				
2.	(a) NH <sub>3</sub> (b) SN <sub>2</sub> mechanism involves:	H <sub>2</sub> O	(c)	BF <sub>3</sub>	(d)	Cl <sub>2</sub>
	(a) 1st order kinetics		(b)	2nd order kinetics		
2	(c) 3rd order kinetics	anaud'a roado	(d)	Zero order kinetics		
3.	Alkyl magnesium halide (Gri (a) Alkane (b)	gnard's reage Alkene	nt) wn (c)	Alkyne	(d)	Alkyl
	halide				( )	, , , , , , , , , , , , , , , , , , ,
4.	Tertiary alcohols are obtained (a) Aldehyde (b)	ed by treating Ketone	Grigna (c)	rd's reagent with: Water	(d)	Amine
5.	Alkyl Halides are:	Recorre	(0)	Water	(4)	Author
6.	(a) Monohaloalkanes (b) Which substance is used to		(c)	Polyhaloalkanes	(d)	All
0.	(a) H <sub>2</sub> O (b)	NH <sub>3</sub>	(c)	Ethyl alcohol	(d)	All of
_	these				C 1 13/2 !	
7.	For the reaction C <sub>2</sub> H <sub>5</sub> OH+ H <sub>2</sub> (a) HBr > HI > HCl	$X \rightarrow C_2H_5-X + 1$	H <sub>2</sub> O the	HI > HCl > HBr	r HX IS:	
	(c) $HCl > HBr > Hl$		(a)\S	HI > HBr > HCl		
8.	Carbanions are:	Nucleanbitas	S (a)	Crown of atoms	(4)	Fron
	(a) Electrophiles (b) radical	Nucleophiles	(c)	Group of atoms	(d)	Free
9.	Secondary alkyl halides follo	WIDE	gar an			
	<ul><li>(a) First order kinetics</li><li>(c) Both a and b</li></ul>	The	(b) (d)	Second order kinetics none of these		
10.	(CH <sub>3</sub> ) <sub>3</sub> CBr preferably underg	oes:	(u)	Horic of these		
	(a) SN <sub>2</sub> reactions		(b)	SN <sub>1</sub> reactions		
11.	(c) both a and b  Ethyl chloride reacts with ale	coholic KOH to	(d)	none of these		
7. ak ac	(a) C <sub>2</sub> H <sub>4</sub> (b)	C <sub>2</sub> H <sub>5</sub> OH	(c)	C <sub>2</sub> H <sub>6</sub>	(d)	None of
10	these	th soci- an	alled b	alide is formed what		wa athaw
12.	When an alcohol reacts wi products:	th SOCI2 and	акуі п	alide is formed wha	t are t	wo other
	(a) SO <sub>2</sub> and HCl	nakaitu	(b)	SO <sub>2</sub> and H <sub>2</sub> O		
13.	(c) HCl and H <sub>2</sub> O Which of the following is not	t a nulceophile	(d)	H <sub>2</sub> S and HCl		
10.	(a) OH (b)	NH <sub>3</sub>	(c)	$C_2H_5O^-$	(d)	Br <sup>+</sup>
14.	Which of the following react				92 - 249 	Tlinain - Li
15.	(a) SN <sub>1</sub> (b)  In primary alkyl halides, th	SN <sub>2</sub> e halogen ato	(c) om is a	Addition ttached to a carbon		Elimination is further
	attached to:	<b>_</b>				
	(a) One Carbon atoms		(q)	Two Carbon atoms		
16.	(c) Three Carbon atoms Reaction of following with G	rignard's reag	(d) <b>jent ca</b> i	Four Carbon atoms n give primary alcoho	l: (LHF	R 2011,14)
-	(a) Epoxide (b)	Peroxide	(c)	Super oxide	7.000	Hydrogen

<b>17.</b>	S <sub>N</sub> 1 rea	actions are easily					(LHR	2012)
	(a)	Primary alkyl halid			(b)	Secondary alkyl halide		
(24 (244))	(c)	Secondary alcohols			(d)	Tertiary alkyl halides		
18.		l chloride reacts v					. IS	■27 <b>3</b> 1 <b>3</b> 1 <b>3</b> 2
	(a)	Mustard gas (b	58	halide	(c)	Aldehyde	(d)	Alcohol
19.	32 (2)	of the following is		124		No. Db	(4)	No Db
20	(a)	NaPb (b			(C)	Na <sub>3</sub> Pb	(d)	Na₄Pb
20.			mea wnei	n Grignare		gent reacts with:		
	(a) (c)	Propanone Ethanal			(b)	Methanal hanoic acid		
21.	7 March 1972	Grignard reagent	reacts wi	th Fnoxid				
<b>4</b> 1.	(a)	1- Alkanal (b		kanol	(c)	Carboxylic acid	(d)	None
22.		THE REAL PROPERTY AND ADDRESS OF THE PERSON ADDRESS OF THE PERSON AND ADDRESS OF THE PERSON AND ADDRESS OF THE PERSON ADDRESS OF THE PERSO				preparing alkyl chloric		
	(a)	HCl+ anhydrous Zi		NaCl	(c)	PCl <sub>5</sub>	(d)	SOCl <sub>2</sub>
23.	. ,	atom holding ha					(-)	_
	(a)	sp <sup>2</sup> - hybridized		54.57	(b)	sp <sup>3</sup> -hybridized		
	(c)	sp-hybridized			(d)	sp <sup>3</sup> d- hybridized		
24.	Which	of the following of	loes not g	ive Iodof	orm te	est:		
	(a)	Ethanol			(b)	Ethanal		
	(c)	Acetophenone			(d)	Benzophenone		
25.	C-X bo	nd is strongest in				6		
2.2	(a)	CH₃Cl (b	<b>5</b> 8		(c)	CH <sub>3</sub> F?	(d)	CH₃I
26.	Fig. 1987	yl halide is conve	erted into	an alcoho		1 1		
	(a)	Addition			(p)	Substitution		
~=	(c)	Dehydrohalogenat			(g)/?	Elimination		
27.	12: 72:	hane reacts with	20 (=5)0 2	1.0	/.	(ES) 1	(4)	Dutana
20	(a)	Ethane (b			(c)	Butene	(d)	Butane
28.	-	Dibromopropane (h		7010	77.5		(d)	Hoyano
29.	(a) Ethyla	Propene (b Icohol gives Ethyl	( )		(c)	Cyclopropane F•	(d)	Hexane
<b>2</b> 3.	(a)	SOCl <sub>2</sub> (b	VAIN		(c)	Cl <sub>2</sub>	(d)	KCl
30.		nitrile is formed	VIV.)				(u)	KCI
30.	(a)	Propyl alcohol (b		l chloride		Butyl alcohol	(d)	Propyl
	Chloride	Tropyr dicorior \db	Ducy	Cilionae	(0)	Bueyr dicorior	(u)	Поруг
31.		omoethane on tre	eatment w	ith alcoh	olic zir	nc gives:		
	(a)	Ethyl bromide (b			(c)	Ethane	(d)	Ethyne
32.		action of Alkyl hal	The state of the s				<b>.</b>	er no es esta e secon
	(a)	Retention of config	guration		(b)	Inversion of configurati	on	
	(c)	Both a and b			(d)	None of these		
33.	Which	one of the follow	ing will ha	ave the m	aximu	m dipole moment:		
	(a)	CH₃F (b			(c)	CH₃Br	(d)	CH₃I
34.		action of an Alkyl	halide wi	th RCOOA				
	(a)	Ester			(b)	Ether		
<b>~</b> =	(c)	Aldehyde	7	. l	(d)	Carboxylic acid	/===	2012
35.		strate of a typical SN		n IS:	/L\	Cooperations - Head to - 1941	(FSD	2012)
	(a)	Primary alkyl halid			(q)	Secondary alkyl halide		
26	(c)	Secondary alcohols		c C1 =0=	(d)	Tertiary alkyl halides		
36.	-	eactive Alkyl halid	ie roward	2 2NT LEGG				
	(a)	n-Butyl chloride ter-Butyl chloride			(b) (d)	Sec-Butyl chloride Allyl chloride		
37.	(c) Which	responds +vely to	nwarde To	doform to		Allyl Cilloride		
J/.	(a)	1-Butene	54441U3 1U	aoioiiii U	(b)	Butanal		
	(a) (c)	Acetic acid			(d)	2-Pentanone		
	(5)	Accur dela			(4)	_ i critarione		

		<b>*</b> /**				•							
38.	Electro (a)	phile ai NH3	mong t	t <b>he fol</b> l (b)	lowing H <sub>2</sub> O	IS:	(c)	В	F <sub>3</sub>			(d)	Cl <sub>2</sub>
39.	The ord	ler of a	typica			n is:	(=)	Σ,	. 3			(4)	O.Z
	(FSD 20 (a)	<b>10)</b> Zero		(b)	First		(c)	S	econd			(d)	Third
40.	Which	of the f	ollowi		ot a N								
41.	(a) <b>In whic</b>	H <sub>2</sub> O	mocha	(b)	H <sub>2</sub> S	ret eta	(c) Lovel		F <sub>3</sub>			(d)	NH <sub>3</sub>
71.		E1 and		11131113,	, the m	اعد عدد	(b)		2 and S				
	(c)	E1 and					(d)		v1 and S				
42.	SN <sub>2</sub> rea						(h)	D	imalacu	lar			
		Unimole Trimole					(b) (d)		imolecu etramol				
43.	Èlimina	tion bi	molecu	ılar rea	action	involv	, ,	etics:				(LHR	2014)
		1 <sup>st</sup> orde					(b)		nd order				
44.		3 <sup>rd</sup> orde		lides	the h	alogen	(d) atom		ero orde hed to		rhon a	atom i	is further
- 11	attache		<b>□</b> 0			83.77	4.VIII			u		II I	
	(a)	Two					(b)		hree	\			
45.	(c) <b>Elimina</b>	One	moleci	ılar re:	action	involv	(d)	F	our				
#YI		First ord			action	IIIVOIV	(b)	\(\sigma\)	econd o	rder kir	netics	® nak	city.org
	(c)	Third or				- 128 - 128	(d)	1 2	ero orde	er kinet	ics	∞ hav	city.org
46.	Ethyl m	,, Res	um bro	omide	reacts	with v	1(5)		<b>1:</b> ethane				
	X 2	Ethane Propane	2			(P)	(b)		-butane				
47.	Primar	y alcoh	ol is ob	otaineo	d by tre	eating		rd's R	egent	with:			
40	\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	HCHO	alama L	(b)	EH3C		(c)		H₃COC⊦			(d)	CO <sub>2</sub>
48.	<b>In SN</b> <sub>2</sub> (a)	sp	ilisili, t	ile liya	riuiza	LIOII OI	(b)	Sp	2	jes II o	1111.	(r	RWP 2009)
	(c)	dsp <sup>2</sup>	<	W,			(d)	n	one of t				
49.	Tertiary	y <b>alcoh</b> HCHO	ol is ob	commence of the					<b>igent v</b> H₃COCH			0822201120	RWP 2009)
50.	(a) <b>Order a</b>		lecular	(b) rity of S	CH₃C SN2 re		of alk			13		(D) <b>(F</b>	CO <sub>2</sub> RWP 2011)
	(a)	1,2		(b)	2,1		(c)		.2			(d) `	0,1
						NICHA	FD I/F						
					A	NSW	EK KE	<u>Y</u>					
		1.	2.	3.	4.	78 <b>5.</b> Ci	.y <sub>6</sub> .	7.	8.	9.	10.		
		С	b	a	b	a	d	d	b	C	b	_	
		11	12	13	14	15	16	<b>17</b>	18	19	20	1	
		21	22	23	24	2 <b>5</b>	2 <b>6</b>	27	28	<b>29</b>	30	1	
		b	b	b	d	C	b	d		a	d	<u> </u>	
		31	32	33	34	35	36	37	38	39	40		
		d	С	a	a	а	С	d	С	С	С		
		41	42	43	44	45	46	47	48	49	50		
		С	b	b	C	b	a	a	d	С	C	_	

22				أخساحت	بلد حد معال	- / 10		4	
1.	2.	3.	4.	5.	6.	7.	8.	9.	10.
С	b	a	b	a	d	d	b	С	b
11	12	13	14	15	16	17	18	19	20
a	a	d	С	a	a	d	b	b	С
21	22	23	24	25	26	27	28	29	30
b	b	b	d	С	b	d	С	а	d
31	32	33	34	35	36	37	38	39	40
d	С	а	а	а	С	d	С	С	С
41	42	43	44	45	46	47	48	49	50
С	b	b	С	b	a	a	d	С	С

## **CHAPTER 11**

# ALCOHOLS, PHENOLS AND ETHERS & pakcity.org



	ALCOHOLS.	PHENOLS AN	ID ETHERS	- <b>₩</b> pakci
4				Og.
1.	Which compound will have m (a) C <sub>6</sub> H <sub>6</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> OH	(b)		c)
2.	When Ether is protonated, the (a) A Carbonion	(b)	Oxide ion	
3.	(c) An Oxonium ion  When ethyl bromide is heated		A Hydronium ion luct formed is:	
	(a) Ethanol (c) Ethanal		) Ethene ) Diethyl ether	
4.	When diethyl ether is treated		· · · · · · · · · · · · · · · · · · ·	
5.	(a) Ethanol (c) Ethyl chloride  2-Alkanol on oxidation forms	(d)	) Triethyl phosphin ) Oxonium ion	e
<b>J</b> .	(a) Ether	(b)	) Aldehyde	
6.	(c) Ketone Which substance is used as a	` '	) Carboxylic acid	
0.	(a) Methanol (c) Acetone	(b)	Ethanol Acetic acid	
7.	Which of the following would		(0)	tate of C
	heated with I <sub>2</sub> in alkaline solution (a) CH <sub>3</sub> OH	05 ( )	○ )C₃H <sub>7</sub> OH	
	(c) CH <sub>3</sub> CH <sub>2</sub> CH(OH)CH <sub>2</sub> CH <sub>3</sub>		C3H7OH CH3CH2OH	
8.	Alcohols on heating with		<b>■</b> 11	
	(a) Oxygen (c) PCl₃	( ) )	) Conc. H <sub>2</sub> SO <sub>4</sub> ) Conc. HNO <sub>3</sub>	
9.	Carbolic acid is the other nan	ne of:		
	(a) Phenol (c) Picric acid		) Biphenyl ) H₂CO₃	
10.		oduce esters:		
	(a) Water (c) Carboxylic acid		) Ethers ) Ketones	
11.	Oxidation ofalcohol give		Retories	
	(a) Primary		Secondary	
12.	(c) Tertiary Which compound is more solu	The same of the sa	Quaternary	
	(a) C <sub>2</sub> H <sub>5</sub> OH		CH₃COCH₃	
13.	(c) CH <sub>3</sub> OH Which compound shows max		) C <sub>6</sub> H₅OH ding with water?	•
	(a) CH <sub>3</sub> -O-CH <sub>3</sub>	(b)	) CH₃-OH	
1/1	(c) C <sub>2</sub> H <sub>5</sub> -OH Which enzyme is not involved		) C <sub>6</sub> H <sub>5</sub> -OH	
17.	(a) Diastase		) Maltase	
4=	(c) Zymase	(d)	) Urease	
15.	Methyl alcohol is used as: (a) a drink	(b	an anti freezing a	aent
	(c) a preservative for biological s	pecimen (d)	a substitute for p	etrol
16.	Only maximum% alcoho			asses:
	(a) 14 (c) 95		) 24 ) 100	

17.	Ethene is formed when Ethyl alcohols reacts	
	(a) 180°C (c) 120°C	(b) 150°C (d) 110°C
18.	Dow's method helps in preparation of:	(d) 110°C
10.	(a) Biphenyl	(b) Benzene
	(c) Phenol	(d) Ester
19.	When phenol react with alkali it forms:	(d) Little
	(a) Acid	(b) Salt
	(c) Base	(d) None of these
20.	Phenol reacts with zinc dust and forms:	(4)
	(a) Benzene	(b) Acetylene
	(c) Molozonide	(d) Ozonide
21.	Bakelite is formed when phenol reacts with	•
	(a) Acetaldehyde	(b) Farmaldehyde
	(c) Propanoic acid	(d) Butanoic acid
22.	Ethers reacts with hydrogen iodide to give:	
	(a) Alcohol	(b) Phenol
	(c) Alcohol+ Alkyl halide	(d) Aldehyde
23.	Which one is primary alcohol:	
	(a) Buten-2-ol	(b) Propan-2-ol
24	(c) Butan-1-ol	(d) 2, 3- Dimethylhexane-4-ol
24.	Ethyl alcohol is industrially prepared from e	
	(a) Permanganate oxidation (c) Absorbing in U.CO. followed by bydrolygic	(b) Catalytic reduction
25	(c) Absorbing in H <sub>2</sub> SO <sub>4</sub> followed by hydrolysis	(d) Fermentation
25.	Ethanol containing some quantity of metha	
	<ul><li>(a) Absolute spirit</li><li>(c) Wood spirit</li></ul>	(b) Rectified spirit (d) Methylated spirit
26.	Hydrolytic conversion of Sucrose into gluco	
20.	(a) Induction	(b) Inversion
	(c) Insertion	(d) Inversion  (d) Inhibition
27.	Alcohols of low molecular weight are:	(d) Inhibition
	(a) Soluble in water	(b) Soluble in water on heating
	(c) Insoluble in water	(d) Insoluble in all solvents
28.	Fermentation is a:	
	(a) Chemical process	(b) Biochemical process
	(c) Engineering process	(d) Physical Process
29.	Ethyl alcohol on oxidation with acidified K20	Cr <sub>2</sub> O <sub>7</sub> gives:
	(a) Acetic Acid	(b) Acetaldehyde
	(c) Formaldehyde	(d) Formic acid
30.	C <sub>2</sub> H <sub>5</sub> OH can be differentiated from CH <sub>3</sub> OH b	· · · · · · · · · · · · · · · · · · ·
	(a) Lucas Test	(b) Baeyer's Test
, <u> </u>	(c) Iodoform test	(d) None
31.	Ethyl alcohol on treating with Conc. H <sub>2</sub> SO <sub>4</sub> a	Add at the second of the secon
	(a) Ethene	(b) Diethyl Ether
22	(c) Ethyl acetate	(d) Ethanoic acid
32.	Which of the following statements is correct	
	(a) Phenol is less acidic than Ethanol	(b) Phenol is more acidic than Ethanol
22	(c) Phenol is more acidic than Acetic acid	(d) None
33.	Isopropyl alcohol on oxidation gives:	(b) Ethor
	(a) Acetone	(b) Ether (d) Acetaldebyde
2/	(c) Ethylene Which of the following is the most suitable of	(d) Acetaldehyde  nethod for removing the traces of water
34.	Which of the following is the most suitable in from ethanol:	nethod for removing the traces of water
		(b) Passing dry HCl through it
	(a) Reacting with Na metal (c) Distilling in presence of CaO	(d) Reacting with Mg

35.	Na reacts with Ethanol to produce:
	(a) H <sub>2</sub> gas (b) Benzene
	(c) CO <sub>2</sub> gas (d) CO gas
36.	
	(a) 2, 4, 6- Trinitrotoluene (b) 2, 4, 6-Tribromethanol
27	(c) 2, 4, 6-Trinitrophenol (d) Para-Nitrophenol
37.	
	(a) Ascorbic acid (b) Vinegar (c) Oxalic acid (d) Picric acid
38.	Methanol can be prepared from hydrogenation of: (LHR 2011)
	(a) CH₃CN (b) CH₃Br
	(c) HCHO (d) CH₃CHO
39.	Phenol reacts with acetyl chloride in the presence of a base to form an: (LHR 2008)
	(a) acid (b) alcohol (c) aldehyde (d)
	ester
40.	Alcohol obtained by fermentation never exceeds: (LHR 2014)
	(a) 10% (b) 16% (c) 14% pleabol (d) 05% pleabol
41.	(c) 14% alcohol (d) 95% alcohol  Phonol reacts with acetyl chloride in the presence of a base to form an:
41.	Phenol reacts with acetyl chloride in the presence of a base to form an: pakcity.or  (a) Acid  (b) Alcohol
	(c) Aldehyde (d) Ester
42.	2 – Hydroxy propanoic acid is called:
<u> </u>	(a) Oxalic acid (b) Lactic acid
	(c) Citric acid (d) Aspartic acid
43.	Methyl alcohol is not used: (GRW 2014)
	(a) As a substitute for Ethanol in drinking (b) As a substitute for petrol
NE ES	(c) For denaturing of ethyl alcohol (d) All
44.	Which liquid is called wood spirit? (GRW 2011)
	(a) CH₃OH (b) C₂H₅OH
45	(c) CH <sub>3</sub> COOH (d) CH <sub>3</sub> OCH <sub>3</sub>
45.	Order of reactivity of alcohols when C-O bond breaks is: (LHR 2012)  (a) Tertiary alcohol>Secondary alcohol>Primary alcohol
	(a) Tertiary alcohol>Secondary alcohol>Primary alcohol  (b) Primary alcohol>Secondary alcohol>Tertiary alcohol
	(c) Secondary alcohol>Primary alcohol>Tertiary alcohol
	(d) Tertiary alcohol > Primary alcohol > Secondary alcohol
46.	Which compound is called a universal solvent?
	(a) H <sub>2</sub> O (b) CH <sub>3</sub> OH
	(c) C₂H₅OH (d) CH₃OCH₃
47.	Ethanol and methanol can be distinguished by (MTN 08, FSD 10)
	(a) Iodoform test (b) Lucas Test
40	(c) Benedicts test (d) Tollen's test
48.	(a) Position isomerism (b) Functional group isomerism
	(c) Metamerism (d) Cis-trans isomerism
49.	Which enzyme is not involved in fermentation of starch? (LHR 2014,15)
	(a) Diastase (b) Zymase
	(c) Urease (d) Invertase
<b>50.</b>	Which of the following reagent will replace -OH group of alcohol by halogen atoms?
	(a) HOCl (b) $Br_2$ (c) $SOCl_2$ (d) $I_2$
	ANSWER KEY
	1. 2. 3. 4. 5. 6. 7. 8. 9. 10.
	a c d c c a d b a c
	11 12 13 14 15 16 17 18 19 20
	a c b d b a a b b a
	21 22 23 24 25 26 27 28 29 30
	b c c d d b a b c 31 32 33 34 35 36 37 38 39 40
	b b a c a c d c d c
	41 42 43 44 45 46 47 48 49 50
	d b a a a a c c c

1.	2.	3.	4.	5.	6.	7.	8.	9.	10.
а	С	d	С	С	a	d	b	а	С
11	12	13	14	15	16	17	18	19	20
а	С	b	d	b	a	a	b	b	a
21	22	23	24	25	26	27	28	29	30
b	С	С	d	d	b	a	b	b	С
31	32	33	34	35	36	37	38	39	40
b	b	a	С	a	С	d	С	d	С
41	42	43	44	45	46	47	48	49	50
d	b	а	a	а	a	a	С	С	С



	CHAPT	TER 12
	ALDEHYDES A	AND KETONES Repakcity.org
1.	appear that the transfer of th	Cannizzaro's reaction in the presence of dilute
	aqueous sodium hydroxide?	(b) CU CUO
	(a) HCHO (c) CH <sub>3</sub> -CH <sub>2</sub> -CHO	(b) CH₃CHO (d) CH₃-COCH₃
	(C) CH3-CH2-CHO	(u) CH3-COCH3
2.	Formaldehyde condenses with phenol in th	e presence of dilute base to yield:
	(a) Nylon-6,6	(b) Urotropine
	(c) Aniline-formaldehyde	(d) Bakelite
3.	Calcium acetate on dry heating yields:	(b) CU CUO
	(a) HCHO	(d) CH₃COOH
1	(c) CH <sub>3</sub> COCH <sub>3</sub> Air oxidation of methanol produces:	(d) CH₃COOH
4.	(a) Ethanol	(b) Methanal
	(c) Mixture of Methanal and Ethanal	(d) Methane
5.	Acetone reacts with HCN to form cyanol	× 2
	(a) Electrophilic addition	(b) Electrophilic substitution
	(c) Nucleophilic addition	(d) Nuclearphilic substitution
6.	Which of the following compounds will i	
	(a) Acetaldehyde	(b) Acetone
	(c) Butanone	(d) 3-Pentanone
7.	Addition of alcohol in carbonyl compour	ds gives acetal; the geometry of acetal is:
	(a) Linear	(b) Trigonal
	c) Tetrahedral	(d) Planar
8.	Which of the following compound will re	eact with Tollen's reagent:
	(a) Acetone	(b) Acetic acid
2 <b>2</b> 0	(c) Methyl ethyl ketone	(d) Acetaldehyde
9.	Which of the reagents will react with Ke	
	(a) Sodium nitroprusside	(b) Tollen's reagent
10	(c) Fehling's reagent  The Nuclear billion ddition reactions of a	(d) Benedict's reagent
10.	The Nucleophilic addition reactions of ca (a) Acids	(b) Bases
	(c) Both a and b	(d) None of these
11.		
	(a) Iodoform test	(b) Tollen's reagent
	(c) Silver mirror test	(d) all of these
12.		A War Coll School College
	(a) Formaldehyde	(b) Acetaldehyde
	(c) Menthol	(d) Menthone
<b>13.</b>	Which is used as an antiseptic inhalant:	
	(a) Formaldehyde	(b) Acetaldehyde
	(c) Formic acid	(d) Acetic acid
14.	125 SE 156 SE 157 SE 15	상태 전 ( ) ( ) ( ) ( ) ( ) ( ) ( ) ( ) ( ) (
	(a) Hydrogen Iodide	(b) Hydrogen gas
2 <u>4</u> 1—-	(c) Hydrogen Chloride	(d) Sodium Hydroxide
15.		
	(a) Aldehyde	(b) Alcohol
	(c) Carboxylic acid	(d) Ketone

4.5		
16.	Isopropyl alcohol on oxidation forms:	
	(a) Acetone	(b) Ether
4 -	(c) Ethylene	(d) Acetaldehyde
17.	_	ed for detecting the presence of carbonyl group?
	(a) Reaction with hydroxylamine	(b) Ammonical cuprous oxide
	(c) Ammonical silver bromide	(d) Ammonical silver nitrate
18.	Propyne on hydrolysis in presence of	
	(a) Acetaldehyde	(b) Acetone
10	(c) Formaldehyde	(d) None
19.	On heating acetaldehyde with Ammor	·
	(a) CH₃OH	(b) Silver acetate
20	(c) HCHO	(d) Silver mirror
20.	At room temperature formaldehyde is	
	(a) Gas	(b) Liquid
24	(c) Solid  The compound obtained by the reduce	(d) Rubber like solid
21.		tion of Propionaldehyde with amalgamated Zind
	and concentrated HCl is:	(h) Duamana
	(a) Propanol	(b) Propane
22	(c) Propane	(d) All
22.		portionation reaction in presence of sodium of
		nding alcohol and acid. The reaction is known as
	(a) Wurtz reaction	(b) Cannizzaro reaction
22	(c) Friedel-Craft reaction	(d) None
23.	5	
	(a) Chloral hydrate	(b) Ethanol Tetramer
24	(c) Ethanol Trimer	(d) both a and c
24.		e passed over heated copper, the major product
	obtained is:	(h) Dropylone
	(a) Propane	(b) Propylene
25	(c) Acetaldehyde	(d) Acetone
25.	A Nucleophilic regent will readily atta	
	(a) Ethylene (c) Ethanol	(b) Ethanal (d) Ethylamine
26.		
20.		
	(a) Ethanol (c) Acetone	(b) Ethanal (d) Acetophenone
27.		
۷/.	(a) An acetal	(b) An aldol
	(c) Paraldehyde	
28.	Which of the following does not give I	(d) Acetone
20.	(a) Formalin	(b) Acetaldehyde
	(c) D-Glucose	(d) Acetaidenyde (d) Acetone
29.	Formalin is 40% aqueous solution of:	
<b>2</b> 3.	(a) Furfural	(b) Formaldehyde
	(c) Formic acid	(d) Methyl iodide
30.	Acetone is oxidized with:	(a) Ficulty locate
50.	(a) Tollen's reagent	(b) Fehling solution
	(c) Acidified dichromate solution	(d) Benedicts solution
31.		
JI.	(a) Benzyl alcohol	(b) Hydrobenzamide
	(c) Cinnamic acid	(d) Benzophenone
		N-9

32.	Wolf-Kishner reduction is used for the reduction	ı of.	
	(a) Nitro compounds	(b)	Carboxylic acids
(G-17)	(c) Carbonyl compounds	(d)	Olefins
33.			
	(a) Phenyl hydrazine	(b)	Hydroxylamine
24	(c) Fehling solution	(d)	Sodium bisulphate
34.			
	(a) H <sub>2</sub> with Pd catalyst	(b)	KOH+N <sub>2</sub> H <sub>4</sub>
35.	(c) LiAlH <sub>4</sub> in water Which of the following organic compounds ex	(d) hibits n	Zn-Hg with conc. HCl
<b>33.</b>	Iodoform test:	ilibits b	ositive remining test as well as
	(a) Methanal	(b)	Ethanol
	(c) Propanone	(d)	Ethanal
36.			
	methyl magnesium iodide?	<b></b>	
	(a) HCHO	(b)	CH₃CHO pakcity.org
	(c) CH₃COCH₃	(d)	CO <sub>2</sub> pakcity.org
<b>37.</b>	Bisulphite adduct is:		
	(a) Yellow ppt	(b)	Crystalline white ppt
	(c) Greenish Crystalline	(d)	Red solid
38.			
	(a) 4-Methylisopropy ketone	(b) <	3-Methyl-2-butanone
	(c) Isopropylmethyl ketone	(qp)	✓ 2- Methyl-2 butanone
39.	Which of the following reagents will react with	both Al	dehydes and Ketones?
	(a) Fehling's reagent	<sup>೨೭/</sup> (b)	Tollen's reagent
	(c) Grignard's reagent	(d)	Benedicts reagent
40.	Tollen's regent is:		
	(a) Alkaline solution containing Potassium tartara		
	(b) Alkaline solution containing Potassium citrate		
	(c) Ammonical AgNO <sub>3</sub>		
44	(d) Ammonical Cu <sub>2</sub> Cl <sub>2</sub>		mont with T /NoOU2
41.	Which compounds will not give Iodoform test of		6-10 E
	(a) Acetaldehyde (c) Butanone	(b) (d)	Acetone 3 – Pentanone
42.	The carbon atom of the carbonyl group is:	(u)	3 - Peritarione
721	(a) sp hybridized	(b)	sp <sup>2</sup> hybridized
	(c) sp <sup>3</sup> hybridized	(d)	None of these
43.	Ketones are prepared by the oxidation of:	N	en anna ann an
	(a) Primary alcohols	(b)	Secondary alcohols
	(c) Tetiary alcohols	(d)	None
44.	Acetone reacts with HCN to form cyanohydrin.		example of:
	(a) Electrophilic addition	(b)	Electrophilic substation
	(c) Nucleophilic addition	(d)	Nucleophlic substitution
45.	Catalyst used for the laboratory preparation of		Con the state of t
	(a) ZnO (b) $Al_2O_3$ (c)	Platin	ized Asbestos (d) None
46.	Hybridization of carbon in carbonyl group is:		(FSD 2010, GRW 2011)
47	(a) $sp^3$ (b) $sp^2$ (c)	dsp <sup>2</sup>	(d) sp
47.	Cannizzaro's reaction is not given by:	200	N 13, 14, LHR 12, 13, BWP 15)
	(a) Formaldehyde (b)		oldehyde
	(c) Benzaldehyde (d)	ırıme	ethyl acetaldehyde

18.	(a)	es are al Primary	alcoho	I	d to:		(b)		econdar		· -	RWP 2008)
19.	(c) Aldol <sub>1</sub> (a)	Tertiary <b>product</b> of Decomp	on hea	iting u	ndergo	oes:	(d) (b)	De	ehydrat	ion		(RWP 2009)
<b>60</b> .	(c) Acidifi (RWP	Rearrandied oxidia 2010)			or the l	labora	(d) <b>tory pr</b>		one of t tion of		dehyd	e is:
	(a) (c)	K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> - K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> -					(b) (d)		a <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> a <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub>		4	
					<u>A</u>	<u>NSW</u>	ER KE	<u>Y</u>				
		1.	2.	3.	4.	5.	6.	7.	8.	9.	10.	]
		a	d	С	b	С	d	С	d	a	С	
		11	12	13	14	15	16	17	18	19	20	
		21	22	23	24	25	2 <b>6</b>	27	28	<b>29</b>	<b>30</b>	
		b		d	d	b	a	b	d	b	С	-
		31	32	33	34	35	36	37	38	39	40	
		a	С	С	d	d	С	b	P	) c	С	-
		<b>41</b>	<b>42</b> b	<b>43</b>	44	<b>45</b>	46 b <sub>2</sub> (	47 b	<b>48</b>	<b>49</b> b	<b>50</b>	<u> </u>
					M	3990						
						DUC C						

### CHAPTER 13

## CARBOXYLIC ACIDS & pakcity.org



	CHAI	PTER 13
	CARBOX	YLIC ACIDS pakcity.org
1.	Which acid of the following is not a	
	<ul><li>(a) Propanoic acid</li><li>(c) Butanoic acid</li></ul>	(b) Phthalic acid (d) Acetic acid
2.	Which reagent is used to reduce a carb	
	(a) H <sub>2</sub> /Ni (c) H <sub>2</sub> /Pd	(b) $H_2/Pt$ (d) LiAl $H_4$
3.	Sodium acetate and Acetyl chloride	
	(a) Acetic acid	(b) Acetone
4	(c) Acetic anhydride	(d) None
4.	(a) Hydrogen bonding	ist as cyclic dimer in vapour state because of:  (b) Polymerization
	(c) Condensation	(d) Distillation
5.	Nature of Glycine amino acid is:	
	(a) Neutral	(b) Acidic
6.	(c) Basic  Acetic acid is obtained when	(d) None
O.	(a) CH <sub>3</sub> OH is oxidized	(b) Calcium acetate is distilled
	(c) Ethanol is oxidized	(d) Ammonium carbonate is heated
7.		is not formed when acetic acid reacts with PCI₅?
	(a) CH₃COCl (c) POCl₃	(b) HCI (d) CH₃Cl
8.		t be prepared directly from acetic acid?
	(a) Acetamide	(b) Acetyl chloride
•	(c) Acetic anhydride	(d) Ethyl acetate
9.	Acetamide is prepared by:  (a) Heating ammonium acetate	(b) Heating methyl cyanide
	(c) Hydrazine	(d) Mixture of NH <sub>4</sub> Cl and acetic acid
10.	Which Ester is used for the flavor of	Banana:
	(a) Benzyl acetate	(b) Amyl acetate
11.	(c) Isobutyl formate  Freezing point of acetic acid is:	(d) Ethyl butyrate
11.	(a) -17°C	(b) 17°C
	(c) -118°C	(d) Room temperature
12.	Setting Contracts of Contract	
	(a) Water (c) Alcohol	(b) Ether (d) All of these
13.	The carbon atom of a carbonyl group	CILV.OI C
— <del>—</del> —	(a) Unhybridized	(b) sp hybridized
<u>y<b>s</b> 194</u>	(c) sp <sup>2</sup> hybridized	(d) sp <sup>3</sup> hybridized
14.	The boiling points of the Carboxylic (a) High	Z1 X 1
	(c) Extremely low	(b) Low (d) Extremely high
15.		
	(a) Aldehyde	(b) Alkene
16	(c) Alcohol  Which of following is commonly kno	(d) Ester
16.	Which of following is commonly kno (a) Formic acid	(b) Butanoic acid
	(c) Formaldehyde	(d) Acetic acid

<b>17.</b>	When Ethyl alcohol is oxidized with K2C1	r2 <b>O</b> 7 and	d dilute H2SO4 , v	which of	f the following
	is produced: (a) Formaldehyde	(b)	Acetic acid		
	(c) Butanoic acid	(b)	Formic acid		
18.	Zwitterion is also called:	(u)	Torrine dela		
10.	(a) Dipolar ion	(b)	Internal Salt	(c)	both a and b
	(d) Non polar ion	(-)	-	(=)	500.100.100
19.	Amino acids can be synthesized by reac	tion of	α-Bromo acid w	rith:	
	(a) Ammonia	(b)	Urea	(c)	Ammonium
	Chloride	(d)	Formaldehyde		
20.	Reaction of acids with alcohols is know	n as:			
	(a) Esterification	(b)	Saponification	(c)	
M. Sec. 1997	Neutralization	(d)	None		
21.	Toluene can be oxidized to Benzoic acid				
	(a) KMnO <sub>4</sub> (acidic medium)		K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> (acidic m	nedium)	
22	(c) Both a and b	(d)	None		- 4 *
22.	When two moles of acetic acid are heat		50.54 549		ea is:
	(a) 2 moles of ethyl alcohol (c) Acetic anhydride		Two moles of Ac		
23.			Ethyl acetate est		hich rospoct?
23.	(a) Replacement of hydrogen by sodium		Formation of este		
	(c) Reduction of Fehling solution		Blue litmus react		ICOTIOI
24.	Glacial Acetic acid at 17°C is:	(u)	Dide litinds react	1011	
\$ <del></del> 74.4	(a) Colourless liquid	(b)	Ice like solid		
	(c) Waxy solid		A gas		
25.	Weakest acid among the followings is:	^ /	110		
	(a) Acetic acid	<(b)	Phenol		
	(c) Water	(C)(d)	Acetylene		
26.	Ethyl alcohol reacts with acetyl chloride	to for	m:		
	(a) Ethyl chloride	(G 20 S	Acetic acid		
	(c) Methyleacetate	(d)	Ethyl acetate		
27.	Which of the following is the strongest	acid:			
	(a) CF <sub>3</sub> COOH	(b)	CBr <sub>3</sub> COOH		
	(c) CH₃COOH	(d)	CCl₃COOH		
28.	HCOOH reacts with conc. H <sub>2</sub> SO <sub>4</sub> to prod	71 N	COLL		
	(a) CO+H <sub>2</sub> O	(b)	CO <sub>2</sub> +H <sub>2</sub>		
29.	(c) HCHO  Hydrolysis of an ester gives a carboxy		None	o's alac	trolycic violde
ZJ.	ethane, the ester is	iic aciu	Willell Oll Kolb	e s elec	ciolysis yleius
	(a) Ethyl methonoate	(b)	Methyl ethanoate	2	
	(c) Ethyl acetate ester	(d)	None	<b>2</b>	
30.	Carboxylic acids are more acidic than pl			use of:	
	(a) Intermolecular hydrogen bonding				
	(b) Formation of Dimers				
	(c) Highly acidic hydrogen				
E P	(d) Resonance stabilization of their conjugate		Carboxylate ion)		
31.	Organic compounds having fruity smell	17232 320	. 200		
	(a) Carboxylic acids	(b)	Alcohols		
~~	(c) Ethers	(d)	Esters		
32.	The solution of the acid used for season				
	(a) Formic acid	(q)	Acetic acid	4	
33	(c) Benzoic acid  Which of the following is not a fatty acid	(d) id2	Butanoic acid	ı	
33.	Which of the following is not a fatty act  (a) Propanoic acid	· 보다	Acetic acid		
	(a) Propanoic acid (c) Phthalic acid	(b) (d)	Butanoic acid	1	
	(c) I fictionic acid	(u)	שמנטווטוכ מכונ	4	

	mistry MCQs www.pakcit
34.	Acetic acid was first isolated from:
ant	(a) Butter (b) vinegar (c) Milk (d) Red
35.	Acetic acid was first isolated from:
26	(a) Butter (b) Vinegar (c) Milk (d) Red ant
36.	The molecular mass of protein is: (a) Less than 10,000 (b) Greater than 10,000
	(c) Equal to 10,000 (d) Equal to 9,000
<b>37.</b>	Amino acids are prepared by:
	(a) Williamson's synthesis (b) Strecker's synthesis (c) Wurtz's synthesis (d) Perkin's reaction
38.	(c) Wurtz's synthesis (d) Perkin's reaction  The molecular mass of protein is:
Salter (Personal	(a) Less than 10,000 (b) Greater than 10,000
20	(c) Equal to 10,000 (d) Equal to 9000
39.	Molar mass of CH₃COOH obtained by elevation of boiling point method is: (LHR 2011)
	(a) 30 (b) 60
40	(c) 120 (d) 180
40.	Which compound is polyprotic acid?  (a) CH₃COOH  (b) C <sub>6</sub> H₄(OH)COOH  (c) pakcity.org
	(c) (COOH) <sub>2</sub> (d) C <sub>6</sub> H <sub>5</sub> OH
41.	Dipolar structure of zwitterions is also called:
	(a) Double salt (c) Internal salt (d) External salt
42.	The flavor of amyl acetate is:
	(a) Orange (b) Apricot
43.	(c) Pine apple  Acetic acid is manufactured by:  Banana  (LHR 14, GRW 14,15, RWP 10)
	(a) Distillation (b) Fermentation
<b>2</b> ,21	(c) Ozonolysis (d) Esterification
44.	The organic acid that does not have carboxyl group is:  (a) Formic acid  (b) Picric acid
	(c) Formaldehyde (d) Acetone
45.	Tyrosine was first isolated from. (RWP 2008)
	(a) Sugar (b) Cheese (c) Butter (d) Milk
46.	Phathalic acid is a: (RWP 2009)
	(a) Monocarboxylic acid (b) Dicarboxylic acid
47.	(c) Tricarboxylic acid (d) None of these  The reagent used to convert carboxylic acid directly to alkane is (RWP 2008)
77.1	(a) HI/P (b) NaBH <sub>4</sub>
	(c) LiAlH <sub>4</sub> (d) H <sub>2</sub> /Ni
48.	Which acid is used in the manufacture of synthetic fiber? (MTN 2015)  (a) Formic acid (b) Oxalic acid
	(c) Carbonic acid (d) Acetic acid
49.	The amino acids which cannot be synthesized by our body are: (MTN 2008)
	(a) 5 (c) 15 (d) 20
<b>50</b> .	Which one of the following is an unsaturated carboxylic acid? (MTN 2009)
	(a) Malonic acid (b) Oxalic acid (c) Oxalic acid
	(c) Succinic acid  ANGWED VEV
	ANSWER KEY
	1.     2.     3.     4.     5.     6.     7.     8.     9.     10.       b     d     c     a     c     d     a     a     b
	11 12 13 14 15 16 17 18 19 20
	b d c a c d b c a a 21 22 23 24 25 26 27 28 29 30
	c c b c d a a c d
	31         32         33         34         35         36         37         38         39         40           d         b         c         b         b         b         b         c
	41         42         43         44         45         46         47         48         49         50           c         d         b         b         b         a         d         b         d

1.	2.	3.	4.	5.	6.	7.	8.	9.	10.
b	d	С	a	a	С	d	а	a	b
11	12	13	14	15	16	17	18	19	20
b	d	С	a	С	d	b	С	а	a
21	22	23	24	25	26	27	28	29	30
С	С	С	b	С	d	a	а	С	d
31	32	33	34	35	36	37	38	39	40
d	b	С	b	b	b	b	b	b	С
41	42	43	44	45	46	47	48	49	50
С	d	b	b	b	b	a	d	b	d

### CHAPTER 14

## MACROMOLECULES pakcity.org



	CH	IAPTER 14
	MACR	OMOLECULES pakcity.org
1.	Which of these polymers is a syn	nthetic polymer:
	(a) Aminal fat	(b) Starch
	(c) Cellulose	(d) Polyester
2.	Which one of the following enzyme	es brings about the hydrolysis of fats:
	(a) Urease	(b) Maltase
	(c) Zymase	(d) Lipase
3.	To which class of compounds cho	olesterol belongs:
	(a) Steroids	(b) Phospholipids
20	(c) Vitamins	(d) Triglycerides
4.	Which of the following is an este	
	(a) Soap	(b) Starch
_	(c) PVC	(d) Triglyceride
5.	Which one of the following is an	
	(a) Graphite	(b) Rubber
6.	(c) DNA  The linear sequence of amino aci	(d) Protein
0.	(a) Primary Structure	(b) Secondary Structure
	(c) Tertiary Structure	(d) Quaternary structure
7.		small organic molecules made into macromolecule:
. <b>5</b> / <b>2</b>	(a) The cracking of petroleum fraction	A     ( )
	(c) Formation of Polyethene	(d) The hydrolysis of proteins
8.	Which of these polymers is synth	
	(a) Animal fat	(b) Starch
	(c) Cellulose	(d) Polyester
9.	Which Carbohydrate can be used	for silvering of mirror:
	(a) Glucose	(b) Fructose
	(c) Maltose	(d) all
10.		carbon atoms:
	(a) 3 to 6	(b) 3 to 7 (d) 7-10
11.	(c) 6-7 Which of the following is not obt	tained by condensation polymerization:
4.4.	(a) Polyester	(b) Nylon
	(c) Polystyrene	(d) None
12.		
<b></b>	(a) Polyester	(b) Orlon
	(c) Polystyrene	(d) Nylon
13.	Orlon is polymer of:	akcity.org
	(a) T.F.E	(b) Acrylonitrile
1557 - 1554 1457 - 1554	(c) Ethanoic acid	(d) Benzene
14.	The raw material used to from Ny	200
	(a) Adipic acid	(b) Butadiene
4 =	(c) Isoprene	(d) Ethylene
15.		(b) Dobyotymono
	(a) Polyethene	(b) Polystyrene
16.	(c) Polyethene  Poly vinyl acetate is used as:	(d) None
10.	Poly vinyl acetate is used as:  (a) Explosive material	(b) Adhesive material
	(c) Rubber	(d) Filler
17.		condensation of poly amines with aliphatic:
17.	(a) Carboxylic acid	(b) Alcohol
	(c) Aldehydes	(d) Dicarboxylic acid
	(5) / 110011/000	(a) Dicarbonyine dela

(a) (c) The (c	Epoxy resin is made by condensation of Diphenylol propane Phenyl acetate etable sugar is: Glucose Maltose Ffinose is an example of: Mono saccharide Trisaccharide Example of compound protein is: Phosphoprotein Albumin Exoptimum PH of salivary amylase is from 6.8 7.0 Exercise reagent which forms crystalline gluctors is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C Example of compound proteins is: An ester linkage The peptide linkage Exircise The fiber which is made from acrylonitrile actions of the corporation of th	(b) (d) (b) (d) (b) (d) (d) (e) (e) (osazon (b) (d) (d) (d) (d) (d) (d) (d)	Phenyl propane Dicarboxylic acid  Sucrose Lactose  Disaccharide None  Peptones Globulins  1: 6.9 None 1e derivative when treated with Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage All
(a) (c) The (c	Diphenylol propane Phenyl acetate Phenyl acetate Phenyl acetate Phenyl acetate Phenyl acetate Phenyl acetate Phospe Phosphose is an example of: Phosphoprotein Albumin Phosphoprotein Benedict solution Benedict solution Phosphoprotein Benedict solution Ben	(b) (d) (b) (d) (b) (d) (d) (d) (e) (osazon (b) (d) (d) (d) (d) (d) (d) (e) (d) (e) (f) (f) (f) (f) (f) (f) (f) (f) (f) (f	Phenyl propane Dicarboxylic acid  Sucrose Lactose  Disaccharide None  Peptones Globulins  Comparise  An ether linkage  An ether linkage  An ether linkage  An ether linkage  Rayon fiber
9. (c) The (c)	Phenyl acetate  a table sugar is: Glucose Maltose  ffinose is an example of: Mono saccharide Trisaccharide a example of compound protein is: Phosphoprotein Albumin a optimum PH of salivary amylase is from 6.8 7.0 a reagent which forms crystalline gluctorse is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C a main structural feature of proteins is: An ester linkage The peptide linkage a fiber which is made from acrylonitrile at PVC Acrylic resins a disaccharide present in milk is:	(d) (b) (d) (b) (d) (b) (d) (expression of the content of the cont	Dicarboxylic acid  Sucrose Lactose  Disaccharide None  Peptones Globulins  C: 6.9 None  Re derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A  Vitamin B6  An ether linkage  All  Dmer is: (LHR-09)  Rayon fiber
19. The contraction of the contr	Glucose Maltose Ffinose is an example of: Mono saccharide Trisaccharide example of compound protein is: Phosphoprotein Albumin optimum PH of salivary amylase is from 6.8 7.0 e reagent which forms crystalline gluctose is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage of fiber which is made from acrylonitrile acid is accharide present in milk is:  The disaccharide present in milk is:  The disaccharide present in milk is:  The peptide linkage	(b) (d) (b) (d) (b) (d) (d) (b) (d) (b) (d) (b) (d) (b) (d)	Sucrose Lactose  Disaccharide None  Peptones Globulins  C: 6.9 None De derivative when treated with Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage All  Dmer is: (LHR-09) Rayon fiber
(a) (b) (c) (c) (d) (c) (d) (d) (d) (d) (d) (d) (d) (d) (d) (e) (e) (e) (e) (e) (e) (e) (e) (e) (e	Glucose Maltose Ifinose is an example of: Mono saccharide Trisaccharide e example of compound protein is: Phosphoprotein Albumin optimum PH of salivary amylase is from 6.8 7.0 e reagent which forms crystalline gluctose is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile acid is accharide present in milk is:	(d) (b) (d) (b) (d) (e) (f) (f) (f) (f) (f) (f) (f) (f) (f) (f	Disaccharide None  Peptones Globulins  6.9 None  Me derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage  All  Mer is: (LHR-09)  Rayon fiber
20. Ra (c) The	Maltose  ffinose is an example of:  Mono saccharide  Trisaccharide  e example of compound protein is:  Phosphoprotein  Albumin  optimum PH of salivary amylase is from  6.8  7.0  e reagent which forms crystalline gluctose is:  Fehling solution  Benedict solution  corbic acid is a chemical name of:  Vitamin D  Vitamin C  e main structural feature of proteins is:  An ester linkage  The peptide linkage  e fiber which is made from acrylonitrile acrylic resins  e disaccharide present in milk is:	(d) (b) (d) (b) (d) (e) (f) (f) (f) (f) (f) (f) (f) (f) (f) (f	Disaccharide None  Peptones Globulins  6.9 None  Me derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage  All  Mer is: (LHR-09)  Rayon fiber
20. Ra (c) The	Mono saccharide Trisaccharide Pexample of compound protein is: Phosphoprotein Albumin Optimum PH of salivary amylase is from 6.8 7.0 Optimum PH of forms crystalline gluctorse is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C Optimum PH of salivary amylase is from 6.8 7.0 Optimum PH of salivary amylase is from	(b) (d) (b) (d) (e) (d) (b) (d) (b) (d) (b) (d) (b) (d)	Disaccharide None  Peptones Globulins  6.9 None  Me derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
(a) (c) The (c	Mono saccharide Trisaccharide e example of compound protein is: Phosphoprotein Albumin e optimum PH of salivary amylase is from 6.8 7.0 e reagent which forms crystalline gluctose is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile at PVC Acrylic resins e disaccharide present in milk is:	(d) (b) (d) (e) (d) (f) (d) (b) (d) (f) (d) (h) (h) (h) (h) (h) (h) (h) (h) (h) (h	Peptones Globulins  6.9 None  Me derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage  An ether linkage  Rayon fiber
21. (c) The (c	Trisaccharide e example of compound protein is: Phosphoprotein Albumin e optimum PH of salivary amylase is from 6.8 7.0 e reagent which forms crystalline gluc cose is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile a PVC Acrylic resins e disaccharide present in milk is:	(d) (b) (d) (e) (d) (f) (d) (b) (d) (f) (d) (h) (h) (h) (h) (h) (h) (h) (h) (h) (h	Peptones Globulins  6.9 None  Me derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage  An ether linkage  Rayon fiber
1. The (a) (b) (a) (c) The (c)	Phosphoprotein Albumin coptimum PH of salivary amylase is from 6.8 7.0 reagent which forms crystalline gluctose is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C main structural feature of proteins is: An ester linkage The peptide linkage Fiber which is made from acrylonitrile application of the period of t	(b) (d) (b) (d) (cosazon (b) (d) (b) (d) (b) (d) (b) (d)	Peptones Globulins  6.9 None  Me derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
2. The (a) (b) (a) (c) The (a)	Albumin c optimum PH of salivary amylase is from 6.8 7.0 c reagent which forms crystalline gluctose is: Fehling solution Benedict solution corbic acid is a chemical name of: Vitamin D Vitamin C c main structural feature of proteins is: An ester linkage The peptide linkage c fiber which is made from acrylonitrile at PVC Acrylic resins c disaccharide present in milk is:	(d) n 6.4 to (b) (d) cosazon (b) (d) (b) (d) (b) (d) (b) (d)	Globulins 6.9 None 1e derivative when treated with Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B6  An ether linkage All  Comer is: (LHR-09) Rayon fiber
2. The (a) (c) The	e optimum PH of salivary amylase is from 6.8 7.0 e reagent which forms crystalline gluctose is: Fehling solution Benedict solution corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile at PVC Acrylic resins e disaccharide present in milk is:	(d) n 6.4 to (b) (d) cosazon (b) (d) (b) (d) (b) (d) (b) (d)	6.9 None  No
(a) (c) The (c	6.8 7.0 e reagent which forms crystalline gluctose is: Fehling solution Benedict solution corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile at PVC Acrylic resins e disaccharide present in milk is:	(b) (d) cosazon (b) (d) (b) (d) (b) (d)	None None Ne derivative when treated with Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B <sub>6</sub> An ether linkage  All  Omer is: (LHR-09) Rayon fiber
(c) The (d) (a) (a) (b) (b) (c) The (c	reagent which forms crystalline gluctose is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C E main structural feature of proteins is: An ester linkage The peptide linkage E fiber which is made from acrylonitrile at PVC Acrylic resins E disaccharide present in milk is:	(d) cosazon (b) (d) (b) (d) (b) (d) (s mono	None  le derivative when treated with  Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B <sub>6</sub> An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
23. The (a) (a) (b) (b) (a) (c) (c) (c) (c) (c) (c) (c) (c) (c) (c	reagent which forms crystalline gluctose is: Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C Emain structural feature of proteins is: An ester linkage The peptide linkage Emain fiber which is made from acrylonitrile at PVC Acrylic resins Endisaccharide present in milk is:	(b) (d) (b) (d) (b) (d) (s mond	Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B <sub>6</sub> An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
glu (a) (c) As (a) (c) As (a) (c) Th (a) (c) Th (a) (c) Th (a) (c) Th (a) (c) Ba (c) B	Fehling solution Benedict solution Corbic acid is a chemical name of: Vitamin D Vitamin C E main structural feature of proteins is: An ester linkage The peptide linkage E fiber which is made from acrylonitrile at PVC Acrylic resins E disaccharide present in milk is:	(b) (d) (b) (d) (b) (d)	Phenyl hydrazine Hydroxyl amine  Vitamin A Vitamin B <sub>6</sub> An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
(a) (c) (d) (d) (e) (e) (e) (e) (f) (e) (f) (f) (f) (f) (f) (f) (f) (f) (f) (f	Fehling solution Benedict solution corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile a PVC Acrylic resins e disaccharide present in milk is:	(d) (b) (d) (b) (s mond	Hydroxyl amine  Vitamin A  Vitamin B <sub>6</sub> An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
(c) As (a) (a) (b) Th (a) (c) Th	Benedict solution corbic acid is a chemical name of: Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile a PVC Acrylic resins e disaccharide present in milk is:	(d) (b) (d) (b) (s mond	Hydroxyl amine  Vitamin A  Vitamin B <sub>6</sub> An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
24. (a) (c) Th (a) (c)	Vitamin D Vitamin C e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile at PVC Acrylic resins e disaccharide present in milk is:	(b) (d) (b) (d) (s mond	Vitamin A Vitamin B <sub>6</sub> An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
(a) (c) The (c	Vitamin D Vitamin C  e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile a PVC Acrylic resins e disaccharide present in milk is:	(d) (b) (d) (s mond	An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
(c) The (c) Th	Vitamin C  e main structural feature of proteins is: An ester linkage The peptide linkage e fiber which is made from acrylonitrile a PVC Acrylic resins e disaccharide present in milk is:	(d) (b) (d) (s mond	An ether linkage  All  Omer is: (LHR-09)  Rayon fiber
25. The (a) (c) 26. (a) (c) 27. (a) (c) 28. (a) (c) 29. (a) (c) 29. (a) (c) 31. (a) (c) 31	An ester linkage The peptide linkage  fiber which is made from acrylonitrile a  PVC  Acrylic resins  disaccharide present in milk is:	(b) (d) (s mond	An ether linkage  All  omer is: (LHR-09)  Rayon fiber
(a) (c) (c) (d) (d) (e) (e) (e) (e) (final dots) (final d	An ester linkage The peptide linkage e fiber which is made from acrylonitrile a PVC Acrylic resins e disaccharide present in milk is:	(d) is mond	omer is: (LHR-09) Rayon fiber
(c) The (a) (c) The (a) (c) St. (b) St. (a) (c) St. (b) St. (b) St. (c) St. (c) St. (d) St. (d	The peptide linkage  e fiber which is made from acrylonitrile a  PVC  Acrylic resins  e disaccharide present in milk is:	(d) is mond	omer is: (LHR-09) Rayon fiber
26. The (a) (b) 27. (a) (c) 28. (a) (c) 29. (a) (c) 31. (a) (c) 32. (a) (c) 33. (a) (c) 34. (a) (a) (b) 34. (a) (a) (b) 34. (a) (a) (b) 34. (a) (b) 34. (b) 34. (a) (c) 34. (b) 34. (a) (c) 34. (b) 34. (c) 34. (c) 34. (c) 34. (c) 34. (c) 34. (d) 34	PVC Acrylic resins e disaccharide present in milk is:	s mond	omer is: (LHR-09) Rayon fiber
(a) (c) The (a) (c) St. (b) St. (a) (c) St. (b) St. (b) St. (c) St. (c) St. (c) St. (d) (c) St. (d) (c) St. (d) (d) St. (d	PVC Acrylic resins e disaccharide present in milk is:	(P)	Rayon fiber
27. The (a) (c) 28. On (a) (c) 29. En: (a) (c) 31. (a) (c) 32. (a) (c) 33. (a) (c) 34. (a) (c) 34. (a)	e disaccharide present in milk is:	100	Polvester fiber
(a) (c) (a) (a) (c) (c) (c) (c) (c) (c) (c) (c) (c) (c	- 111		1 01/ 00001 11001
(c) 28. On (a) (c) 29. En: (a) (c) 31. (a) (c) 32. (a) (c) 33. (a) (c) 34. (a) (c) 34. (a)	Cucroco		
8. On (a) (c) (d) (d) (d) (d) (d) (d) (d) (d) (d) (d	Sucrose	(b)	Maltose
(a) (c) (c) (a) (c) (c) (c) (c) (c) (c) (c) (c) (c) (c	Lactose	(d)	Cellobiose
(c) (a) (c) (c) (c) (c) (c) (c) (c) (c) (c) (c	hydrolysis of starch, we finally get:	B 20 CO	
9. En: (a) (c) (b) (c) (c) (c) (d) (c) (d) (d) (d) (d) (d) (d) (d) (d) (d) (d	Glucose	(b)	Fructose
(a) (c) (a) (a) (c) (c) (c) (c) (c) (c) (c) (c) (c) (c	Both a and b	(d)	Sucrose
(c) Ba (a) (c) B1. (a) (c) (a) (a) (c) (a) (a) (a) (a) (b) (a) (a) (b) (a) (b) (b) (c) (c) (c) (c) (c) (d) (d) (d) (d) (d) (d) (d) (d) (d) (d	zymes in the living systems:	(b)	Dravida immunity
(a) (c) (d) (d) (d) (d) (d) (d) (d) (d) (d) (d	Provide energy Transport evergen	(p)	Provide immunity Catalyze biochemical processes
(a) (c) 31. En: (a) (c) 32. Wh (a) (c) 34. The (a)	Transport oxygen kelite is obtained from phenol by reactin	(d)	
(c) 81. En: (a) (c) 82. Wh (a) (c) 83. (a) (c) 84. The (a)	Acetaldehyde	(b)	Acetal
(a) (c) (a) (c) (a) (c) (a) (c) (a) (c) (c) (a) (c) (d) (d) (d) (d) (d) (d) (d) (d) (d) (d	Formaldehyde	(d)	Chlorobenzene
(a) (c) (a) (a) (c) (a) (c) (c) (a) (a)	zyme used for the treatment of blood car		
(c) (a) (a) (c) (a) (a) (c) (c) (a)	Cellulose	(b)	Urease
(a) (c) (3. Wh (a) (c) (c) (4. The	L-asparaginase	(d)	lactic dehydrogenase
(a) (c) 33. Wh (a) (c) 44. The	ich of the following is not present in nu		
3. Wh (a) (c) 4. The (a)	Guanine	(b)	Cytosine
(a) (c) <b>84. Th</b> (a)	Adenine	(d)	Tyrosine
(c) <b>84. Th</b> (a)	ich of the following has magnesium in it	s struct	
<b>34.</b> The (a)	Carbonic anhydrase	(b)	Haemocyanin
(a)	Chlorophyll	(d)	Vitamin B <sub>12</sub>
	e number of amino acids found in proteir		
/_\		(p	
(c)		d) tobovos	3.6 ·
969 050	ich of the following is an example of Ke	:212: 0	20 NEC 1127
935 500	Clucoco	(d	24.00 (Heading Control of the Contro
(C)	Glucose	(d	i) Huctose
	Glucose Maltose		

		10 <mark>111111111111111111111111111</mark> 111111111				HIIIIIIIIIIIIIIIIIIIIIIIIIIIIIIIIIIIII				HIIIIIIIIIIII <mark>I</mark> IIIII		
36.	The gr	roup linka	age pr	esent	in Car	bohyd	rates:					
	(a) Per	otide linka	ge					(b	)	Ester	linkag	e
		cosidic lin						(d	)	Diest	er linka	age
37.	Teflon	, styrene	and N	leopre	ene ar	e all:						
		polymers						(b				on polymers
		mopolyme				· · · · · · · · · · · · · · · · · · ·	•	. (d		Mond	omers	
38.	M44	of the fo	ilwolld	ng is fi	ully flu	iorina	ted po			т. а.		
	(Se) (Se)	oprene						(q)	-	Teflo	n	
39.	(c) Thi	okol e <b>gree of I</b>	ıncəti	ıratior	1 of a 1	fat or :	an oil	d) is chec	<u> </u>	PVC		
33.		drogenatio			i Oi a i	at or a	all Oll	(b)	e e		e numl	her
		onification						(d			ne num	_
40.	500 St.	h is a pol							,			
	(a)	Fructose						(b	)	$\alpha$ -D-	Glucose	9
	(c)	Sucrose						(d	)	Lacto		
41.	Which	n one of t	he fol	lowing	g enzy	mes b	rings	about	the hy			ats?
	(a)	Urease						(b	-81	Malta		
40	(c)	Zymase	<b>=</b> 1			J.11		(d	)	Lipas	e	
42.	9 6	of these		mers i	s an a	aditioi	n poly	D:0	`	Del	<b></b>	d b
	(a)	Nylon-6						(q		1	tyrene	- pakcity.org
43.	(c)	Terylene resins a		damo	ntallv			(d	)	∟pox	y resin	3
101	(a)	Polyami	8		aiiyi	Į.		(b	) /	√ Polve	ethers	
	(c)	Polyeste						(d	10	Poly		
44.		n one of t		lowing	g acid	is use	d to p	reparé	synth			
	(a)	Carbonio	c acid				^	(b)		Form	ic acid	
	(c)	Acetic a					45	(3)(d	)		ric acid	
<b>45</b> .	Which	n stateme				and s	ucrose		-			-
	(a)	Both are					5	(b	5/2			naturally
16	(c)	Both are			A	000	Luila a	(d	•		are dis	accharides
46.		ber whic PVC	n is m	(b)		on fibe				: Fiber	(d)	polyester
fiber	(a)	1 4 C		(0)	Mich	OII IIDC		c)	ACITIC	, i ibci	(u)	polyester
47.	Which	n of these	ıylog s	ners i	s an a	dditio	ylog r	mer?				
	(a)	Polystyr		4.1		on-6,6			Polyes	ter	(d)	Epoxy resin
48.	Enzyn	ne used f	or the	treat	ment o	of bloc	d can	cer in	childr	en is:		
	(a)	Cellulose	9		S EI			(b	)	Urea	se	
	(c)	L-Aspar	- h		The last			(d				drogenase
49.	7 7	1-6,6 is ol		d by h	eating	, hexa	meth					(FSD 2010)
	(a)	Adipic a						(p)			c acid	
50.	(c)	Vinyl chi nerization		nyloni	trilo a	ivo		(d)		ACTY	ic acid	(FSD 2011)
JU.	POIVII	ici izaliUl	. oi at	y 10111	ane g	akai						(130 2011)
	(a)	PVC						(b)		Rayo	n Fiber	
	(c)	Acrylic f	iber					(d)			ester fib	
					11=							
					<u>A</u>	<u>NSW</u>	EK KE	<u>Y</u>				
		1	2	3	4	5	6	7	8	9	10	
		d	d	а	d	а	а	С	d	a	b	
		11	12	13	14	15	16	17	18	19	20	
		С	d	b	a	a	b	d	a	b	С	
		21	22	23	24	25	26	27	28	29	30	
		a	b	b	С	С	С	С	a	d	С	
		31	32	33	34	35	36	37	38	39	40	
		С	d	С	b	d	С	С	b	b	b	
		41	42	43	44	45	46	47	48	49	50	
			L	h		d	_	<b>a</b>	_	3	C	
		d	b	b	C	u	С	a	C	а		J

1	2	3	4	5	6	7	8	9	10
d	d	a	d	a	a	С	d	a	b
11	12	13	14	15	16	17	18	19	20
С	d	b	a	a	b	d	a	b	С
21	22	23	24	25	26	27	28	29	30
a	b	b	С	С	С	С	a	d	С
31	32	33	34	35	36	37	38	39	40
С	d	С	b	d	С	С	b	b	b
41	42	43	44	45	46	47	48	49	50
d	b	b	С	d	С	a	С	а	С



	CHAPTER	pakcity.org
C	COMMON CHEMICAL INDUS	TRIES IN PAKISTAN
1.	The percentage of nitrogen in NH₃ is:	
	(a) 46% (c) 82%	(b) 60% (d) 100%
2.	Potassium fertilizers are useful for:	(a) 10070
	(a) Tobacco	(b) Coffee
	(c) Potato	(d) all of these
3.	Which one of the following raw materials is no	t used in the cement manufacture:
	(a) Lime stone	(b) Gypsum
_	(c) Blast furnace slag	(d) Lead
4.	White water is:	
	(a) Hard water (b) Water obtained from a crystal by beating	
	<ul><li>(b) Water obtained from a crystal by heating</li><li>(c) Water which is obtained from pulp, through scre</li></ul>	en at Fourdrinier table
	(d) Water which is removed from amino acids when	
5.	The % of Nitrogen in Ammonium Nitrate is:	(GRW 2009)
-	(a) 46%	(b) 82%
	(c) 33%	(d)\16%
6.	Which is not a calcareous material?	(SGD 14, MTN 11,13, LHR 14,15)
	(a) lime	(b) clay
	(c) marble	(d) marine shell
7.	The nitrogenous fertilizer easily taken up by p	
	(a) Urea	(b) Ammonium nitrate
8.	(c) Liquid ammonia  Ammonium nitrate fertilizer is not used for wh	(d) Ammonia solution
0.	(a) Cotton	ich crops? (LHR 2010) (b) Wheat
	(c) Sugarcane	(d) Paddy rice
9.	Point out raw material which is most suitable	
	(a) CO <sub>2</sub> , N <sub>2</sub> , H <sub>2</sub>	(b) N <sub>2</sub> , H <sub>2</sub> , CO
	(c) CH <sub>4</sub> , N <sub>2</sub> , H <sub>2</sub>	(d) H <sub>2</sub> , N <sub>2</sub>
10.	Which one of the following fertilizers has max	mum manufacturing plants in Pakistan:
	(a) Urea	(b) Ammonia
	(c) Ammonium phosphate	(d) Ammonium nitrate
11.		(LHR 14, MTN, FSD, SWL 15)
	(a) Root (c) Stem	(b) Leave (d) Seed
12.		
	(a) NaNO <sub>3</sub>	(b) KNO₃
	(c) NH <sub>4</sub> NO <sub>3</sub>	(d) Urea
13.	Urea contains about	
	(a) 48% Nitrogen	(b) 46% Nitrogen
	(c) 44% Nitrogen	(d) 42% Nitrogen
14.	Potassium nitrate is prepared by direct reaction	sans and a second of the contract of the contr
	(a) Nitric acid	(b) Nitrosyl chloride
	(c) Sodium nitrate	(d) Nitrate ions

15.	The total production of urea fertilizer in Paki	stan is:	
	(a) 56,20,10 metric tons /annum	(b) 56,25,100 metric t	
(a) * a=0	(c) 56, 23,100 metric tons/annum	(d) 56, 30,100 metric	tons /annum
16.	For Chemical pulping, the principal methods		
	(a) Five	(b) Four	
17.	(c) Three  During setting of cement, Tricalcium silic	(d) Two	aluminate are get
17.	hydrolyzed to produce calcium hydroxide and		alullillate are get
	(a) Ca(OH) <sub>2</sub> .2H <sub>2</sub> O	(b) CaCO <sub>3</sub> .2H <sub>2</sub> O	
	(c) Al(OH) <sub>3</sub>	(d) CaCO <sub>2</sub>	
<b>18.</b>	For manufacture of Portland cement main m	aterials are:	
	(a) Lime and Silica	(b) Alumina and Magr	nesia
	(c) Silica and FeO	(d) Clay and Shale	
19.	In cement manufacture, 75% Lime stone is (	1	
	(a) 62% Lime (c) 75% Lime	(b) 60% Lime (d) None	
20.	The natural fertilizer is called:	(u) None	
	(a) Manure	(b) Urea	
	(c) Super phosphate	(d) Ammonium sulpha	ate
21.	Which of the following elements is not a mid		(RWP 2009)
		c) Mg	(d) Mo
22.	Macronutrients are required for an acre of la	and in quantity ranging	g from:
	(SGD 14, RWP 10)	(P) 3 300Ka	
	(a) 2-200Kg (c) 4-200Kg	(b) 3-200Kg (d) 5-200Kg	
23.	Calendar stock is the stage of paper making		
	(a) Paper is stored	(b) Thickness is reduc	ed
	(c) water is removed	(d) Stock is reduced to	
24.	Which non-woody raw material is used for n	naking pulp and paper	?
	(a) Fur (b) Rice/wheat straw	(c) Eucalyptus	s (d) Poplar
25.	The hottest zone in rotary kiln is:	(I) Design	(MTN 2008)
	(a) Drying	(b) Pre-heating	
26.	(c) Burning Which sequence of steps is correct for the m	(d) Decomposition	•
20.	(a) Mixing, heating, grinding, crushing	ianulacture of cement	=
	(b) Crushing, heating, mixing, grinding	ON 22 II	
	(c) Crushing, heating, mixing, grinding		
	(d) Crushing, grinding, mixing, heating		
27.	It is not used in paper and pulp industry:		(MTN 2009)
	(a) Bamboo (b) Cotton stalk	(c) Poplar	(d) He gas
28.	A manure is:  (a) An organic compound	rg (h) An ingressia saus	(MTN 2009)
	<ul><li>(a) An organic compound</li><li>(c) A mixture of organic and inorganic compounds</li></ul>	(b) An inorganic comp (d) A mixture of inorg	
29.	The fertilizer that contains largest amount of		Jame Compounds
.a. 55	(a) Liquid nitrogen	(b) Urea	
	(c) Liquid ammonia	(d) Ammonium nitrate	9
<b>30.</b>	Argillaceous material in the following is:		(MTN 2011)
	(a) Lime	(b) Clay	
, 4	(c) Marble	(d) Marine shell	/BILLE 6555
31.	Which substance in cement has greater perc	\$ <del>1 - 5</del> 40	(BWP 2008)
	(a) Silica (SiO <sub>2</sub> ) (c) Iron Oxide (Fe <sub>2</sub> O <sub>3</sub> )	(b) Lime (CaO) (d) Alumina (Al <sub>2</sub> O <sub>3</sub> )	
	(C) 2.3.1 C/1.30 (1 C/2.5)	(a) Filannia (Fil2O3)	

32.	Which	one is a	n ora	anic fe	rtilizer	•						(BWP 2009)
<b>5</b>	(a) Manure								mmoniu	ım nitra	ite	(2 2000)
22	(c) Urea							(d) b	oth a &	С		/DWD 2040\
33.	(a) Bord	one is a	micro	onutrie	ent?			(h) N	itrogen			(BWP 2010)
		sphorous	5						otassiur			
34.		rtilizer v	which	contai	ns 46%	% nitro	gen is:	0.22 1.22 1.20				(DGK 2009)
	(a) Urea	a nonium ı	nitrate						mmonia one of t			
35.		ceous n		al used	for th	e manı	ufactu				es	
	(a) Basi	c compo	nents					(b) A	mphote	ric com	pounds	
36.		ic compo <b>ronium</b>		hate c	ontain	C'		(d) b	oth acid	lic and I	oasic co	omponent (BWP 2012)
50.		6 Nitroge		ace c	Jiidiii	<b>_</b> 1		(b) 4	8% P <sub>2</sub> O	5		(2111 2012)
	(c) 88%	plant n	utrient		·		<b>.</b> •		0% Nitr			
37.	<b>The %</b> (a) 1.0	of lime	(CaO	) in Po	rtland	cemen	t is:	(b) 2	5			(SGD 2012)
	(a) 1.0 (c) 62.0	)						(d) 6				
38.		of gyps	sum in	Portla	nd cer	nent is	S <b>:</b>		22.1			(LHR 2012)
	(a) 4-59 (c) 2-39							(b) 9 (d) 1				
	(c) 2 3 /	70						(u) 1	100			
					Δ	WZN	ER KE	Y (	0/			
					=		Q.{	1/2	>			
						_	18					
		1	2	3	4	50	6	7	8	9	10	7
		c	d	d	С	() c	b	C	d	a	a	<u>.</u> -
		11	12	13	14	15	16	17	18	19	20	
		а	d	ba	C	d	С	С	a	a	a	<u> </u>
		21	22	23	24	25	26	27	28	29	30	
		<b>31</b>	<b>32</b>	<b>33</b>	<b>34</b>	35	<b>36</b>	<b>37</b>	<b>38</b>	С	b	
		b	d	a	a	C	b	C	a	7. 10.		
						oakci	ty.or					

## CHAPTER 16

# ENVIRONMENTAL CHEMISTRY & pakcity.org



	CHAPT	ER 16	
			<u>*</u>
	ENVIRONMENTA	AL CHEMISTRY	pakcity.org
1.	Disinfection of water by chlorine is done (a) NH <sub>2</sub> Cl	e by the production of: (b) NCl <sub>3</sub>	(LHR 2005)
	(c) HOCl	(d) NHCl <sub>2</sub>	
2.	Following is better to disinfect water:		(LHR 2011)
	(a) Cl <sub>2</sub>	(b) O <sub>2</sub>	
3.	(c) O <sub>3</sub> In which layer of the atmosphere is Ozo	(d) KMnO <sub>4</sub>	
Yı	(a) Thermosphere	(b) mesosphere	
	(c) Stratosphere	(d) Troposphere	
4.	Which one of the following is not a pollo	utant:	
	(a) CO <sub>2</sub>	(b) NO <sub>2</sub>	
5.	(c) CO  The smea which has high contents of So	(d) SO <sub>2</sub>	
Э.	The smog which has high contents of Social Reducing smog	(b) Oxidizing smog	
	(c) Natural smog	(d) Neutral smog	
6.	Which one of the following diseases is r	. ,	esticides:
	(a) Sleeping sickness	(b) Rickets	
	(c) Malaria	(d) Yellow fever	
7.	C.O.D of water can be determined direction of the O = 2	tly:	
	(a) $Cr_2O_3^{-2}$ (c) $Cr^{+3}$	(d) Cr <sub>2</sub> O <sub>7</sub> -2	
8.	Cracking of Polyethene at high tempera		
	(a) Allotropes	(b) Isomorphs	
	(c) Polymers	(d) Monomers	
9.	The residual ash after incineration of in	ndustrial waste is disp	
	which is lined with:	(b) Clay and plactic	(LHR 2011)
	(a) Portland cement (c) Stone-ware	<ul><li>(b) Clay and plastic</li><li>(d) Methyl silicone</li></ul>	
10.			eration of industrial
	hazardous waste process has a range:		(DGK 09, GRW 08)
	(a) 950 to 1300 ℃	(b) 900 to 1000 ℃	
4.4	(c) 250 to 500 ℃	(d) 500 to 900 ℃	(DWD 004 E)
11.	and the second s	(b) 1000 Km	(BWP 2015)
	(a) 1500 Km (c) 500 Km	(b) 1000 Km (d) 100 Km	
12.	The fresh water being used for domesti		(FSD-10, BWP-09)
	(a) 8% pakcit	(b) 23%	(7 <b>-</b> 8
) <u>F</u> 252	(c) 69%	(d) 100%	
13.	The normal amount of overhead ozone	AGE AT SOUTH TOUR VICES	N-08,10, DGK-08)
	(a) 300 Du (c) 400 Du	(b) 350 Du (d) 450 Du	
14.			
	(a) 250 D.U	(b) 300 D.U	
	(c) 350 D.U	(d) 400 D.U	
<b>15.</b>	The mean residence time of methane in	the atmosphere:	
	(a) 3-6 years	(b) 3-5 years	
4 %	(c) 3-7 years	(d) 3-8 years	
16.	Photochemical smog consists of high co		
	(a) Oxidants (c) Acids	<ul><li>(b) Reductants</li><li>(d) Bases</li></ul>	
	(c) Acids	(d) Dases	

17.		(h)	Lave
	(a) High (c) 10 °C	(b) (d)	Low -10 °C
18.		` /	
10.	(a) 4-7 ppm		4-9 ppm
	(c) 4-8 ppm		4-10 ppm
19.		(u)	т то ррпп
	(a) Re-processing	(b)	Depolymerization
	(c) Transformation	(d)	All
20.	Which of following is three times light	¥	
	(a) Carbon monoxide		Carbon dioxide
	(c) Both a and b	(d)	None
21.	Atmosphere of big/metropolitan cities	is pollu	ted most by:
	(a) Automobile exhausts	22	Pesticide residue
	(c) Household waste		Radio-active fall out
22.	The Laboratory Section 1. Description 1. The Control of the Contro	V	No. 420 1 1 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2
	(a) Sulphur dioxide	(b)	Carbon monoxide
	(c) Carbon dioxide	(d)	NO <sub>X</sub> + CFC's
23.	Pollutant of automobile exhausts that	affects i	nervous system and produces menta
	disease is:		
	(a) Mercury	(b)	Lead
	(c) Nitrogen oxide		Sulphur oxide
24.	A STATE OF THE STA	<i>2</i> 1.3	
	(a) Smog	(b)	Acidic Rain
	(c) Both a and b	(d)	None
25.	* *		1/0
	(a) Inactivates nerves		Inhibits glycolysis
20	(c) Combines with oxygen	$\sqrt{\mathcal{Q}(a)}$	Combines with hemoglobin
26.		, fuele	
	(a) Excess NO <sub>2</sub> and SO <sub>2</sub> from burning fossil		
	(b) Excess production of NH <sub>3</sub> by industry ar	33.	
	(c) Excess release of carbon monoxide by in	-	
27.	(d) Excess formation of CO <sub>2</sub> by combustion Atmospheric pollutant is:	and anim	lai respiration
<b>4</b> /•	(a) CO <sub>2</sub>	(b)	CO
	(a) CO <sub>2</sub> (c) O <sub>2</sub>	(d)	N <sub>2</sub>
28.			
	(a) Inhalation of seasonal pollen		Eating of seasonal vegetables
	(c) Low temperature	(d)	Wet and dry environment
29.			
	(a) Forest fires	(b)	Increased incidence of skin cancer
	(c) Global warming	(d)	None
30.	Pollution is: pakci	tv.org	
	(a) Removal of top soil		
	(b) Release of toxic/undesirable materials in	n environi	ment
	(c) Wastage of energy		
	(d) All of above		
31.	Which causes water pollution:		
	(a) Smoke	(b)	Automobile exhausts
	(c) Pesticides	(d)	All
32.		page 440	we store and
	(a) Organic matter	(b)	Microbes
	(c) Microbes and organic matter	(d)	None
33.		<u> </u>	
	(a) Skin cancer	(b)	Mouth cancer
	(c) Lung cancer	(d)	Liver cancer

	nistry MCQs	www.pako
34.	BOD is:	
	(a) Biological oxygen deficit	(b) Biosphere oxygen demand
25	(c) Biological oxygen demand	(d) None of the above
35.		(b) Carbon dioxide
	(a) Sulphur dioxide (c) Carbon particles	<ul><li>(b) Carbon dioxide</li><li>(d) Industrial discharges</li></ul>
36.		
	(a) Free radicals	(b) -Ve ions
	(c) +Ve ions	(d) All
37.	Increasing skin cancer and high mutation rate a	re due to:
	(a) Acid rain	(b) Ozone depletion
20	(c) CO	(d) Smog
38.		(b) Antarctica
	(a) Europe (c) Asia	(b) Antarctica (d) Africa
39.	Ozone hole refers to:	(a) milea
25 Th 20 h	(a) Physical hole in ozone layers	
	(b) Reduction in thickness of ozone layer in stratosphe	re
	(c) Reduction of thickness of ozone in troposphere	
<u> 22</u> 8,522	(d) Increase concentration of ozone	
40.	Environmental pollution affects:	
	(a) Biotic components	
	(b) Plants only (c) Man only	200
	(d) Biotic and abiotic components of environment	
41.	Water is often treated with chlorine to:	
<del>303 3 7</del>	(a) Increase oxygen content	(b) Kill germs
	(c) Remove hardness	(d) Remove suspended particles
42.	Result of ozone hole is	
	(a) Greenhouse effect	(b) Global warming
40	(c) Acid rain	(d) UV rays reach the earth
43.	Co-ordination number of Fe in [Fe(CN) <sub>6</sub> ] <sup>4</sup> :	(h) 2
	(a) 4 (c) 6	(b) 2 (d) –4
44.	Disinfection of water by chlorine is done by the	
	(a) $NH_2Cl$ (b) $NCl_3$ (c)	HOCI (d) NHCl <sub>2</sub>
45.	The residence time of NO in atmosphere is:	(SGD 2010)
75. Visar	(a) 30 minutes (b) one day (c)	Three days (d) Four days
46.	In purification of potable water the coagulant u	
	(a) Nickle sulphate	(b) Copper sulphate
47.	(c) Barrium sulphate  Atmosphere contains carbon dioxide:	(d) Alum
	(a) 0.01% (b) 0.02% (c)	0.03% (d) None
48.	Biochemical oxygen demand is the capacity of	and the state of t
	consume oxygen with in a period of:	(RWP 2008)
5-	(a) 3 days (b) 4 days (c)	5 days (d) 6 days
49.	Half of the mass of the atmosphere is concentrated	
FO	(a) 4.6 Km (b) 5.6 Km (c)	3.6 Km (d) 15 Km
50.	Lithosphere extends in Earth up to the depth of	**
51.	(a) 50 Km (b) 100 Km (c) Chlorofluorocarbons play an effective role in re	150 Km (d) 30 Km
<b>J1.</b>	(a) Troposphere (b) Stratosphere (c)	Polar region (d) Equator
	(-) (-)	(-)

52.		phere ntage i		inly m	nade u	p of :	11 ele	ments	, the e	elemer		nd in highest 11, BWP 10)
	(a)	Sodiur		(b)	Cal	lcium	(	c)	Carbon		(d)	Silicone
<b>53.</b>	Which	one o	f the fo	ollowir	ng is no	ot the	affect	of acid	l rain?			(MTN 2008)
	(a)			he perc	_							
	(b)					um, me	ercury a	ind lead	d from s	oil		
	(c)			ne build	-	I						
54.	(d) <b>How n</b>			the pH			Ocean					(MTN 2008)
J <b>T.</b>	(a)	97%	ai tii S	(b)	87 <sup>0</sup>		نز.	c)	77%		(d)	67%
<b>55.</b>			eatly a	affects		, ,	•	<b>-</b> /	,,,,		(3)	(MTN 2009)
	(a)	Aquati	7-50				(	b)	moderr	life		
	(c)	Terres	trial life	е					plant's	life		
<b>56.</b>	The O						-					
	(a)		Km hig				172		26-29 k			
E7	(c)		Km hig	era ne	do ef e	711BB	•		20-28 k	m high	)	(DWD 2000)
<b>57.</b>				capab	ne ot s	suppor	,		V	ro		(BWP 2008)
	(a) (c)	Atmos Lithos	( <u>*</u> )						Biosphe Hydrosi			
58.	` '	6		one co	oncent	ration	`				is occ	urring due to
												p to the year:
	(a)	1960					,	b)	1970	1		
	(c)	1980					(	d)	1990	×.		
<b>59.</b>	-		of envi	ronme	nt whi	ich cor	nsists (	of all w	vater b	odies i	s:	
	(DGK 2	12-0.0	entre de la constante de la co				,	(	0)\\	-1		
	(a)	Biosph					(	MANIN	Hydros			
60.	(c)	Lithos		recycle	ene he	in and	aggin	1//	Atmosp w man		·c2	(LHR 12,15)
00.	(a)	2	an be	(b)	eu aga 3			c)	4	y cillic	.J :	(d) $5$
61.		one is	secon	dary p	olluta	nt??	in ,	-,	·			(-)
	(a)	$SO_2$		(b)	H <sub>2</sub> (	ÇO <sub>3</sub>	(	c)	CO			(d) CO <sub>2</sub>
<b>62.</b>	Which	one is	most		M	30			(1284)/ 1224 <sup>(1</sup>			(LHR 2015)
<b>60</b>	(a)	SO <sub>2</sub>		(b)				c)	CO			(d) CO <sub>2</sub>
63.		<b>5.00</b> 5.00	pollut	ed rain (b)	5.6				6.50			( <b>LHR 2014</b> ) (d) 7.00
	(a)	3.00		(D)	3.0	10		c)	0.50			(d) 7.00
					A	NSW	ER KE	<u>Y</u>				pakcity.org
		1	2	3	4	5	6	7	8	9	10	
		С	С	С	a	а	b	d	d	<u>b</u> _	a	
		11	12	13	14	15	16	17	18	19	20	
		b	a	b	С	С	a	b	C	d	a	
		21	22	23	24	25	26	27	28	29	30	
		a	d	b	С	d	a	b	a	b	b	
		31	32	33	34	35	36	37	38	39	40	
		С	С	a	С	d	a	b	b	b	d	
		41	42	43	44	45	46	47	48	49	50	
		b	d	С	С	d	d	С	С	b	b	
		51	52	53	54	55	56	57	58	59	60	
		С	d	а	a	а	a	b	С	b	d	
		61	62	63								
		b	С	b								
				•								



9 <b>1</b>	2	3	4	5	6	7	8	9	10
С	С	С	a	a	b	d	d	b	а
11	12	13	14	15	16	17	18	19	20
b	a	b	C	С	а	b	C	d	а
21	22	23	24	25	26	27	28	29	30
a	d	b	С	d	a	b	а	b	b
31	32	33	34	35	36	37	38	39	40
С	С	а	С	d	а	b	b	b	d
41	42	43	44	45	46	47	48	49	50
b	d	С	С	d	d	С	С	b	b
51	52	53	54	55	56	57	58	59	60
С	d	а	а	a	а	b	С	b	d
61	62	63							
h		h	1						