

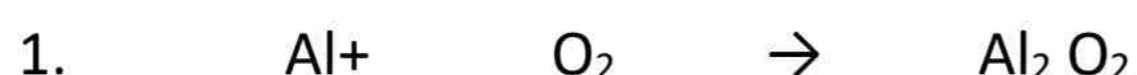
Chemistry Class IX

Guess Paper



Section B (Short Question Answer)

Q1) Balance the following chemical reaction.



Q2) What is the molarity of the solution prepared by dissolving 1.25 g of HCl gas into enough water to make 30 cm³ of solution?

OR----->What volume of 0.5M acid is needed to neutralize 200ml of 4M base?

Q3) A 600 ml sample of gas is heated from 27 °C to 77 °C at constant pressure. What is the final volume?

OR----->The 700 cm³ of a gas is enclosed in a container under a pressure of 650 mm of Hg. If the volume is reduced to 350 cm³, what will be the pressure then?

Q4) Calculate Number Of Moles in 40 gram of HNO₃

OR----->Calculate mass in 2.5 moles of Al

Q5) Define three different branches of chemistry

Q6) Distinguish between periods and groups

OR

Give properties of I-A OR VII-A, VIII-A group of modern periodic table

Q7) Differentiate between Amorphous and crystalline solid.

OR----->Define Allotropy. Name the different allotropies of carbon

OR----->What do you know about plasma state? Also give examples of plasma state.

Q8) State Faraday's first law of Electrolysis. Also explain it. Also explain Faraday's second law of Electrolysis.

Q9) Compare the properties of Alkali and Alkaline Earth Metals.

OR----->Differentiate between metals and nonmetals.

Q10) Write the Rutherford atomic model defects.

OR----->What is an isotope? Discuss the isotopes of Hydrogen

Q11) Distinguish polar and non-polar bond.

OR----->Write some properties of Coordinate Covalent Bond.

Q12) Differentiate between oxidation and reduction reaction

OR----->Differentiate between Saturated and Unsaturated Solution

Section C (Detailed Question Answer)

Q13) Describe the discovery of cathode rays.

OR----->Discuss Rutherford gold metal foil experiment in the light of structure of atom.

OR----->Write the Rutherford atomic model defects. Also Describe the Schrodinger atomic model.

Q14) Define Ionic bonds. Explain the formation of ionic bond in sodium chloride NaCl.

OR----->Discuss in detail the long form of periodic table.

Q15) State and explain Boyle's law. Also establish a relation between volume and pressure of a gas

Q16) Define solubility. Give the general principles of solubility.

Q17) What is corrosion. Give methods for prevention form corrosion.

